

GASES BREAKTHROUGH ADSORPTION ON ACTIVATED  
CARBON OF POROUS SYNTHESIZED RENEWABLE MATERIAL

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POROUS SYNTHESIZED RENEWABLE MATERIAL

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Specially dedicated to MY LOVELY MOM AND DAD,

*Mohd Zain Bin Mahmood*

*and*

*Siti Rokiah Binti Abdullah.*

“To my family, lecturers and friends, who supported me each step of the way”

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## ABSTRACT

Acidic and alkaline gases are the types of pollutant which may give negative impact on human life as well as environment without proper treatment. It is important to ensure that these types of gases concentration in oil and gas industry are up to regulatory standard before it can be released to surrounding. Activated carbon is one of the most effective adsorbents used in hazardous environmental treatments and it is considered to be sustainable, environmentally friendly, economical and efficient making it a superior and the most commonly used in adsorption process compared to other adsorbents. The research is aimed to synthesize agro-based solid waste materials to activated carbon (AC) and to determine the breakthrough time adsorption isotherm of single poisoning gases on the AC. In this study, palm empty fruit bunch and palm kernel shell were selected as the raw materials to produce AC. Char was produced through carbonization process and then undergone chemical treatments [potassium hydroxide (KOH), iron (III) chloride hexahydrate ( $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ ) and zinc nitrate hexahydrate ( $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ )] followed by microwave treatment in order to produce three different AC. Effect of different gases on the sorbent performance for adsorption breakthrough study were investigated. Testing of the prepared samples was accomplished using adsorption column breakthrough experiment. The breakthrough time adsorption testing was performed at 100 mL/min flowrate for 3.6 g of sample. The samples were characterized by thermo-gravimetric analysis, scanning electron microscope, Fourier transform infrared and nitrogen adsorption isotherm. Nitrogen adsorption isotherm results showed that the commercial activated carbon, palm kernel activated carbon treated with KOH-Zn (PKAC-KOH-Zn), palm kernel activated carbon treated with KOH- $\text{FeCl}_3$  (PKAC-KOH- $\text{FeCl}_3$ ) and palm empty fruit bunch activated carbon treated with KOH (EFBAC-KOH) have Brunauer-Emmett-Teller surface area of 1,005.87, 259.17, 122.61 and 3.48  $\text{m}^2/\text{g}$ , respectively. For adsorption breakthrough study, the longer breakthrough time for acidic and alkaline gas adsorption implied better adsorption breakthrough performance. The end results showed that breakthrough times for three type of AC were not much different as compared to commercial AC. PKAC-KOH- $\text{FeCl}_3$  was the best material for carbon monoxide and sulfur dioxide adsorption with 6.33 s and 51 s breakthrough time, respectively. PKAC-KOH-Zn produced the longest breakthrough time for chlorine adsorption of 1002.5 s and EFBAC-KOH produced the longest breakthrough time for ammonia adsorption of 175.67 s. As a conclusion, the findings revealed the potential of AC derived from palm kernel shell and palm empty fruit bunch as poisoning gases sorbents based on the good results obtained from characterization analysis and adsorption breakthrough study.

## ABSTRAK

Gas berasid dan beralkali adalah jenis bahan cemar yang boleh memberi kesan negatif kepada kehidupan manusia dan juga alam sekitar tanpa rawatan yang betul. Adalah penting untuk memastikan yang kepekatan gas-gas ini dalam industri minyak dan gas mencapai standard peraturan sebelum dilepaskan ke persekitaran. Karbon teraktif (AC) adalah salah satu bahan penjerap yang paling berkesan digunakan dalam rawatan-rawatan alam sekitar yang berbahaya dan dianggap mampan, mesra alam, menjimatkan dan berkesan yang menjadikannya lebih bagus dan yang paling biasa digunakan dalam proses penjerapan berbanding dengan penjerap-penjerap yang lain. Kajian ini adalah bertujuan untuk mensintesis bahan sisa pepejal asas tani kepada karbon teraktif dan menentukan jangka masa perintisan isoterma penjerapan gas beracun tunggal ke atas karbon teraktif. Dalam kajian ini, tandan kelapa sawit kosong dan tempurung kelapa sawit dipilih sebagai bahan-bahan mentah untuk menghasilkan AC. Arang dihasilkan melalui proses pengkarbonan dan kemudiannya melalui rawatan kimia [kalium hidroksida (KOH), besi klorida heksahidrat ( $\text{FeCl}_3 \cdot 6\text{H}_2\text{O}$ ) and zink nitrat heksahidrat ( $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ )] diikuti dengan rawatan gelombang mikro dalam menghasilkan tiga jenis AC yang berbeza. Kesan gas berbeza ke atas prestasi penjerap untuk kajian penjerapan perintisan telah dilakukan. Pengujian sampel-sampel yang telah disediakan dicapai dengan menggunakan eksperimen perintisan terus penjerapan. Ujian penjerapan masa perintisan telah dijalankan pada kadar alir 100 mL/min untuk sampel 3.6 g. Sampel telah dicirikan melalui analisis termo-gravimetri, mikroskop imbasan elektron, inframerah transformasi Fourier dan penjerapan nitrogen isoterma. Keputusan penjerapan nitrogen isoterma menunjukkan sampel karbon teraktif komersial, karbon teraktif tempurung kelapa sawit yang dirawat dengan menggunakan KOH-Zn (PKAC-KOH-Zn), karbon teraktif tempurung kelapa sawit yang dirawat dengan menggunakan KOH- $\text{FeCl}_3$  (PKAC-KOH- $\text{FeCl}_3$ ) dan karbon teraktif tandan kelapa sawit kosong yang dirawat dengan menggunakan KOH (EFBAC-KOH) mempunyai luas permukaan Brunauer-Emmett-Teller masing-masing 1,005.87, 259.17, 122.61 dan 3.48  $\text{m}^2/\text{g}$ . Untuk kajian perintisan penjerapan, semakin lama masa perintisan untuk penjerapan gas berasid dan beralkali bermakna prestasi perintisan penjerapan yang lebih baik. Keputusan akhir menunjukkan bahawa masa perintisan untuk tiga jenis AC adalah tidak banyak beza berbanding dengan karbon teraktif komersial. PKAC-KOH- $\text{FeCl}_3$  adalah bahan terbaik untuk penjerapan gas karbon monoksida dan sulfur dioksida dengan masa perintisan masing-masing 6.33 s dan 51 s. PKAC-KOH-Zn menghasilkan masa perintisan yang paling lama untuk penjerapan gas klorin dengan 1002.5 s dan EFBAC-KOH menghasilkan masa perintisan yang paling lama untuk penjerapan gas ammonia dengan 175.67 s. Kesimpulannya, hasil kajian mendedahkan potensi AC yang dihasilkan daripada tempurung kelapa sawit dan tandan kelapa sawit kosong sebagai penjerap gas beracun berdasarkan keputusan yang baik diperolehi daripada analisis pencirian dan kajian perintisan penjerapan.

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**LIST OF ABBREVIATIONS**

AC	-	Activated Carbon
ACFs	-	Activated carbon fiber
BET	-	Brunauer, Emmett and Teller
EFB	-	Empty Fruit Bunches
FTIR	-	Fourier Transform Infra-Red Spectroscopy
IDLH	-	Immediately Dangerous to Life and Health values
IUPAC	-	International Union of Pure and Applied Chemistry
MPOB	-	Malaysia Palm Oil Berhad
PAH	-	Polycyclic Aromatic Hydrocarbon
PEL	-	Permissible Exposure Limits
PKS	-	Palm kernel shell
PKC	-	Palm Kernel Char
POME	-	Palm Oil Mill Effluent
SEM	-	Scanning Electron Microscopy
STEL	-	Short Term Exposure Limit
TGA	-	Thermo-gravimetric analysis
TLV	-	Threshold Limit Values

