

**EFFECT OF LEACHING ON LIME STABILISED CONTAMINATED SOIL  
AND MINING WASTE**

**BABA BIN MUSTA**

**A thesis submitted in fulfilment of the requirement for  
the award of the Degree of Doctor of Philosophy**

**Faculty of Civil Engineering  
Universiti Teknologi Malaysia**

**MARCH, 2003**

## Universiti Teknologi Malaysia

**BORANG PENYESAHAN STATUS TESIS**

**JUDUL: EFFECT OF LEACHING ON LIME STABILISED CONTAMINATED SOIL AND MINING WASTE**

**SESI PENGAJIAN: 2002/2003**

Saya

**BABA BIN MUSTA**  
**(HURUF BESAR)**

mengaku membenarkan tesis (~~PSM/Sarjana~~/Doktor Falsafah)\* ini disimpan di Perpustakaan Universiti Teknologi Malaysia dengan syarat-syarat kegunaan seperti berikut:

1. Tesis adalah hakmilik Universiti Teknologi Malaysia.
2. Perpustakaan Universiti Teknologi Malaysia dibenarkan membuat salinan untuk tujuan pengajian sahaja.
3. Perpustakaan dibenarkan membuat salinan tesis ini sebagai bahan pertukaran antara institusi pengajian tinggi.
4. \*\*Sila tandakan (✓)

SULIT

(Mengandungi maklumat yang berdarjah keselamatan atau kepentingan Malaysia seperti termaktub di dalam AKTA RAHSIA RESMI 1972)

TERHAD

(Mengandungi maklumat TERHAD yang telah ditentukan oleh organisasi/badan di mana penyelidikan dijalankan)

TIDAK TERHAD

(TANDATANGAN PENULIS)

(TANDATANGAN PENYELIA)

**Alamat Tetap:**  
**A39-29, Lorong Indah Permai 8**

**Taman Indah Permai**

**Off Jalan Sepanggar**

**Menggatal Kota Kinabalu**

**SABAH.**

**PROF. MADYA DR. KHAIRUL ANUAR B. KASSIM**

Nama Penyelia

Tarikh: **02-05-2003**

Tarikh: **02/05/2003**

CATATAN:

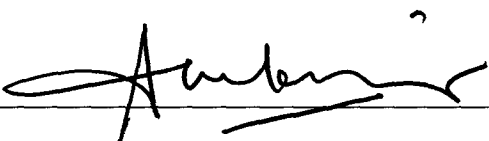
\*  
\*\*

Potong yang tidak berkenaan.

Jika tesis ini SULIT atau TERHAD, sila lampirkan surat daripada pihak berkuasa/organisasi berkenaan dengan menyatakan sekali sebab dan tempoh tesis ini perlu dikelaskan sebagai SULIT atau TERHAD.

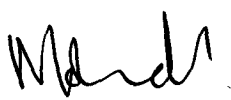
Tesis dimaksudkan sebagai tesis bagi Ijazah Doktor Falsafah dan Sarjana secara penyelidikan, atau disertasi bagi pengajian secara kerja kursus dan penyelidikan, atau Laporan Projek Sarjana Muda (PSM)

We certify that the work undertaken by the candidate has been carried out under our supervision.

Signature:  \_\_\_\_\_

Name of supervisor: Assoc. Prof. Dr. Khairul Anuar bin Kassim

Date: 02/05/2003

Signature:  \_\_\_\_\_

Name of supervisor: Assoc. Prof. Dr. Mohd. Razman bin Salim

Date: 02-05-2003

## BAHAGIAN A – Pengesahan Kerjasama\*

Adalah disahkan bahawa projek penyelidikan tesis ini telah dilaksanakan melalui kerjasama antara \_\_\_\_\_ dengan \_\_\_\_\_

Disahkan oleh:

Tandatangan : \_\_\_\_\_ Tarikh : \_\_\_\_\_

Nama : \_\_\_\_\_

Jawatan : \_\_\_\_\_  
(Cop rasmi)

\* Jika penyediaan tesis/projek melibatkan kerjasama.

