

MODELLING AND SIMULATION OF ADSORPTION PROCESS FOR  
REMOVAL OF CO<sub>2</sub> FROM NATURAL GAS IN AN OFFSHORE PLATFORM

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*To my beloved husband  
for understanding the days I was away and to my beloved mother for devoting her  
time to make sure I complete my studies  
Thank You!*

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## ABSTRACT

The presence of CO<sub>2</sub> in natural gas causes pipeline corrosion and increases operating costs during transfer from the offshore production platforms to the storage terminal. Due to space limitation and harsh operating environment, a robust and compact process such as pressure swing adsorption is preferable. To facilitate the study of process dynamics, simulation studies based on a derived mathematical model on a MATLAB software is presented. The effect of design parameters, focusing on the column height is considered, and it is found that for a typical laboratory scaled apparatus having diameter of 0.5 m. The maximum height required to adsorb 99 % CO<sub>2</sub> is 3 m when the feed flow rate is fixed at 2.5 m<sup>3</sup>/s. The size of adsorbent particles is also impacting separation efficiency, and the optimum particle radius is found to be 1.25x10<sup>-3</sup> m and the bed porosity was 0.2. Sensitivity analyses on the main operating parameters are also investigated. It is found that the initial CO<sub>2</sub> feed composition has positive relationship to the adsorption efficiency. The 0.4 mole fraction was found to have sufficient separation efficiency of 90 %. The model is also tested for representing a typical industrial operation with 120 mmscfd. In this case, for a 4 m diameter column, a column height of 20 m is required. This is achieved with a 4 bed PSA system at a flow rate of 10.05 m<sup>3</sup>/s for each, and an optimum separation of 87 % is established. Based on the results obtained in this work it can be concluded that the model is a reasonable representation of the system and can be used to obtain the necessary process insights for further process development.

## ABSTRAK

Kehadiran CO<sub>2</sub> dalam gas asli menyebabkan kakisan saluran paip dan meningkatkan kos operasi semasa pemindahan dari platform pengeluaran luar pesisir ke terminal penyimpanan. Oleh kerana batasan ruang dan persekitaran operasi yang teruk, proses yang mantap dan padat seperti penjerapan tekanan swing lebih baik. Untuk memudahkan kajian dinamik proses, kajian simulasi berdasarkan model matematik yang diperolehi pada perisian MATLAB dibentangkan. Kesan parameter reka bentuk, yang memberi tumpuan kepada ketinggian lajur adalah dipertimbangkan, dan didapati bahawa untuk alat ukur makmal tipikal yang mempunyai diameter 0.5m. Ketinggian maksimum yang diperlukan untuk menyerap 99% CO<sub>2</sub> ialah 3 m apabila kadar aliran suapan ditetapkan pada 2.5m<sup>3</sup>/s. Saiz zarah penyerap juga memberi kesan kepada kecekapan pemisahan, dan radius zarah optimum didapati 1.25x10<sup>-3</sup> m dan porositi katil adalah 0.2. Analisis kepekaan terhadap parameter operasi utama juga disiasat. Telah didapati bahawa komposisi makanan CO<sub>2</sub> awal mempunyai hubungan positif dengan kecekapan penjerapan. Pecutan 0.4 mol didapati mempunyai kecekapan pemisahan yang mencukupi sebanyak 90%. Model ini juga diuji untuk mewakili operasi perindustrian biasa dengan 120 MMSCFD. Dalam kes ini, bagi lajur diameter 4 m, ketinggian lajur 20 m diperlukan. Ini dicapai dengan sistem PSA 4 katil pada kadar aliran 10.05 m<sup>3</sup>/s untuk setiap satu, dan pemisahan optimum 87% ditubuhkan. Berdasarkan hasil yang diperolehi dalam karya ini, dapat disimpulkan bahawa model adalah representasi yang wajar dari sistem dan dapat digunakan untuk memperoleh wawasan proses yang diperlukan untuk pengembangan proses selanjutnya.