## LIST OF SYMBOLS

A	-	Cross sectional area
$C_b$	-	Breakthrough concentration
$C_i$	-	Inlet concentration
$Cl_2$	-	Chlorine
CO	-	Carbon Monoxide
$CO_2$	-	Carbon Dioxide
C	-	Concentration of adsorbate at time t
cm	-	Centimeter
[C]	-	Ceiling Limit (never exceed)
$C_o$	-	Feed concentration of adsorbate
$^{\circ}C$	-	Degree Celsius
d	-	Diameter
%	-	Percentage
$FeCl_3$	-	Iron (II) chloride
GHz	-	Gigahertz
$H_2$	-	Hydrogen
$H_2O$	-	Water
$H_2S$	-	Hydrogen sulfide
$H_3PO_4$	-	Phosphoric acid
$H_2SO_4$	-	Sulfuric acid
$k_v$	-	Adsorption rate constant
kV	-	Kilo volt
KOH	-	Potassium hydroxide
$K_2CO_3$	-	Potassium carbonate
M	-	Molar / Molecular weight
$m_c$	-	Mass of adsorbent used inside the bed

mL/min	-	Milliliter per minute
n	-	Number of filter
N <sub>2</sub>	-	Nitrogen
NH <sub>3</sub>	-	Ammonia
NO <sub>x</sub>	-	Nitrogen oxide
NaOH	-	Sodium hydroxide
Na <sub>2</sub> CO <sub>3</sub>	-	Sodium carbonate
O <sub>2</sub>	-	Oxygen
Pa	-	Pascal
ρ <sub>c</sub>	-	Carbon density
PEEK	-	Polyetheretherketone
ppmv	-	Part per millions volume
q	-	Adsorbent adsorption capacity
Q	-	Inlet flow rate
Q <sub>f</sub>	-	Volumetric feed flow rate at STP
SO <sub>2</sub>	-	Sulfur Dioxide
S	-	Second
t	-	Time
t <sub>p</sub>	-	Breakthrough time
t <sub>t</sub>	-	Time equivalent to the total or stoichiometric capacity
W	-	Total weight of carbon
W <sub>e</sub>	-	Adsorption capacity
ZnCl <sub>2</sub>	-	Zinc chloride

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## **CHAPTER 1**

### **INTRODUCTION**

#### **1.1 Background of Study**

Malaysia generates abundant amounts of agricultural residues. The main agricultural crops planted in Malaysia are oil palm, coconut, cocoa, and rubber. Agricultural residue means that the by-products which are produced from processing of agricultural crops. The examples of the residues are paddy straw, palm empty fruit bunch, coconut shell, coconut husk, paddy husk, fiber and palm oil shell. Malaysia is presently the world's main exporter of palm oil despite the fact that it is the second largest manufacturer of the oil after Indonesia. Malaysia represents 39% of the world's palm oil generation and 44% of the world's exports (Malaysian Palm Oil Council, MPOC, 2012). The Malaysian oil palm industry is one of the most highly organised sectors of any national agriculture system of the world. Today, though, the focus has shifted to how well agriculture also meets universally accepted standards of sustainability. Oil palm is the most important product from Malaysia which enhances the agriculture and economy sectors. In order to create the probable for improvement of energy in the palm oil industry, there is necessary to consider disposal of mill residues and available uses (Abdullah *et al.*, 2009). Palm oil mills can produce many products such as palm kernels, fiber, and empty fruit bunches. These products are produced in huge quantities. It has the ability as a source of renewable energy and also can be converted from dried oil palm wastes into the products which contain various values. Usage of oil palm unwanted materials as a source of energy will lead another environmental advantage like decreasing in carbon dioxide that discharged (Lim, 2000). Many methods have been utilized in the

removal of the gas and liquid such as membrane separation, an adsorption method, chemical separation, distillation and others which give benefits to the industry. These methods contributed to the production of high-quality products, environmental friendly, and energy saving. However, adsorption is commonly used method due to its cost effective and outstandingly efficacious.

Adsorption on activated carbon has been proven to be better than others process for some treatments based on its simple design, fast adsorption kinetics and its ability to adsorb a wide range of pollutants efficiently. However, the available commercial activated carbons are still costly because of the starting material used are non-renewable and expensive i.e. coal, which is unjustified in pollution control applications (Yin *et al.*, 2007). Therefore, nowadays, many researchers have tried to produce activated carbons from renewable and cheaper precursors which were mainly agricultural by-products like biomass wastes. They produced activated carbon for removal of various pollutants. Many studies have been reported on the application of palm kernel shell (PKS) and empty fruit bunch (EFB). One of the potential products from PKS and EFB is activated carbon, and more interest has been shown in this area.

The present research focuses on the practical application. The activated carbon produced will be applied in the safety respiratory devices like gas masks. The utilization of the products commonly used in industries like manufacturing, oil and gas and petrochemicals which involve activities in the hazardous environment. The product is necessary and is required for all the people that involved in the unhealthy environment. Consequently, this study was planned to determine the ability of AC from PKS and EFB as acidic and alkaline gases adsorbents for the gas separation application.

## 1.2 Problem Statement

There are abundance of waste material from the palm-oil industry especially empty fruit bunches (EFB) material and palm kernel shell (PKS). The wastes will affect the environment which they are burning illegally. Therefore, some step should be taken to overcome the crisis in managing this waste which is by using the scientific method. Among all the wastes, PKS have been proven that the materials which having highest carbon content and surface area (Noor *et al.*, 2014). Moreover, previous studies (Sumathi *et al.*, 2015; Hesas *et al.*, 2013; Nasri *et al.*, 2014; Hamza *et al.*, 2015 and Lee *et al.*, 2015 ) reported that its ability to adsorb the gas and produced the good performance in term of adsorption capacity. Therefore, PKS has the potential to be converted into activated carbon (AC). Meanwhile, EFB wastes were selected as a precursor for the production of activated carbon due to the low cost and large quantity from palm oil industry.

From the scientific method, the EFB and PKS wastes were converted into value added products like activated carbon with proper method. Furthermore, nowadays, the commercial activated carbons were derived from natural materials like coal that contribute to the high costs of the adsorbents. The production of activated carbon by using biomass waste materials like EFB and PKS will minimize the environmental impacts and lead to the low cost of the adsorbents moreover, they are renewable and sustainable materials.

Acidic gases like carbon monoxide, carbon dioxide, chlorine, chlorine dioxide, sulfur dioxide and hydrogen sulfide are the types of contamination which may give the negative effect on human life and the environment without proper treatment. It is important to ensure that these types of gases concentration on oil and gas industry are up to regulatory standard before it can be removed to surrounding. Currently, the gases are adsorbed by conventional activated carbon. Nevertheless, the commercially activated carbon which predominantly produced by coal and petroleum is very costly and become depleting subsequently, has increased the treatment cost. In this manner, production of activated carbon derived by PKS and EFB can be applied as an adsorbent for acidic and alkaline gases treatment.

Therefore, this study focused on the best working conditions in a conversion of PKS and EFB to activated carbon for the mean of gas adsorption breakthrough.

Adsorption is a process of preferential removal of gas or liquid by adsorbing it on the surface of a solid. Activated carbon is most chosen as adsorbent due to many sources can be obtained from agricultural wastes which used as the precursor for the production of AC. The agricultural wastes are converted into activated carbon using chemical or physical activation or both.

Adsorption technology especially activated carbon is one of the most popular methods used in environmental pollution treatments and it is considered to be sustainable, environmentally friendly, simple, economical and efficient which makes it a superior and the most commonly used technique in adsorption compared to other methods (Auta and Hameed, 2013; Hesas *et al.*, 2013; Tan *et al.*, 2008; Tan *et al.*, 2009; Tham *et al.*, 2011), they also possess large surface area, high porosity, and rapid adsorption capabilities (Tham *et al.*, 2011; Dehdashti *et al.*, 2010; Prauchner and Rodriguez-Renoso, 2012).

### **1.3 Objective of Research**

This research is focused on synthesizing activated carbon from palm empty fruit bunch (EFB) and palm kernel shell (PKS) for capturing and separating the acid gases. The objectives include the following:

- i. To synthesize agro-base solid waste materials to activated carbon and to characterize the synthesized materials.
- ii. To determine the breakthrough time adsorption isotherm of single acid gases on the activated carbon.
- iii. To predict the saturation time based on simulation results.