---

## BAHAGIAN B – Untuk Kegunaan Pejabat Sekolah Pengajian Siswazah

Tesis ini telah diperiksa dan diakui oleh:

Nama dan Alamat  
Pemeriksa luar : **Prof. Madya Dr. Mohd Raihan bin Taha**  
**Timbalan Pengarah**  
**LESTARI**  
**Universiti Kebangsaan Malassia**  
**43600 Bangi**  
**Selangor**

Nama dan Alamat  
Pemeriksa Dalam I : **Prof. Madya Dr. Norhan bin Abd. Rahman**  
**Fakulti Kejuruteraan Awam**  
**UTM, Skudai, Johor**

Nama Penyelia Lain  
(Jika ada) : **Prof. Madya Dr. Mohd. Razman bin Salim**  
**Fakulti Kejuruteraan Awam**  
**UTM, Skudai, Johor**

Disahkan oleh Timbalan Pendaftar di SPS:

Tandatangan : \_\_\_\_\_

Tarikh: 10/9/03

Nama : HJ. DAHARI BIN HJ. DERANI

"I declare that this thesis is the result of my research except for cited references. The thesis has not been accepted for any degree and is not concurrently submitted in candidate of any other degree"

Signature



Name : Baba Bin Musta

Date : March, 2003

*Dedication to my beloved wife Suraya Sintang  
My children Hilal Hakimy, Bahrul Ilmy and Ain Nazifah  
To my parents, sisters and brother*

## ACKNOWLEDGEMENTS.

I would like to express my gratitude to my supervisor Assoc. Prof. Dr. Khairul Anuar Kassim for his supervision, advice, guidance, and useful comments throughout the completion of this thesis. I also wish to express my appreciation to my co-supervisor Assoc. Prof. Dr. Mohd. Razman Salim for his guidance and relevant suggestions. I am thankful to University Malaysia Sabah (UMS), for granting me the study leave and scholarship during my study. This research is also funded by the Intensification Research of Priority Area Coded 72277 lead by Assoc. Prof. Dr. Khairul Anuar Kassim and short term Research Grant (UTM) Coded 71574 lead by Assoc. Prof Z'aba Ismail. Acknowledgement to Kota Kinabalu City Hall, Sabah for allowing the writer to collect soil samples near the Kayumadang, Landfill, Telipok and also to Mamut Copper Mine Sdn. Bhd for allowing collection of mining waste samples in Lohan Valley, Ranau. Thanks to the technicians in the Geotechnical Laboratory Faculty of Civil Engineering, Universiti Teknologi Malaysia (UTM), Skudai for their assistance. Also to my research clique Mr. Kamaruddin Ahmad and Dr. Ayob Katimon for sharing ideas during the completion of this thesis. I would like acknowledge Mr Jerfi and Mr. Ayub from the Faculty of Mechanical Engineering, UTM for the SEM analyses and also to Mr. Abd. Aziz and Mr. Abd Hamid from the Program of Geology, Universiti Kebangsaan Malaysia (UKM) for helping with the XRF analyses. Thanks to Mr. Zainal Abidin from Faculty of Mechanical Engineering, UTM and Mr. Azhari from the Geology Programme, UKM for providing the XRD analyses, not forgetting Mr. Zaimi from Petronas Reseach Scientific Services, Bangi, and Mr. Mohd. Nazri from the Faculty of Science, UTM for helping with the SSA analyses. A lot of thanks to Mr. Jalaludin and Mr. Muhin from School of Science and Technology, UMS for helping with the soil collection and AAS analysis. Thanks to Mrs. Junedah from the Minerals and Geoscience Department, Kota Kinabalu for AAS analyses.