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**LIST OF ABBREVIATIONS**

CCS	-	Carbon Capture and Storage
CO <sub>2</sub>	-	Carbon dioxide
FPSO	-	Floating Production Storage Off-loading
GDP	-	Gross Domestic Product
LPH	-	Litter per hour
MEA	-	Monoethanol Amine
MOFs	-	Metal-organic frameworks
MOPs	-	Microporous organic polymers
MSA	-	Microwave Swing Adsorption
PI	-	Process Intensification
Ppmv	-	parts per million volume
PSA	-	Pressure Swing Adsorption
TSA	-	Temperature Swing Adsorption

**LIST OF SYMBOLS**

$L$	-	Column Length
$D$	-	Column Diameter
$\varepsilon$	-	Voidage
$R_p$	-	Particle Radius
$kt$	-	Volumetric Mass Transfer Coefficient
$u$	-	Gas Velocity
$K$	-	Freundlich Constants
$n$	-	Freundlich Constants
$T$	-	Temperature
$R$	-	Universal Gas Constant
$\rho_s$	-	Density of Solid
$P$	-	Pressure
$\dot{V}$	-	Volumetric Flow Rate
$\dot{V}^0$	-	Standard Volumetric Flow Rate
$\dot{M}$	-	Mass Flow Rate
$C$	-	CO <sub>2</sub> Concentration
$q$	-	Amount of CO <sub>2</sub> Adsorbed
$q_e$	-	Equilibrium Quantity Adsorbed
$C_p$	-	Specific heat capacity

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## CHAPTER 1

### INTRODUCTION

#### 1.1 Background of the Study

Natural gas is a good candidate replacement fuel for heavy fossil fuel such as fuel oil or diesel, commonly used in power plants (Beronich *et al.*, 2009). It is a gas formed from decomposition of the animals and plants that died over 400 million years. It is used for both lighting and heating. Nearly one quarter of global energy production is from natural gas, with which at the year 2004 had an increasing rate of 3.3 % (Group, 2005). Natural gas is very attractive source of energy because it has low CO<sub>2</sub> emission as compared to other fuels like coal and oil (Beronich *et al.*, 2009). Moreover, NG has been widely studied as an important alternative for the expansion of the world's energy supply (Rios *et al.*, 2011; Walton *et al.*, 2006).

In natural gas, CO<sub>2</sub> is found at an average composition of 0.5 to 10 vol.% (Rios *et al.*, 2013). However, in certain fields, the content can be much higher. The presence of CO<sub>2</sub> in natural gas has several disadvantages such as lowering the heating value. Also, the volume of natural gas to be handled is increased and thus increasing the handling cost (Ahmed and Ahmad, 2011). The presence of water in natural gas causes reactions between water and CO<sub>2</sub> producing acid which is responsible for corrosion of metal pipes (Birkelund, 2013). Thus, most of the gas treatment facilities are designed with the limitation to treat natural gas with reduced CO<sub>2</sub> content by 30 to 40 mole %. Therefore, there is a need to remove the CO<sub>2</sub> at source to reduce transportation load and the risks of pipeline corrosion. This is even

more important for the case of natural gas produced in offshore fields, where gases are transported by pipelines to onshore gas terminals.

Typically, CO<sub>2</sub> is removed either using membrane processes, chemical absorption or pressure swing adsorption. However, membrane separation suffers from gas losses due to membrane saturation, the life time and durability of membrane and chemicals used for absorption are poor (Ahmed and Ahmad, 2011). Also, aqueous amine solution has effects of corrosion, high energy consumption for regeneration and requires large volume of absorber (Veawab et al., 1999). In addition, there is high operation cost due the fact that both membrane and chemical solvents separation requires pre- treatment to remove the impurities that may degrade the solvent or membrane. Based on these disadvantages more stable, efficient and cost-effective gas separation methods are to be developed to overcome such challenges.

A preferred alternative, especially from robustness perspectives is adsorption. However, there are several issues needed to be addressed. These include developments of better adsorbent and better process configurations especially in the context of offshore production platforms where space is limited. A compact and efficient process is therefore needed as the system is constrained by the limited space in these offshore production platforms.