## 1.4 Scope of Research

Palm empty fruit bunch and palm kernel shell were used as precursors for the production of the activated carbon. The raw palm EFB and PKS were washed and dried first before undergoing carbonization process under 150 cm<sup>3</sup>/min nitrogen gas flows at temperature 700<sup>0</sup>C and heating rate 10<sup>0</sup>C/min. The char produced from carbonization process was treated with chemical treatment by using potassium hydroxide (KOH) solution, Iron (III) chloride hexahydrate (FeCl<sub>3</sub>.6H<sub>2</sub>O) and Zinc nitrate hexahydrate, Zn (NO<sub>3</sub>)<sub>2</sub>.6H<sub>2</sub>O respectively. The activation process of char was prepared by microwave treatment. The physical activation of the char through microwave heating was done at 200 cm<sup>3</sup>/min CO<sub>2</sub> flow rate with 400W power level. The prepared activated carbon was then being characterized by Thermogravimetric Analysis (TGA), Scanning Electron Microscope (SEM), Fourier Transformation Infrared (FTIR), Nitrogen adsorption isotherm analysis and determination kit of density. The performance of the sample was then analyzed by poisoning gases adsorption test with 10ppm concentration at low flow rate (100mL/min) and at a high flow rate (6L/min) for mixing with the air. The activated carbon was tested by acidic gases like CO, SO<sub>2</sub>, Cl<sub>2</sub> and alkaline gas of NH<sub>3</sub>. Oxygen was tested to be used as the reference in the normal human breathing rate.

## 1.5 Significance of Study

Adsorption is one of the most effective and economic techniques. Additionally, adsorption can occur at ambient temperatures and low pressures, so it is relatively safe moreover in air purification also activated carbon is being used in controlling potential harmful, environmentally damaging and unpleasant odours. Moreover, carbon adsorption technology is very effective for many applications and the major attribute of activated carbon treatment is its ability to remove a wide variety of toxic organic compounds to non-detectable levels. Furthermore, in Malaysia, palm empty fruit bunch generates a large part of municipal waste and produce a problem of considerable concern for the environment. The use of this cheap, viable and readily carbon source for production of activated carbon for

subsequent use in the adsorption of acid gases will go a long way in promoting environmental pollution mitigation and waste management which helps in environmental sustainability and in turn curb the potential danger of health risks.

The costing for managing the uncontrollable waste from palm oil production can be reduced by using palm kernel shell and empty fruit bunch as the raw material for the production of activated carbon. The activated carbon produced will be very economical by converting low value materials (agricultural wastes) to high value products (activated carbon). Besides that, renewable material sources like biomass-waste have been recognized all around the world as a key driver to accomplish economic development while guaranteeing minimum environmental damage. At the same time, the present development of green innovation and its related strategies have improved the development of renewable material sources in the nation (Foo-Yuen *et al.*, 2011). This study is special application in gas mask respiratory system using agricultural wastes as precursors for production of activated carbon. This special product used for the gas poisoning adsorption which follows the standard requirement from NIOSH and OSHA. The application and utilisation on these products is commonplace and widespread throughout the industry activities involving contaminated air and hazardous environments. The product is essential and in fact mandatory as required by NIOSH for certain environments.

The most important is there is no research and report available for the adsorption of the selected gases ( $\text{Cl}_2$ ,  $\text{SO}_2$ ,  $\text{CO}$ , and  $\text{NH}_3$ ) on the palm empty fruit bunch activated carbon that impregnated with potassium hydroxide (KOH) especially in adsorption breakthrough study. From the adsorption breakthrough study, the time for the adsorbent material to completely saturate with the adsorbate will be obtained and how long for the adsorbent can be used before it will be replaced or regenerated will be known. This result will be applied in the gas mask respirator application and the results will be compared to the application using commercial activated carbon.

- Aksu, Z. and Gönen, F. (2004). Biosorption of phenol by immobilized activated sludge in a continuous packed bed: Prediction of breakthrough curves. *Process Biochemistry*, 39(5), 599-613.
- Allen, S. J., Koumanova, B., Kircheva, Z., Nenkova, S. (2005). Adsorption of 2-Nitrophenol by Technical Hydrolysis Lignin: Kinetics Mass Transfer and Equilibrium Studies. *Industrial Engineering Chemical Resource*, 44, 2281-2287.
- Alias, N., Ibrahim, N., Hamid, M. K. A., Hasbullah, H., Ali, R. R., Sadikin, A. N., Asli, U. A. (2014). Thermogravimetric analysis of rice husk and coconut pulp for potential biofuel production by flash pyrolysis. *The Malaysian Journal of Analytical Sciences*, 18, 705-710.
- Altenor, S., Carene, B., Emmanuel, E., Lambert, J., Ehrhardt, J. J., and Gaspard, S. (2009). Adsorption studies of methylene blue and phenol onto vetiver roots activated carbon prepared by chemical activation. *Journal of Hazardous Materials*, 165(1), 1029-1039.
- Amin, N. K. (2009). Removal of direct blue-106 dye from aqueous solution using new activated carbons developed from pomegranate peel: Adsorption equilibrium and kinetics. *Journal of Hazardous Materials*, 165(1), 52-62.
- Ania, C. O., Parra, J. B., Menendez, J. A., Pis, J. J. (2005). Effect of microwave and conventional regeneration on the microporous and mesoporous network and on the adsorptive capacity of activated carbons. *Microporous and Mesoporous Materials*, 85, 7-15.
- Arami-Niya, A., Daud, W. M. A. W., and Mjalli, F. S. (2010). Using granular activated carbon prepared from oil palm shell by  $ZnCl_2$  and physical activation for methane adsorption. *Journal of Analytical and Applied Pyrolysis*, 89(2), 197-203.
- Arami-Niya, A., Daud, W. M. A. W., Mjalli, F. S., Abnisa, F., and Shafeeyan, M. S. (2012). Production of microporous palm shell based activated carbon for methane adsorption: Modeling and optimization using response surface methodology. *Chemical Engineering Research and Design*, 90(6), 776-784.
- Aranovich, G. and Donohue, M. (1998). Analysis of Adsorption Isotherms: Lattice Theory Predictions, Classification of Isotherms for Gas-Solid Equilibria, and Similarities in Gas and Liquid Adsorption Behavior. *Journal of Colloid and Interface Science*, 200, 273-290.

- Aslaibi, T. M., Abustan, I., Ahmad, M. A., and Foul. A. A.(2013). A Review: Production of Activated Carbon from Agricultural byproducts via conventional and microwave heating. *Jurnal Chemical Technology Biotechnology*, 88, 1183-1190.
- Asenjo, N. G., Álvarez, P., Granda, M., Blanco, C., Santamaría, R., and Menéndez, R. (2011). High performance activated carbon for benzene/toluene adsorption from industrial wastewater. *Journal of Hazardous Materials*, 192(3), 1525-1532.
- Atkins, P. and de Paula, J. (2006), *Physical Chemistry for the Life Sciences*, New York: Oxford University Press.
- Auta, M. and Hameed, B. H. (2002). Modified mesoporous clay adsorbent for adsorption isotherm and kinetics of methylene blue. *Chemical Engineering Journal*, 198-199, 219-227.
- Auta, M. and Hameed, B. H. (2013). Coalesced chitosan activated carbon composite for batch and fixed-bed adsorption of cationic and anionic dyes. *Colloids and Surfaces B: Biointerfaces*, 105, 199-206.
- Ayoob, S., Gupta, A. K., and Bhakat, P. B. (2007). Analysis of breakthrough developments and modeling of fixed bed adsorption system for As (V) removal from water by modified calcined bauxite (MCB). *Separation and Purification Technology*, 52(3), 430-438.
- Bandosz, T. J. and Ania, C. O. (2006). *Activated Carbon Surfaces in Environmental Remediation*. Interface Science and Technology, New York U.S.A.: Elsevier.
- Bansal, R. C. and Goyal, M. (2005). *Activated Carbon Adsorption*. CRC Press.
- Bhatnagar, A. and Minocha, A. K. (2006). Conventional and non-conventional adsorbents for removal of pollutants from water-a review. *Indian Journal of Chemical Technology*, 13(3), 203-217.
- Bhuvaneshwari, S. and Sivasubramanian, V. (2014). Equilibrium, kinetics, and breakthrough studies for adsorption of Cr (vi) on chitosan. *Chemical Engineering Communications*, 201(6), 834-854.
- Bradshaw, S.M., van Wyk, E.J., Swardt, J. B. (1998). Microwave heating principles and the application to the regeneration of granular activated carbon. *Journal of The South African Institute of Mining and Metallurgy*, 4, 201-210.

- Buasri, A., Chaiyut, N., and Nakweang, C. (2011). Preparing activated carbon from palm shell for biodiesel fuel production. *Chiang Mai Journal of Science*, 38, 572-578.
- Byrne, J. F. and Marsh, H. (1995). Introductory Overview. In Patrick, J. W. (Ed.) *Porosity in Carbons*. Edward Arnold: London.
- Calvo-Muñoz, E. M., García-Mateos, F. J., Rosas, J. M., Rodríguez-Mirasol, J., and Cordero, T. (2002). Biomass Waste carbon Materials as adsorbents for CO<sub>2</sub> capture under Post-combustion conditions. *Frontiers in Material*, 3(23), 1-14.
- Cansado, I. P., Mourão, P. A., Falcão, A. I., Carrott, M. R., and Carrott, P. J. (2012). The influence of the activated carbon post-treatment on the phenolic compounds removal. *Fuel Processing Technology*, 103, 64-70.
- Carabineiro, S. A. C., Ramos, A. M., Vital, J., Loureiro, J. M., Orfao, J. J. M., Foneseca, I. M. (2003). Adsorption of SO<sub>2</sub> using vanadium and vanadium-copper supported on activated carbon. *Catalysis Today*, 78, 203–210.
- Chandra, T. C., Mirna, M. M., Sunarso, J., Sudaryanto, Y., and Ismadji, S. (2009). Activated carbon from durian shell: Preparation and characterization. *Journal of the Taiwan Institute of Chemical Engineers*, 40(4), 457-462.
- Chiang, H. L., Huang, C. P., Chiang, P. C., and You, J. H. (1999). Effect of metal additives on the physico-chemical characteristics of activated carbon exemplified by benzene and acetic acid adsorption. *Carbon*, 37(12), 1919-1928.
- China Activated Carbon Industry Report, 2014-2017. Available from: <http://www.rnrmarketresearch.com/china-activated-carbon-industry-report-2014-2017-market-report.html>
- Chingombe, P., Saha, B., and Wakeman, R. J. (2005). Surface modification and characterisation of a coal-based activated carbon. *Carbon*, 43(15), 3132-3143.
- Choo, H. S., Lau, L. C., Mohamed, A. R., and Lee, K. T. (2013). Hydrogen sulfide adsorption by alkaline impregnated coconut shell activated carbon. *Journal of Engineering Science and Technology*, 8, 741-753.
- Cuerda-Correa, E. M., Díaz-Díez, M. A., Macías-García, A., and Gañán-Gómez, J. (2006). Preparation of activated carbons previously treated with sulfuric acid:

- A study of their adsorption capacity in solution. *Applied Surface Science*, 252(17), 6042-6045.
- Czarneski, M. A. (2008). Facility Decontamination with Gaseous and Liquid Chlorine Dioxide. In Dunn, S. and Wood, J. P. (2008) *Report on the 2008 Workshop on Decontamination Research and Associated Issues for Sites Contaminated with Chemical, Biological, or Radiological Materials*. Environmental Protection Agency: Research Triangle Park, NC.
- Dandekar, A., Baker, R. T. K., and Vannice, M. A. (1998). Characterization of activated carbon, graphitized carbon fibers and synthetic diamond powder using TPD and DRIFTS. *Carbon*, 36(12), 1821-1831.
- Darmstadt, H. and Roy, C. (2003). Surface spectroscopic study of basic sites on carbon blacks. *Carbon*, 41, 2662–2665.
- Das, D., Samal, D. P., Meikap, B. C. (2015). Preparation of Activated Carbon from Green Coconut Shell and its Characterization. *Journal Chemical Engineering Process Technology*, 6, 5.
- Daud, W. M. A. W., and Ali, W. S. W. (2004) Comparison on pore development of activated carbon produced from palm shell and coconut shell. *Bioresource Technology*, 93, 63–69.
- Davini, P. (2001). The effect of certain metallic derivatives on the adsorption of sulphur dioxide on activated carbon. *Carbon*, 39, 419–424.
- Dechant, D. (2008). Expedited Fumigation of a Large Hospital as Related to Biological Contamination Scenarios. In Dunn, S. and Wood, J. P. (2008) *Report on the 2008 Workshop on Decontamination Research and Associated Issues for Sites Contaminated with Chemical, Biological, or Radiological Materials*. Environmental Protection Agency: Research Triangle Park, NC.
- Dehdashti, A., Khavanin, A., Rezaee, A., and Assilian, H. (2011). Regeneration of granular activated carbon saturated with gaseous toluene by microwave irradiation. *Turkish Journal of Engineering and Environmental Sciences*, 35(1), 49-58.
- Demirbas, A. (2004). Effects of temperature and particle size on bio-char yield from pyrolysis of agricultural residues. *Journal of Analytical and Applied Pyrolysis*, 72(2), 243-248.

- Demirbas, A. (2004). Adsorption of Lead and Cadmium Ions in Aqueous Solutions onto Modified Lignin from Alkali Glycerol Delignification. *Journal of Hazardous Material*, 109, 221-226.
- Demirbas, A. (2009). Agricultural based activated carbons for the removal of dyes from aqueous solutions: a review. *Journal of Hazardous Materials*, 167(1), 1-9.
- Deng, H., Zhang, G., Xu, X., Tao, G., Dai, J. (2010). Optimization of preparation of activated carbon from cotton stalk by microwave assisted phosphoric acid-chemical activation. *Journal of Hazardous Materials*, 182, 217–224.
- Diao, Y., Walawender, W. P., and Fan, L. T. (2002). Activated carbons prepared from phosphoric acid activation of grain sorghum. *Bioresource Technology*, 81(1), 45-52.
- Din, A. T. M., Hameed, B. H., and Ahmad, A. L. (2009). Batch adsorption of phenol onto physiochemical-activated coconut shell. *Journal of Hazardous Materials*, 161(2), 1522-1529.
- Duan, X., Peng, J., Srinivasakannan, C., Zhang, L., Xia, H., Yang, K., and Zhang, Z. (2011). Process optimization for the preparation of activated carbon from Jatropha hull using response surface methodology. *Energy Sources, Part A: Recovery, Utilization, and Environmental Effects*, 33(21), 2005-2017.
- Emsley, J. (1998). *The Elements*, (3rd ed.). Oxford: Clarendon Press, 120-121, 198-199.
- Ezeoha, S.L., Akubuo, C.O., Ani, A.O. (2012). Proposed Average Values of Some Engineering Properties of Palm Kernels. *Nigerian Journal of Technology (NIJOTECH)*, 31(2), 167-173.
- Fan, X., Li, C., Zeng, G., Gao, Z., Chen, L., Zhang, W., and Gao, H. (2010). Removal of gas-phase element mercury by activated carbon fiber impregnated with CeO<sub>2</sub>. *Energy Fuels*, 24(8), 4250-4254.
- Foo, K. Y. and Hameed, B. H. (2011). Preparation of oil palm (Elaeis) empty fruit bunch activated carbon by microwave-assisted KOH activation for the adsorption of methylene blue. *Desalination*, 275(1), 302-305.
- Foo, K. Y. and Hameed, B. H. (2012). Porous structure and adsorptive properties of pineapple peel based activated carbons prepared via microwave assisted KOH and K<sub>2</sub> CO<sub>3</sub> activation. *Microporous and Mesoporous Materials*, 148(1), 191-195.