## **1.2 Problem Statement**

On-site removal of CO<sub>2</sub> is advantageous as it reduces all the associated costs and problems in transporting the overall costs of transporting natural gas using pipelines from offshore production platforms. This is however, a challenging task because of space limitations (Dalane et al., 2017; Mazzetti et al., 2014). Although there are many options available, most are not practicable due to various issues. Membrane separation suffers from durability issue, while amine absorption liquid amine that requires large space, energy intensive regeneration cycle and waste disposal issue. On the contrary, Pressure Swing Adsorption (PSA), which is a robust

physical process offers a clear advantage. However, in order to design a compact system to be applied in harsh processing environment. The process must be well understood. This requires detailed studies on the sensitivity and dynamics of the process variables, which can be facilitated through mathematical modelling and simulation studies of the PSA system. Based on these understandings, detailed equipment design can later be carried out to suit the desired application.

### **1.3 Research Objectives**

The aim of the work is to study the characteristics of PSA process for separating CO<sub>2</sub> from natural gas mixture. This is achieved by satisfying the following research objectives:

- i. To develop mathematical model of the CO<sub>2</sub> adsorption process and simulate the model using MATLAB software
- ii. To carry out sensitivity analysis of key design and process variables and their influence on the separation of CO<sub>2</sub> from natural gas stream that contains high CO<sub>2</sub>

### **1.4 Scope of Study**

This research focused on simulation of CO<sub>2</sub> adsorption. There are many parameters that can be studied to determine the optimum adsorption conditions for CO<sub>2</sub>. For the purpose of this research, effects of gas velocity, height of the column and initial CO<sub>2</sub> composition in natural gas was studied. The sizing of the bed was optimized and the optimum operating conditions were examined. The height variation was from 1m to 4m at the small scale which was further optimized for the industrial operation. The initial CO<sub>2</sub> composition (mole fraction) was varied from 0.2 to 0.6 this range was selected because the given CO<sub>2</sub> mole fraction natural gas is around 0.4 in most of the fields. In addition, the feed flow rate was varied from



1m<sup>3</sup>/s to 5 m<sup>3</sup>/s which was then optimized to accommodate the actual flow rate at the gas processing plants.

## **1.5 Significance of Study**

The ability to remove CO<sub>2</sub> on-site in an offshore platform with a reduced equipment size offer major benefits to the industry. The outcome of this study provides a useful guide for studying various alternatives for system configurations to be applied in space scarce processing environment such as off-shore production platform and Floating Production Storage Off-loading (FPSO) system.

The compact size of PSA at the offshore gas processing plants is more efficient since less space is needed. Also there will be a needed of less construction materials since the size of columns needed are smaller as compared to the conventional one. Thus reducing both installation and maintenance costs. Moreover the use of new adsorbent will also be significant because smaller column size require less adsorbent and that most of the time new adsorbent are available in abundant form.

## REFERENCES

- Adams, D. (2010). *Flue gas treatment for CO<sub>2</sub> capture*: IEA Clean Coal Centre London.
- Adewole, J., Ahmad, A., Ismail, S., and Leo, C. (2013). Current challenges in membrane separation of CO<sub>2</sub> from natural gas: a review. *International Journal of Greenhouse Gas Control*, 17, 46-65.
- Ahmad, F., Lau, K., Shariff, A., and Murshid, G. (2012). Process simulation and optimal design of membrane separation system for CO<sub>2</sub> capture from natural gas. *Computers & Chemical Engineering*, 36, 119-128.
- Ahmed, T. Y., and Ahmad, M. M. (2011). Flowsheet development and simulation of off-shore carbon dioxide removal system at natural gas reserves. *International Journal*, 2(1).
- Anbia, M., and Hoseini, V. (2012). Enhancement of CO<sub>2</sub> adsorption on nanoporous chromium terephthalate (MIL-101) by amine modification. *Journal of Natural Gas Chemistry*, 21(3), 339-343.
- Barchas, R., and Davis, R. (1992). The Kerr-McGee/ABB Lummus Crest technology for the recovery of CO<sub>2</sub> from stack gases. *Energy Conversion and Management*, 33(5), 333-340.
- Bastin, L., Barcia, P. S., Hurtado, E. J., Silva, J. A., Rodrigues, A. E., and Chen, B. (2008). A microporous metal–organic framework for separation of CO<sub>2</sub>/N<sub>2</sub> and CO<sub>2</sub>/CH<sub>4</sub> by fixed-bed adsorption. *The Journal of Physical Chemistry C*, 112(5), 1575-1581.
- Ben-Mansour, R., Habib, M., Bamidele, O., Basha, M., Qasem, N., Peedikakkal, A., et al. (2016). Carbon capture by physical adsorption: materials, experimental investigations and numerical modeling and simulations—a review. *Applied Energy*, 161, 225-255.