- Foo, K.Y. and Hameed, B.H. (2012). Dynamic adsorption behavior of methylene blue onto oil palm shell granular activated carbon prepared by microwave heating. *Chemical Engineering Journal*, 203, 81–87.
- Foo-Yuen, N., Foong-Kheong, Y., Basiron, Y., and Sundram, K. (2011). A renewable future driven with Malaysian palm oil-based green technology. *Journal of Oil Palm and The Environment*, 2, 1-7.
- Garba, A., Basri, H., Nasri, N. S., and Isma'il, R. (2006). Synthesis and characterization of porous carbon from biomass using KOH and K<sub>2</sub>CO<sub>3</sub> chemical activation. *Journal of Engineering and Applied Sciences*, 11, 1613-1616.
- García, A. B., Martínez-Alonso, A., y Leon, C. A. L., and Tascón, J. M. (1998). Modification of the surface properties of an activated carbon by oxygen plasma treatment. *Fuel*, 77(6), 613-624.
- García, S., Gil, M. V., Martín, C. F., Pis, J. J., Rubiera, F., and Pevida, C. (2011). Breakthrough adsorption study of a commercial activated carbon for pre-combustion CO<sub>2</sub> capture. *Chemical Engineering Journal*, 171(2), 549-556.
- Gaur, V. and Shankar, P. A. (2008). Surface modification of activated carbon for the removal of water impurities. *Water Technology Online*.
- Gbadamosi, L. (2006). Some engineering properties of palm kernel seeds (PKS). *Journal Agricultural Engineering and Technology*, 14, 58-66.
- Ghorai, S. and Pant, K. K. (2005). Equilibrium, kinetics and breakthrough studies for adsorption of fluoride on activated alumina. *Separation and Purification Technology*, 42(3), 265-271.
- Gonçalves, M., Molina-Sabio, M., and Rodriguez-Reinoso, F. (2010). Modification of activated carbon hydrophobicity by pyrolysis of propene. *Journal of Analytical and Applied Pyrolysis*, 89(1), 17-21.
- Guo, J., and Lua, A. C. (2000). Effect of surface chemistry on gas-phase adsorption by activated carbon prepared from oil-palm stone with pre-impregnation. *Separation and Purification Technology*, 18, 47–55.
- Guo, J. and Lua, A. C. (2002). Characterization of adsorbent prepared from oil-palm shell by CO<sub>2</sub> activation for removal of gaseous pollutants. *Materials Letters*, 55(5), 334-339.



- Guo, J. and Lua, A. C. (2003). Adsorption of sulphur dioxide onto activated carbon prepared from oil-palm shells with and without pre-impregnation. *Separation and Purification Technology*, 30(3), 265-273.
- Guo, J., Xu, W. S., Chen, Y. L., and Lua, A. C. (2005). Adsorption of NH<sub>3</sub> onto activated carbon prepared from palm shells impregnated with H<sub>2</sub>SO<sub>4</sub>. *Journal of Colloid and Interface Science*, 281(2), 285-290.
- Gupta, A., Gaur, V., and Verma, N. (2004). Breakthrough analysis for adsorption of sulfur-dioxide over zeolites. *Chemical Engineering and Processing: Process Intensification*, 43(1), 9-22.
- Hamza D. U., Nasri, N. S., Abdul, K. A., Ahmed, M. M. and Mohammed, J. (2012). Modification of activated carbon for enhancement of gas contaminant removal: A review. In *UMT 11th International Annual Symposium on Sustainability Science and Management 9-11<sup>th</sup> July 2012*, Terengganu, Malaysia, 1336-1342.
- Hamza U. D., Nasri N. S., Mohammed J., Zain H. M. (2015). Thermal regeneration of CO<sub>2</sub> spent palm shell-polyetheretherketone activated carbon sorbents. *International Journal of Chemical, Nuclear, Materials and Metallurgical Engineering*, 9, 492-495.
- Haridass, C., Aw-Musse, A., Misra, P., and Jordan, J. (2000). Fourier transform infrared (FT-IR) spectroscopy of trace molecular species of importance for the elucidation of atmospheric phenomena. *Computers & Electrical Engineering*, 26(1), 47-65.
- Hariharan, A. B. A. and H. P. S. Abdul-Khalil. (2005). Lignocellulose based hybrid bi-layer laminate composites: Part 1-studies on tensile and impact behavior of oil palm fiber/glass fiber reinforced epoxy. *Journal of Composite Materials*. 39(8), 663-684.
- Haykiri-Acma, H., Yaman, S, and Kucukbayrak, S. (2010). Comparison of the thermal reactivities of isolated lignin and holocellulose during pyrolysis. *Fuel Process Technology*, 91, 759-64.
- Hesas, R. H., Arami-Niya, A., Daud, W. M. A. W., and Sahu, J. N. (2013). Preparation and characterization of activated carbon from apple waste by microwave-assisted phosphoric acid activation: Application in methylene blue adsorption. *Bio Resources*, 8(2), 2950-2966.

- Hesas, R. H., Arami-Niya, A., Daud, W. M. A. W., and Sahu, J. N. (2013). Preparation of granular activated carbon from oil palm shell by microwave-induced chemical activation: Optimisation using surface response methodology. *Chemical Engineering Research and Design*, 91(12), 2447-2456.
- Ho, C. K. and Webb, S. W. (2006). *Gas Transport in Porous Media*, Springer.
- Hsu, T. C. (2008). Adsorption of an acid dye onto coal fly ash. *Fuel*, 87(13), 3040-3045.
- Hunt, J., Ferrari, A., Lita, A., Crosswhite, M., Ashley, B., and Stiegman, A. E. (2013). Microwave-Specific Enhancement of the Carbon–Carbon Dioxide (Boudouard) Reaction. *The Journal of Physical Chemistry*, 117, 26871–26880.
- Hu, Z. and Srinivasan M. P. (1999). Preparation of high-surface-area activated carbon from coconut shell. *Microporous and Mesoporous Materials*, 27, 11-18.
- Huang, C. C., Li, H. S., and Chen, C. H. (2008). Effect of surface acidic oxides of activated carbon on adsorption of ammonia. *Journal of Hazardous Materials*, 159, 523-527.
- Iron (III) chloride 6-hydrate MSDS. Available from: <http://www.sciencelab.com/msds.php?msdsId=9924034>
- Islam, N., and Ghosh, D. C. (2011). The electronegativity and the global hardness are periodic properties of atoms. *Journal of Quantum Information Science*, 1, 135-141.
- Karagöz, S., Tay, T., Ucar, S., and Erdem, M. (2008). Activated carbons from waste biomass by sulfuric acid activation and their use on methylene blue adsorption. *Bioresource Technology*, 99(14), 6214-6222.
- Khalil, H. A., Fazita, M. N., Jawaid, M., Bhat, A. H., and Abdullah, C. K. (2010). Empty fruit bunches as a reinforcement in laminated bio-composites. *Journal of Composite Materials*, 45(2), 219-236.
- Kim, D. J. and Yie, J. E. (2005). Role of copper chloride on the surface of activated carbon in adsorption of methyl mercaptan. *Journal of Colloid and Interface Science*, 283(2), 311-315.

- Kowalczyk, Z., Sentek, J., Jodzis, S., Diduszko, R., Presz, A., Terzyk, A., and Suwalski, J. (1996). Thermally modified active carbon as a support for catalysts for NH<sub>3</sub> synthesis. *Carbon*, 34(3), 403-409.
- Lee, C. S., Ong, Y. L., Aroua, M. K., and Daud, W. W. (2013). Impregnation of palm shell-based activated carbon with sterically hindered amines for CO<sub>2</sub> adsorption. *Chemical Engineering Journal*, 219, 558-564.
- Lee, T., Ooi, C-H., Othman, R., and Yeoh, F-Y. (2014). Activated Carbon Fiber-The Hybrid of Carbon Fiber and Activated Carbon. *Review Advanced Material Science*, 36, 118-136.
- Lehman, J. and Joseph, S. (Eds.) (2009). *Biochar for Environmental Management*. Earthscan: London.
- Li, L., Liu, S., and Junxin, L. (2011). Surface modification of coconut shell based activated carbon for the improvement hydrophobic VOC removal. *Journal of Hazardous Mats*, 192, 683-690.
- Lim, K. O. (2000). Meeting Malaysia's future energy demand with energy plantations. *Proceedings of the World Renewable Energy Congress VI (Part III)*. Brighton, United Kingdom, 1-7<sup>th</sup> July, 1404-1407.
- Liou, T-H. (2010). Development of mesoporous structure and high adsorption capacity of biomass-based activated carbon by phosphoric acid and zinc chloride activation. *Chemical Engineering Journal*, 158, 129-142.
- Lu, X., Jiang, J., Sun, K, Xie, X., and Hu, Y. (2012). Surface modification, characterization and adsorptive properties of a coconut activated carbon. *Applied Surface Science*, 258, 8247-8252.
- Ludin, N. A., Bakri, M. A. M., Kamaruddin, N., Sopian, K., Deraman, M. S., Hamid, N. H., Asim, N., Othman, M. Y. (2014). Malaysian oil palm plantation sector: exploiting renewable energy toward sustainability production. *Journal of Cleaner Production*, 65, 9-15.
- Luftman, H. S., Regits, M. A., Lorcheim, P., Czarneski, M. A., Boyle, T., Aceto, H., Dallap, B., Munro, D., and Faylor, K. (2006). Chlorine dioxide gas decontamination of large animal hospital intensive and neonatal care units. *Appl. Biosafety*, 11, 144-154.
- Ma, J., Li, L., Ren, J., and Li, R. (2010). CO adsorption on activated carbon-supported Cu-based adsorbent prepared by a facile route. *Separation and Purification Technology*, 76, 89-93.