- Beronich, E. L., Abdi, M. A., and Hawboldt, K. A. (2009). Prediction of natural gas behaviour in loading and unloading operations of marine CNG transportation systems. *Journal of Natural Gas Science and Engineering*, 1(1), 31-38.
- Biegler, L. T., Jiang, L., and Fox, V. G. (2005). Recent advances in simulation and optimal design of pressure swing adsorption systems. *Separation & Purification Reviews*, 33(1), 1-39.
- Birkelund, E. S. (2013). *CO2 Absorption and Desorption Simulation with Aspen HYSYS*. Universitetet i Tromsø.
- Bonjour, J., Chalfen, J.-B., and Meunier, F. (2002). Temperature swing adsorption process with indirect cooling and heating. *Industrial & engineering chemistry research*, 41(23), 5802-5811.
- Cao, D., and Wu, J. (2005). Modeling the selectivity of activated carbons for efficient separation of hydrogen and carbon dioxide. *Carbon*, 43(7), 1364-1370.
- Chavez, R.-H., and Guadarrama, J. d. J. (2011). A numerical simulation study of CO2 capture process for an electric central. *Energy Procedia*, 4, 2096-2103.
- Chen, Z., Deng, S., Wei, H., Wang, B., Huang, J., and Yu, G. (2013). Activated carbons and amine-modified materials for carbon dioxide capture—a review. *Frontiers of Environmental Science & Engineering*, 7(3), 326-340.
- Cherbański, R., and Molga, E. (2009). Intensification of desorption processes by use of microwaves—an overview of possible applications and industrial perspectives. *Chemical Engineering and Processing: Process Intensification*, 48(1), 48-58.
- Chronopoulos, T. (2016). *Microwave Swing Adsorption for post-combustion CO2 capture from flue gases using solid sorbents*. Heriot-Watt University.
- Couck, S., Denayer, J. F., Baron, G. V., Rémy, T., Gascon, J., and Kapteijn, F. (2009). An amine-functionalized MIL-53 metal-organic framework with large separation power for CO2 and CH4. *Journal of the American Chemical Society*, 131(18), 6326-6327.
- Dąbrowski, A. (2001). Adsorption—from theory to practice. *Advances in colloid and interface science*, 93(1), 135-224.
- Dalane, K., Dai, Z., Mogseth, G., Hillestad, M., and Deng, L. (2017). Potential applications of membrane separation for subsea natural gas processing: A review. *Journal of Natural Gas Science and Engineering*, 39, 101-117.

- De Groot, S. R., and De Groot, S. R. (1951). *Thermodynamics of irreversible processes* (Vol. 336): North-Holland Amsterdam.
- Ding, Y., and Alpay, E. (2001). High temperature recovery of CO<sub>2</sub> from flue gases using hydrotalcite adsorbent. *Process Safety and Environmental Protection*, 79(1), 45-51.
- Dortmundt, D., and Doshi, K. (1999). Recent development in CO<sub>2</sub> removal membranes. *UOP LLC*.
- Dotto, G., Santos, J., Rodrigues, I., Rosa, R., Pavan, F., and Lima, E. (2015). Adsorption of Methylene Blue by ultrasonic surface modified chitin. *Journal of Colloid and Interface Science*, 446, 133-140.
- Duwahir, Z. M. (2016). *Capturing CO<sub>2</sub> from an integrated steel mill: a techno-economic analysis through process modelling*. University of Nottingham.
- EIA. (2017). International Energy Outlook 2017.
- Emrani, A. S., Saber, M., and Farhadi, F. (2011). A Decision Tree for Technology Selection of Nitrogen Production Plants.
- Figueroa, J. D., Fout, T., Plasynski, S., McIlvried, H., and Srivastava, R. D. (2008). Advances in CO<sub>2</sub> capture technology—the US Department of Energy's Carbon Sequestration Program. *International journal of greenhouse gas control*, 2(1), 9-20.
- Finsky, V., Ma, L., Alaerts, L., De Vos, D., Baron, G., and Denayer, J. (2009). Separation of CO<sub>2</sub>/CH<sub>4</sub> mixtures with the MIL-53 (Al) metal-organic framework. *Microporous and Mesoporous Materials*, 120(3), 221-227.
- Galizia, M., Chi, W. S., Smith, Z. P., Merkel, T. C., Baker, R. W., and Freeman, B. D. (2017). 50th Anniversary Perspective: Polymers and Mixed Matrix Membranes for Gas and Vapor Separation: A Review and Prospective Opportunities. *Macromolecules*, 50(20), 7809-7843.
- Gibson, J. A., Mangano, E., Shiko, E., Greenaway, A. G., Gromov, A. V., Lozinska, M. M., et al. (2016). Adsorption materials and processes for carbon capture from gas-fired power plants: AMPGas. *Industrial & Engineering Chemistry Research*, 55(13), 3840-3851.
- Glueckauf, E. (1955). Theory of chromatography. Part 10.—Formulae for diffusion into spheres and their application to chromatography. *Transactions of the Faraday Society*, 51, 1540-1551.

- Golmakani, A., Fatemi, S., and Tamnanloo, J. (2017). Investigating PSA, VSA, and TSA methods in SMR unit of refineries for hydrogen production with fuel cell specification. *Separation and Purification Technology*, 176, 73-91.
- Grande, C. A. (2012). Advances in pressure swing adsorption for gas separation. *ISRN Chemical Engineering*, 2012.
- Group, B. (2005). BP statistical review of world energy. *June 2010*.
- Hansen, J., Johnson, D., Lacis, A., Lebedeff, S., Lee, P., Rind, D., et al. (1981). Climate impact of increasing atmospheric carbon dioxide. *Science*, 213(4511), 957-966.
- Hauchhum, L., and Mahanta, P. (2014). Carbon dioxide adsorption on zeolites and activated carbon by pressure swing adsorption in a fixed bed. *International Journal of Energy and Environmental Engineering*, 5(4), 349-356.
- Ho, M. T., Allinson, G. W., and Wiley, D. E. (2008). Reducing the cost of CO<sub>2</sub> capture from flue gases using pressure swing adsorption. *Industrial & Engineering Chemistry Research*, 47(14), 4883-4890.
- Ho, Y., and McKay, G. (1998). A comparison of chemisorption kinetic models applied to pollutant removal on various sorbents. *Process safety and environmental protection*, 76(4), 332-340.
- Hwang, K. S., Jun, J. H., and Lee, W. K. (1995). Fixed-bed adsorption for bulk component system. Non-equilibrium, non-isothermal and non-adiabatic model. *Chemical Engineering Science*, 50(5), 813-825.
- Kim, S.-N., Son, W.-J., Choi, J.-S., and Ahn, W.-S. (2008). CO<sub>2</sub> adsorption using amine-functionalized mesoporous silica prepared via anionic surfactant-mediated synthesis. *Microporous and Mesoporous Materials*, 115(3), 497-503.
- Lee, S.-Y., and Park, S.-J. (2015). A review on solid adsorbents for carbon dioxide capture. *Journal of Industrial and Engineering Chemistry*, 23, 1-11.
- Leung, D. Y., Caramanna, G., and Maroto-Valer, M. M. (2014). An overview of current status of carbon dioxide capture and storage technologies. *Renewable and Sustainable Energy Reviews*, 39, 426-443.
- Li, B., He, G., Jiang, X., Dai, Y., and Ruan, X. (2016). Pressure swing adsorption/membrane hybrid processes for hydrogen purification with a high recovery. *Frontiers of Chemical Science and Engineering*, 10(2), 255-264.