- Ma, S., Zhao, Y., Guo, T., Ma, X. (2009). Experimental Study on Microwave-Induced NO Decomposition on Activated Carbon Bed. *Power and Energy Engineering Conference, APPEEC Asia-Pacific*, 27-31 March, Wuhan, China.
- Malaysian Palm Oil Board (2010). Malaysian Palm Oil Statistics 2009. Available from: <http://www.mpob.gov.my/>
- Malaysian Palm Oil Board (2016). Malaysian Palm Oil Statistics 2015. Available from: <http://www.mpob.gov.my/>
- Malaysian Palm Oil Council (2012). MPOC 2012.
- Market Research Report on Global and Chinese Activated Carbon Industry, 2009-2019. Available from: <http://www.rnrmarketresearch.com/market-research-report-on-global-and-chinese-activated-carbon-industry-2009-2019-market-report.html>
- Matos, J., Nahas, C., Rojas, L., and Rosales, M.(2011). Synthesis and characterization of activated carbon from sawdust of Algarroba wood. 1. Physical activation and pyrolysis. *Journal of Hazardous Materials*, 196, 360-369.
- Menendez, J. A., Phillips, J., Xia, B., and Radovic, L. R. (1996). The modification and characterization of chemical surface-properties of activated carbon-in the search of carbons with stable basic properties. *Langmuir*, 12, 4404-4410.
- Mohamad, A. B., Iyuke, S. E., Daud, W. R. W., Kadhum, A. A. H., Faisal, Z., Al-Khatib, M. F., and Shariff, A. M. (2000). Adsorption of carbon monoxide on activated carbon–tin ligand. *Journal of Molecular Structure*, 550, 511-519.
- MSA the Safety Company (2012). *Gas Detection Handbook* (6<sup>th</sup> ed.). U.S.A.
- Mubarak, N. M., Kundu, A., Sahu, J. N., Abdullah, E. C., and Jayakumar, N. S. (2014). Synthesis of palm oil empty fruit bunch magnetic pyrolytic char impregnating with FeCl<sub>3</sub> by microwave heating technique. *Biomass and Bioenergy*, 61, 265-275.
- Nabais, J.M.V., Laginhas, C.E.C., Carrott, P.J.M., Carrott, R. (2011). Production of activated carbon from almond shell. *Fuel Proc. Tech.* 92, 234-240.
- Nagano, S., Tamon, H., Adzumi, T., Nakagawa, K., and Suzuki, T. (2000). Activated carbon from municipal waste. *Carbon*, 38(6), 915-920.
- Nasri, N. S., Hamza, U. D., Amin, S., Aishah, N., Ahmed, M. M., Mohammed, J., and Mohd Zain, H. (2014). Methane adsorption on chemically modified

- microwave irradiated palm shell porous carbon. *Advanced Materials Research*, 1043, 175-179.
- Nasri, N. S., Tatt, E. C., Hamza, U. D., Mohammed, J., and Zain, H. M. (2015). Kinetic rate comparison of methane catalytic combustion of palladium catalysts impregnated onto  $\gamma$ -alumina and bio-char. *International Scholarly and Scientific Research & Innovation*, 9(4), 565.
- National Institute for Occupational Safety (2000). *NIOSH Pocket Guide to Chemical Hazards*. DIANE Publishing.
- Ninduangdee, P., Kuprianov, V. I., Cha, E. Y., Kaewrath, R., Youngyuen, P., and Atthawethworawuth, W. (2015). Thermogravimetric studies of oil palm empty fruit bunch and palm kernel shell: TG/DTG analysis and modeling. *Energy Procedia*, 79, 453-458.
- Ning, P., Li, F., Yi, H., Tang, X., Peng, J., Li, Y., He, D., and Deng, H. (2012). Adsorption equilibrium of methane and carbon dioxide on microwave-activated carbon. *Separation and Purification Technology*, 98, 321-326.
- Nor, N. M., Chung, L. L., Teong, L. K., and Mohamed, A. R. (2013). Synthesis of activated carbon from lignocellulosic biomass and its applications in air pollution control-A review. *Journal of Environmental Chemical Engineering*, 1, 658-666.
- Norikazu, K., Yamada, H., Yajima, T., and Sugiyama, K. (2007). Surface properties of activated carbon treated by cold plasma heating. *Thin Solid Films*, 515(9), 4192-4196.
- O'Connor, T. P. and Mueller, J. (2001). Modelling competitive adsorption of chlorinated volatile organic compounds with Dubinin-Radushkevich equation. *Microporous Mesoporous Mater*, 46, 314.
- Okman, I., Karagöz, S., Tay, T., and Erdem, M. (2014). Activated carbons from grape seeds by chemical activation with potassium carbonate and potassium hydroxide. *Applied Surface Science*, 293, 138.
- Oliveira, L. C. A., Rios, R. V. R. A., Fabris, J. D., Gargc, V., Sapag, K., and Lago, R. M. (2002). Activated carbon iron oxide magnetic composites for the adsorption of contaminants in water, *Carbon* 40, 2177-2183.
- Omar, R., Idris, A., Yunus, R., Khalid, K., Isma, M. I. A. (2011). Characterization of empty fruit bunch for microwave-assisted pyrolysis. *Fuel*, 90, 1536-1544.

- Önal, Y. (2006). Kinetics of adsorption of dyes from aqueous solution using activated carbon prepared from waste apricot. *Journal of Hazardous Materials*, 137(3), 1719-1728.
- Onay, Ö., Beis, S. H., and Koçkar, Ö. M. (2001). Fast pyrolysis of rape seed in a well-swept fixed-bed reactor. *Journal of Analytical and Applied Pyrolysis*, 58, 995-1007.
- Onay, O. (2007). Influence of pyrolysis temperature and heating rate on the production of bio-oil and char from safflower seed by pyrolysis using a wellswept fixed-bed reactor. *Journal Fuel Proc.*, 88, 523-531.
- Pakseresht, S., Kazemeini, M., and Akbarnejad, M. M. (2002). Equilibrium isotherms for CO, CO<sub>2</sub>, CH<sub>4</sub> and C<sub>2</sub>H<sub>4</sub> on the 5A molecular sieve by a simple volumetric apparatus. *Separation and Purification Technology*, 28(1), 53-60.
- Park, S. J. and Jang, Y. S. (2002). Pore structure and surface properties of chemically modified activated carbons for adsorption mechanism and rate of Cr (VI). *Journal of Colloid and Interface Science*, 249(2), 458-463.
- Park, S. J. and Kim, B. J. (2004). Influence of oxygen plasma treatment on hydrogen chloride removal of activated carbon fibers. *Journal of Colloid and Interface Science*, 275(2), 590-595.
- Peljhan, S. and Kokalj, A. (2009). Adsorption of Chlorine on Cu (111): A density-functional theory study. *The Journal of Physical Chemistry C*, 113(32), 14363-14376.
- Potassium hydroxide MSDS. Available from:<http://www.sciencelab.com/msds.php?msdsId=9927230>
- Prauchner, M. J. and Rodríguez-Reinoso, F. (2012). Chemical versus physical activation of coconut shell: A comparative study. *Microporous and Mesoporous Materials*, 152, 163-171.
- Rafatullah, M., Ahmad, T., Ghazali, A., Sulaiman, O., Danish, M., and Hashim, R. (2013), Oil Palm Biomass as a Precursor of Activated Carbons: A Review. *Critical Reviews in Environmental Science and Technology*, 43, 1117–1161.
- Rahman, M. and Mohd Yusof, A. (2011). Preparation and modification of activated carbon from oil-palm shell and its adsorption capacity through speciation of chromium. *Research Journal of Chemistry and Environment*, 15(4), 49-51.