- Liu, Y., Bisson, T. M., Yang, H., and Xu, Z. (2010). Recent developments in novel sorbents for flue gas clean up. *Fuel processing technology*, 91(10), 1175-1197.
- Lu, A.-H., Hao, G.-P., and Zhang, X.-Q. (2014). Porous carbons for carbon dioxide capture. In *Porous Materials for Carbon Dioxide Capture* (pp. 15-77): Springer.
- Luis, P., and Bruggen, B. (2013). The role of membranes in post-combustion CO<sub>2</sub> capture. *Greenhouse Gases: Science and Technology*, 3(5), 318-337.
- Mazlee, M. (2016). *Sustainable catalyst supports for carbon dioxide gas adsorbent*. Paper presented at the AIP Conference Proceedings, 020002.
- Mazzetti, M. J., Neksa, P., Walnum, H. T., and Hemmingsen, A. K. T. (2014). Energy-Efficiency Technologies for Reduction of Offshore CO<sub>2</sub> Emissions. *Oil and Gas Facilities*, 3(01), 89-96.
- Metz, B., Davidson, O., De Coninck, H., Loos, M., and Meyer, L. (2005). *IPCC special report on carbon dioxide capture and storage: Intergovernmental Panel on Climate Change, Geneva (Switzerland). Working Group IIIo. Document Number)*
- Nouh, S., Lau, K., and Shariff, A. (2010). Modeling and simulation of fixed bed adsorption column using Integrated CFD Approach. *Journal of Applied Sciences(Faisalabad)*, 10(24), 3229-3235.
- Nurbas, M., Kaçar, Y., and Kutsal, T. (2002). Determining the overall mass transfer coefficient for adsorption of Cu<sup>2+</sup> ions onto Ca-Alginate in fixed bed column. *Eur. J. Miner. Process. Environ. Prot*, 2, 55-60.
- Oreggioni, G., Friedrich, D., Luberti, M., Ahn, H., and Brandani, S. (2016). Development of an equilibrium theory solver applied to pressure swing adsorption cycles used in carbon capture processes. *Computers & Chemical Engineering*, 94, 18-27.
- Özcan, A., Öncü, E. M., and Özcan, A. S. (2006). Kinetics, isotherm and thermodynamic studies of adsorption of Acid Blue 193 from aqueous solutions onto natural sepiolite. *Colloids and Surfaces A: Physicochemical and Engineering Aspects*, 277(1), 90-97.
- Park, J. H., Kim, J. N., and Cho, S. H. (2000). Performance analysis of four-bed H<sub>2</sub> PSA process using layered beds. *AIChE Journal*, 46(4), 790-802.