- Rezaee, A., Rangkooy, H. A., Khavanin, A., Jonidi-Jafari, A., Soltani, R. D. C., and Nili-Ahmadabadi, A. (2011). Adsorption properties and breakthrough model of formaldehyde on bone char. *International Journal of Environmental Science and Development*, 2(6), 423.
- Rodrigues, C. C., De Moraes, D., Da Nobrega, S. W., and Barboza, M. G. (2007). Ammonia adsorption in a fixed bed of activated carbon. *Bioresource Technology*, 98(4), 886-891.
- Rodriguez-Reinoso, F. and Molina-Sabio, M. (1998). Textural and chemical characterization of microporous carbons. *Advances in Colloid and Interface Science*, 76, 271-294.
- Rugayah, A. F., Astimar, A. A., and Norzita., N. (2014). Preparation and Characterization of Activated Carbon from Palm Kernel Shell by Physical Activation with Steam. *Journal of Oil Palm Research*, 6(3), 251-264.
- Safety Equipment Australia (SEA) (1997). *The Practical Use of Some Existing Models for Estimating Service Life of Gas Filter*.
- Savova, D., Apak, E., Ekinci, E., Yardim, F., Petrov, N., Budinova, T., ... and Minkova, V. (2001). Biomass conversion to carbon adsorbents and gas. *Biomass and Bioenergy*, 21(2), 133-142.
- Şensöz, S., Demiral, İ., and Gerçel, H. F. (2006). Olive bagasse (*Olea europea* L.) pyrolysis. *Bioresource Technology*, 97(3), 429-436.
- Shafeeyan, M. S., Daud, W. M. A. W., Houshmand, A., and Shamiri, A. (2010). A review on surface modification of activated carbon for carbon dioxide adsorption. *Journal of Analytical and Applied Pyrolysis*, 89(2), 143-151.
- Sharma, R. K., Wooten, J. B., Baliga, V. L., Lin, X., Chan, W. G., and Hajaligol, M. R. (2004). Characterization of chars from pyrolysis of lignin. *Fuel*, 83(11), 1469-1482.
- Shen, W., Li, Z., and Liu, Y. (2008). Surface chemical functional groups modification of porous carbon. *Recent Patents on Chemical Engineering*, 1(1), 27-40.
- Shiue, A., Den, W., Kang, Y. H., Hu, S. C., Jou, G. T., Lin, C. H., ... and Lin, S. I. (2011). Validation and application of adsorption breakthrough models for the chemical filters used in the make-up air unit (MAU) of a cleanroom. *Building and Environment*, 46(2), 468-477.

- Sirisha D., Smita A., Gandhi N., and Hasheena, M. (2015). Adsorptive removal of aqueous SO<sub>2</sub> by using Orange Peel Powder. *Indian Journal of Science*, 12, 39-51.
- Skodras, G., Diamantopoulou, I., Zabaniotou, A., Stavropoulos, G., and Sakellariopoulos, G. P. (2007). Enhanced mercury adsorption in activated carbons from biomass materials and waste tires. *Fuel Processing Technology*, 88(8), 749-758.
- Somy, A., Mehrnia, M. R., Amrei, H. D., Ghanizadeh, A., and Safari, M. (2009). Adsorption of carbon dioxide using impregnated activated carbon promoted by Zinc. *International Journal of Greenhouse Gas Control*, 3, 249 – 254.
- Suhas, P. J. M., Carrot, M. M. L., and Carrot, R. (2007). Lignin- From Natural Adsorbent to Activated Carbon-A Review. *Bioresource Technology*, 98, 2301-2312.
- Sukiran, M. A., Chin, C. M., and Abu Bakar, N. K. (2009). Bio-oils from pyrolysis of oil palm empty fruit bunches. *Journal Applied Science*, 6(5), 869-875.
- Sulfur dioxide, Carbon Monoxide, Chlorine and Ammonia SDS. Available from: <https://www.calgaz.com/pages/safety-data-sheet>
- Sumathi, S., Bhatia, S., Lee, K. T., and Mohamed, A. R. (2010). Adsorption isotherm models and properties of SO<sub>2</sub> and NO removal by palm shell activated carbon supported with cerium (Ce/PSAC). *Chemical Engineering Journal*, 162(1), 194-200.
- Sumathi, S., Bhatia, S., Lee, K. T., and Mohamed, A. R. (2010). Selection of best impregnated palm shell activated carbon (PSAC) for simultaneous removal of SO<sub>2</sub> and NO<sub>x</sub>. *Journal of Hazardous Materials*, 176(1), 1093-1096.
- Takeuchi, Y., Hino, M., Yoshimura, Y., Otowa, T., Izuhara, H., and Nojima, T. (1999). Removal of single component chlorinated hydrocarbon vapor by activated carbon of very high surface area. *Separation and Purification Technology*, 15(1), 79-90.
- Tan, I. A. W., Ahmad, A. L., and Hameed, B. H. (2008). Optimization of preparation conditions for activated carbons from coconut husk using response surface methodology. *Chemical Engineering Journal*, 137(3), 462-470.
- Tan, I. A. W., Ahmad, A. L., and Hameed, B. H. (2009). Adsorption isotherms, kinetics, thermodynamics and desorption studies of 2,4,6-trichlorophenol on



- oil palm empty fruit bunch-based activated carbon. *Journal of Hazardous Materials*, 164, 473-482.
- Tham, Y.J., Latif, P.A., Abdullah, A.M., Shamala-Devi, A., and Taufiq-Yap, Y.H. (2011). Performances of toluene removal by activated carbon derived from durian shell. *Bioresource Technology*. 102, 724-728.
- Thommes, M., Kaneko, K., Neimark, A. V., Olivier, J. P., Rodriguez-Reinoso, F., Rouquerol, J., and Sing, K. S. W. (2015). Physisorption of gases, with special reference to the evaluation of surface area and pore size distribution (IUPAC Technical Report). *Pure Applied Chemistry*, 3-10.
- Tsai, J. H., Jeng, F. T. and Chiang, H. L. (2001). Removal of H<sub>2</sub>S from exhaust gas by use of alkaline activated carbon. *Adsorption*, 7(4), 357-366.
- Tseng, R.-L., and Tseng, S.-K. (2005). Pore structure and adsorption performance of the KOH-activated carbons prepared from corncob. *Journal of Colloid and Interface Science*, 287, 428–437.
- Van Zutphen, J. and Wijbrans, R. A. (2007). The CO<sub>2</sub> and energy balance of Malaysian palm oil. *Current Status and Potential for Future Improvements*, Carbon Capital Solutions.
- Varol, M., Atimtay, A. T., Bay, B., and Olgun, H. (2010). Investigation of co-combustion characteristics of low quality lignite coals and biomass with thermogravimetric analysis. *Thermochimica Acta*, 510(1), 195-201.
- Vidic, R. D. and Siler, D. P. (2001). Vapor-phase elemental mercury adsorption by activated carbon impregnated with chloride and chelating agents. *Carbon*, 39(1), 3-14.
- Wang, T., Tan, S., and Liang, C. (2009). Preparation and characterization of activated carbon from wood via microwave-induced ZnCl<sub>2</sub> activation. *Carbon*, 47(7), 1880-1883.
- Wankhade, A. A. and Ganvir, V. N. (2013). Preparation of low cost activated carbon from tea waste using sulphuric acid as activating agent. *Int. Res. J. Environment Sci*, 2(4), 53-55.
- Wood, J. P., Ryan, S. P., Snyder, E. G., Serre, S. D., Touati, A., and Clayton, M. J. (2010). Adsorption of chlorine dioxide gas on activated carbons. *Journal of the Air & Waste Management Association*, 60(8), 898-906.