- Rashidi, N. A., Yusup, S., and Lam, H. L. (2013). Kinetic studies on carbon dioxide capture using activated carbon.
- Ravanchi, M. T., Kaghazchi, T., and Kargari, A. (2009). Application of membrane separation processes in petrochemical industry: a review. *Desalination*, 235(1), 199-244.
- Rios, R., Stragliotto, F., Peixoto, H., Torres, A., Bastos-Neto, M., Azevedo, D., et al. (2013). Studies on the adsorption behavior of CO<sub>2</sub>-CH<sub>4</sub> mixtures using activated carbon. *Brazilian journal of chemical engineering*, 30(4), 939-951.
- Rios, R. B., Bastos-Neto, M., Amora, M. R., Torres, A. E. B., Azevedo, D. C., and Cavalcante, C. L. (2011). Experimental analysis of the efficiency on charge/discharge cycles in natural gas storage by adsorption. *Fuel*, 90(1), 113-119.
- Roy, P. S., and Amin, M. R. (2012). Aspen-HYSYS simulation of natural gas processing plant. *Journal of Chemical Engineering*, 26(1), 62-65.
- Rubin, E. S., and Rao, A. B. (2002). *A technical, economic and environmental assessment of amine-based CO<sub>2</sub> capture technology for power plant greenhouse gas control*: National Energy Technology Lab., Pittsburgh, PA (US); National Energy Technology Lab., Morgantown, WV (US). (Document Number)
- Ruthven, D. M. (1984). *Principles of adsorption and adsorption processes*: John Wiley & Sons.
- Sayari, A., Belmabkhout, Y., and Serna-Guerrero, R. (2011). Flue gas treatment via CO<sub>2</sub> adsorption. *Chemical Engineering Journal*, 171(3), 760-774.
- Schueller, B. S., and Yang, R. T. (2001). Ultrasound enhanced adsorption and desorption of phenol on activated carbon and polymeric resin. *Industrial & engineering chemistry research*, 40(22), 4912-4918.
- Shafeeyan, M. S., Daud, W. M. A. W., Houshmand, A., and Arami-Niya, A. (2011). *Statistical Modeling and Optimization of Amination Conditions of Activated Carbon for Carbon Dioxide Adsorption Using Response Surface Methodology*. Paper presented at the Proceedings of 3rd International Conference on Chemical, Biological and Environmental Engineering (ICBEE 2011), 13-18.

- Shafeeyan, M. S., Daud, W. M. A. W., and Shamiri, A. (2014). A review of mathematical modeling of fixed-bed columns for carbon dioxide adsorption. *Chemical engineering research and design*, 92(5), 961-988.
- Sheindorf, C., Rebhun, M., and Sheintuch, M. (1981). A Freundlich-type multicomponent isotherm. *Journal of Colloid and Interface Science*, 79(1), 136-142.
- Shimekit, B., and Mukhtar, H. (2012). Natural gas purification technologies-major advances for CO<sub>2</sub> separation and future directions. In *Advances in Natural Gas Technology*: InTech.
- Siefers, A., Wang, N., Sindt, A., Dunn, J., McElvogue, J., Evans, E., et al. (2010). A novel and cost-effective hydrogen sulfide removal technology using tire derived rubber particles. *Proceedings of the Water Environment Federation*, 2010(12), 4597-4622.
- Simo, M., Brown, C. J., and Hlavacek, V. (2008). Simulation of pressure swing adsorption in fuel ethanol production process. *Computers & Chemical Engineering*, 32(7), 1635-1649.
- Singh, V. K., and Kumar, E. A. (2016). Comparative Studies on CO<sub>2</sub> Adsorption Kinetics by Solid Adsorbents. *Energy Procedia*, 90, 316-325.
- Skarstrom, C. W. (1960). Method and apparatus for fractionating gaseous mixtures by adsorption: Google Patents.
- Solomon, I., Spiridon, M., Secula, M. S., and Petrescu, S. (2010). STUDY OF GAS SEPARATION PROCESS BY DYNAMIC ADSORPTION IN FIXED BED. *Scientific Study & Research. Chemistry & Chemical Engineering*, 11(3), 329-340.
- Song, C., Kansha, Y., Ishizuka, M., Fu, Q., and Tsutsumi, A. (2015). Conceptual design of a novel pressure swing CO<sub>2</sub> adsorption process based on self-heat recuperation technology. *Chemical Engineering and Processing: Process Intensification*, 94, 20-28.
- Spigarelli, B. P., and Kawatra, S. K. (2013). Opportunities and challenges in carbon dioxide capture. *Journal of CO<sub>2</sub> Utilization*, 1, 69-87.
- Stankiewicz, A. I., and Moulijn, J. A. (2000). Process intensification: transforming chemical engineering. *Chemical engineering progress*, 96(1), 22-34.
- Stewart, E., and Lanning, R. (1994). Reduce amine plant solvent losses; Part 1. *Hydrocarbon Processing;(United States)*, 73(5).



- Tabe-Mohammadi, A. (1999). A review of the applications of membrane separation technology in natural gas treatment. *Separation science and technology*, 34(10), 2095-2111.
- Talebian-Kiakalaieh, A., and Amin, N. A. S. (2016). Theoretical and experimental evaluation of mass transfer limitation in gas phase dehydration of glycerol to acrolein over supported HSiW catalyst. *Journal of the Taiwan Institute of Chemical Engineers*, 59, 11-17.
- Thomas, W. J., and Crittenden, B. D. (1998). *Adsorption technology and design*: Butterworth-Heinemann.
- Trivedi, H., Patel, V., and Patel, R. (1973). Adsorption of cellulose triacetate on calcium silicate. *European Polymer Journal*, 9(6), 525-531.
- Ünveren, E. E., Monkul, B. Ö., Sariođlan, Ş., Karademir, N., and Alper, E. (2016). Solid amine sorbents for CO<sub>2</sub> capture by chemical adsorption: a review. *Petroleum*.
- Veawab, A., Tontiwachwuthikul, P., and Chakma, A. (1999). Corrosion behavior of carbon steel in the CO<sub>2</sub> absorption process using aqueous amine solutions. *Industrial & engineering chemistry research*, 38(10), 3917-3924.
- Wakao, N., and Funazkri, T. (1978). Effect of fluid dispersion coefficients on particle-to-fluid mass transfer coefficients in packed beds: correlation of Sherwood numbers. *Chemical Engineering Science*, 33(10), 1375-1384.
- Walton, K., Cavalcante Jr, C., and LeVan, M. D. (2006). Adsorption of light alkanes on coconut nanoporous activated carbon. *Brazilian Journal of Chemical Engineering*, 23(4), 555-561.
- Wang, Q., Luo, J., Zhong, Z., and Borgna, A. (2011). CO<sub>2</sub> capture by solid adsorbents and their applications: current status and new trends. *Energy & Environmental Science*, 4(1), 42-55.
- Weiland, R. H., and Sivasubramanian, M. (2004). *Effect of Heat-Stable Salts on Amine Absorber and Regenerator Performance*. Paper presented at the Fall Meeting of AIChE, Austin, Texas.
- Whitney, G., and Behrens, C. E. (2010). *Energy: Natural Gas: The Production and Use of Natural Gas, Natural Gas Imports and Exports, Epack Project, Liquefied Natural: The Capitol Net Inc*.
- Williams, R. H. (2001). Toward zero emissions from coal in China. *Energy for Sustainable Development*, 5(4), 39-65.

- Won, W., Lee, S., and Lee, K.-S. (2012). Modeling and parameter estimation for a fixed-bed adsorption process for CO<sub>2</sub> capture using zeolite 13X. *Separation and Purification Technology*, 85, 120-129.
- Xu, D., Zhang, J., Li, G., Xiao, P., Webley, P., and Zhai, Y.-C. (2012). Adsorption equilibrium and kinetics of CO (2) and H (2) O on activated carbon. *Wuji Cailiao Xuebao(Journal of Inorganic Materials)*, 27(2), 139-145.
- Yaumi, A., and Mustafa, M. A. B. Solid materials for carbon dioxide Adsorption: A Review.
- Yeo, Z. Y., Chew, T. L., Zhu, P. W., Mohamed, A. R., and Chai, S.-P. (2012). Conventional processes and membrane technology for carbon dioxide removal from natural gas: a review. *Journal of Natural Gas Chemistry*, 21(3), 282-298.
- Yong, Z., Mata, V., and Rodrigues, A. r. E. (2002). Adsorption of carbon dioxide at high temperature—a review. *Separation and Purification Technology*, 26(2), 195-205.
- Yu, C.-H., Huang, C.-H., and Tan, C.-S. (2012). A review of CO<sub>2</sub> capture by absorption and adsorption. *Aerosol Air Qual. Res*, 12(5), 745-769.
- Zheng, F., Addleman, R. S., Aardahl, C. L., Fryxell, G. E., Brown, D. R., and Zemanian, T. S. (2007). Amine functionalized nanoporous materials for carbon dioxide (CO<sub>2</sub>) capture. *Environmental applications of nanomaterials*, 285-312.