- Wu, F. C., Tseng, R. L., and Hu, C. C. (2005). Comparisons of pore properties and adsorption performance of KOH-activated and steam-activated carbons. *Microporous and Mesoporous Materials*, 80(1), 95-106.
- Xiao, Y., Wang, S., Wu, D., and Yuan, Q. (2008). Experimental and simulation study of hydrogen sulfide adsorption on impregnated activated carbon under anaerobic conditions. *Journal of Hazardous Materials*, 153(3), 1193-1200.
- Xin-hui, D., Srinivasakannan, C., Jin-hui, P., Li-bo, Z., and Zheng-yong, Zhang. (2011). Comparison of activated carbon prepared from *Jatropha* hull by conventional heating and microwave heating. *Biomass and Bioenergy*, 35, 3920-3926.
- Xue, Y., Lu, G., Guo, Y., Guo, Y., Wang, Y., and Zhang, Z. (2008). Effect of pretreatment method of activated carbon on the catalytic reduction of NO by carbon over CuO. *Applied Catalysis B: Environmental*, 79(3), 262-269.
- Yacob, A. R., Wahab, N., Suhaimi, N. H., and Mustajab, M. K. A. A. (2013). Microwave induced carbon from waste palm kernel shell activated by phosphoric acid. *International Journal of Engineering and Technology*, 5(2), 214.
- Yagmur, E., Ozmak, M., and Aktas, Z. (2008). A novel method for production of activated carbon from waste tea by chemical activation with microwave energy. *Fuel*, 87(15), 3278-3285.
- Yang, K., Peng, J., Srinivasakannan, C., Zhang, L., Xia, H., and Duan, X. (2010). Preparation of high surface area activated carbon from coconut shells using microwave heating. *Bioresource Technology*. 101: 6163-6169.
- Yang, T. and Lua, A. C. (2006). Textural and chemical properties of zinc chloride activated carbons prepared from pistachio-nut shells. *Materials Chemistry and Physics*, 100, 438-444.
- Yin, C. Y., Aroua, M. K., and Daud, W. M. A. W. (2007). Review of modifications of activated carbon for enhancing contaminant uptakes from aqueous solutions. *Separation and Purification Technology*, 52(3), 403-415.
- Yusof Basiron (2010). Palm oil - a success story in green technology. *Plenary paper at the ASM International Conference (ASMIC) 2010*.
- Yusufu, M. I., Ariaahu, C. C., Nkpa, N. N., and Igbabul, B. D. (2012). Chlorine adsorption kinetics of activated carbon from selected local raw materials. *Journal of Chemical Engineering and Materials Science*, 3(2), 23-29.

- Zhang, L., Mi, M., Li, B., and Dong, Y. (2013). Modification of Activated Carbon by Means of Microwave Heating and its Effects on the Pore Texture and Surface Chemistry. *Research Journal of Applied Sciences, Engineering and Technology*, 5(5), 1836-1840.
- Zhang, P., Wanko, H., and Ulrich, J. (2007). Adsorption of SO<sub>2</sub> on activated carbon for low gas concentrations. *Chemical Engineering & Technology*, 30(5), 635-641.
- Zhang, Z., Xu, M., Wang, H., and Li, Z. (2010). Enhancement of CO<sub>2</sub> adsorption on high surface area activated carbon modified by N<sub>2</sub>, H<sub>2</sub> and ammonia. *Chemical Engineering Journal*, 160(2), 571-577.
- Zhou, H. F. and Haynes, R. J. (2010). Sorption of heavy metals by inorganic and organic components of solid wastes: Significance to use of wastes as low-cost adsorbents and immobilizing agents. *Critical Reviews in Environmental Science and Technology*, 40, 909–977.
- Zinc nitrate hexahydrate MSDS. Available from:  
<https://www.elac.edu/academics/departments/chemistry/chemistrydocuments/docs/Z/zinc%20nitrate%20hexahydrate.pdf>

## PUBLICATIONS

- Zain, H. M.**, Nasri, N. S., Sidik, H. U., Abdurashed, A., Mohsin, R., Majid, Z. A., Rashid, N. M., and Sharer. Z. (2016). Sulphur Dioxide and Oxygen Adsorption Isotherm Breakthrough Time on Surface Porous Palm Shell Activated Carbon, *Chemical Engineering Transactions*, 56, Accepted.
- Nasri, N. S., **Zain, H. M.**, Hamza, U. D., Majid, Z. A., Sharer. Z., Sazali, N. A., and Anirman, N. L. (2015). CO<sub>2</sub> Adsorption-breakthrough Study on Activated Carbon derived from Renewable Oil Palm Empty Fruit Bunch, *Australian Journal of Basic and Applied Sciences*, 9(26) Special 2015, 67-71.
- Nasri, N. S., **Zain, H. M.**, Abdurashed, A., Sidik, H. U., and Rashid, N. M. (2017). Acidic Gases Adsorption-breakthrough Study on Adsorbent derived from Sustainable Solid Waste Materials, *Research Journal of Applied Sciences*, Submitted.
- Nasri, N. S., **Zain, H. M.**, Sidik, H. U., Abdurashed, A., and Rashid, N. M.(2017). Adsorption Isotherm Breakthrough Time of Acidic and Alkaline Gases on Treated Porous Synthesized KOH-FeCl<sub>3</sub>.6H<sub>2</sub>O Sustainable Agro-Based Material, *Chemical Engineering Transactions*, 61, Submitted.
- Nasri, N. S., **Zain, H. M.**, Mohammed, J., Hamza, U. D., and Ahmed, M. M. (2015). Sustainable recycle solid waste to synthetic renewable solid energy. *Applied Mechanics and Materials*, 705, 24-28.
- Nasri, N. S., Martel, H., Abbas, I. M. H. I., Sidik, H. U., **Zain, H. M.**, Abdurashed, A., Mohsin, R., Majid, Z. A., Rashid, N. M., Sharer, Z., and Garba, A.(2016). Comparative Study of Natural Gas Adsorption Isotherms on KOH and H<sub>3</sub>PO<sub>4</sub> Palm Kernel Shells Porous Activated Carbon, *Chemical Engineering Transactions*, Accepted.
- Nasri, N. S., Abbas, I. M. H. I., Martel, H., Abdurashed, A., **Zain, H. M.**, Sidik, H. U., Mohsin, R., Majid, Z. A., Rashid, N. M., Sharer, Z., and Garba, A.(2016). CO<sub>2</sub> Adsorption and Desorption on KOH, H<sub>3</sub>PO<sub>4</sub> and FeCl<sub>3</sub>.6H<sub>2</sub>O

Surface Irradiated Porous Activated Carbon Palm Shell Kernel, Chemical Engineering Transactions, Accepted.

- Nasri, N. S., Hamza, U. D., Amin, N. A. S., Ahmed, M. M, Mohammed, J., and **Zain, H. M.** (2014). Kinetics of CO<sub>2</sub> adsorption on microwave palm shell activated carbon. *Advanced Materials Research*, 1043, 224-228.
- Nasri, N. S., Mohammed, J., Zaini, M. A. A., Hamza, U. D., **Zain, H. M.**, and Ani, F. N. (2014). Characteristics of potassium acetate - activated coconut shell carbon. *Advanced Materials Research*, 1043, 193-197.
- Nasri, N. S., Ramlan, N. S., Hamza, U. D., Mohammed, J., Ahmed, M. M., and **Zain, H. M.** (2015). Enhancing sustainable recycle solid waste to porous activated carbon for methane uptake, *Applied Mechanics and Materials*, 705, 19-23.
- Nasri, N. S., Mohammed, J., Zaini, M. A. A., Hamza, U. D., **Zain, H. M.**, and Ani, F. N. (2014). Equilibrium and kinetic studies of benzene and toluene adsorption onto microwave irradiated - coconut shell activated carbon. *Advanced Materials Research*, 1043, 219-223.
- Nasri, N. S., Hamza, U. D., Amin, N. A. S., Ahmed, M. M, Mohammed, J., and **Zain, H. M.** (2014). Methane adsorption on chemically modified microwave irradiated palm shell porous carbon. *Advanced Materials Research*, 1043, 175-179.
- Hamza, U. D., Nasri, N. S., Amin, N. S., Mohammed, J., and **Zain, H. M.** (2015). Thermal regeneration of CO<sub>2</sub> spent palm shell-PEEK activated carbon Sorbent , ICEERET, Marrakech, Morocco, 8 April, 2015.
- Nasri, N. S., Musa, A. M., Noor, N. M., Mohammed, J., Hamza, U. D., and **Zain, H. M.** (2014),Hydrophobicity Characterization of Bio-Wax Derived from Taro Leaf for Surface Coating Applications, *Advanced Materials Research* Vol. 1043, p.184-188.