## ABSTRACT

Degradation of environment due to industrial and mining related activities is identified as one of the major problem facing by most industrial and developing countries. Therefore, scientific research are needed to acquire knowledge to manage such toxic waste, which has a potential to degrade on the environment as well as on human health. Leaching column tests were conducted on two soil samples and one mining waste sample in order to investigate the potential of lime stabilisation technique to immobilise heavy metals. The soil samples were obtained from Kg. Kayumadang, Telipok, Sabah (KMT) and Kg. Bongkud, Ranau, Sabah (KGB), whereas the mining waste was obtained from Lohan Dam, Ranau, Sabah (LDA). The amounts of lime used for the stabilisation were 2%, 4%, 6%, and 8% of dry weight. Five selected heavy metals namely Cu, Cr, Ni, Pb, and Zn were chosen to understand their behavioural changes with the long term leaching process. These heavy metals were spiked into the unstabilised and lime stabilised samples in the leaching columns. The pH values of leachate from all stabilised KMT, KGB, and LDA samples were found to be alkaline, indicating that the heavy metals were adsorbed on the surface of the cementitious minerals, whilst the Ca ions were released from the cementitious compound. The trend of migration profiles showed that the concentrations of the heavy metals decreased with depth for all samples. On the other hand, very low concentrations of heavy metals were detected in the leachate of the lime stabilised samples. Mass balance calculations showed that Zn had the highest mobility in KMT sample, whereas Ni was the highest in KGB and LDA samples. Calculations also showed that Pb was the least mobile in KMT, whereas Cr was the slowest in both KGB and LDA samples. The permeability of all stabilised samples decreased significantly immediately after leaching processes due to the clogging of pore by fine materials and the development of cementitious mineral. Scanning electron microscope (SEM) study showed that the bottom part of the leached columns had tight structure with low pore spaces, which were clogged-up by fine materials. Using X-ray diffraction (XRD) technique, quartz, kaolinite, feldspar and cementitious minerals were detected in the stabilised samples. Nevertheless, SEM method illustrated better image of the cementitious mineral. Soil classification tests, chemical characterisations, mineralogical identifications, and engineering properties verification were also performed on both unstabilised and stabilised samples to examine the liming effect on their properties. Maximum unconfined compressive strength (UCS) was achieved by adding 6% lime to the KMT, KGB, and LDA samples, and this strength increased with age. Based on the leaching column and UCS tests, 6% lime is suggested as the optimum value for stabilisation of KMT and KGB samples. Hence, lime stabilised KMT and KGB could be used as a clay liner due to their low permeability and high strength. It is also possible to stabilise LDA with 6% lime with the addition of fine material to facilitate stabilisation process, thus immobilising heavy metals within the stabilised mining waste.



## ABSTRAK

Penurunan kualiti alam sekitar akibat aktiviti-aktiviti berkaitan perindustrian dan perlombongan menjadi masalah utama yang dihadapi oleh kebanyakan negara industri dan negara membangun. Oleh yang demikian kajian saintifik adalah perlu untuk memperolehi pengetahuan bagi mengurus bahan buangan bertoksik yang berupaya menurunkan kualiti alam sekitar dan kesihatan manusia. Ujian kolum larutlesapan telah dijalankan ke atas dua sampel tanah dan satu sampel sisa lombong bertujuan mengkaji keupayaan kaedah penstabilan kapur untuk mentakmobilkan logam-logam berat. Sampel-sampel tanah diperolehi daripada Kg. Kayumadang, Telipok, Sabah (KMT) dan Kg. Bongkud, Ranau, Sabah (KGB), manakala sisa lombong diperolehi daripada Empangan Lohan, Ranau, Sabah (LDA). Peratusan kapur yang digunakan untuk penstabilan adalah 2%, 4%, 6%, dan 8% dengan berat kering sampel. Lima logam berat terpilih iaitu Cu, Cr, Ni, Pb, and Zn telah dipilih untuk kajian perubahan perlakuannya dengan proses larutlesapan jangka panjang. Logam-logam berat ini digaul kedalam sampel-sampel tidak distabilkan dan distabilkan dalam kolum larutlesapan. Nilai pH cecair larutlesapan daripada kesemua sampel-sampel KMT, KGB, dan LDA yang distabilkan didapati beralkali, menandakan logam-logam berat telah dijerap di permukaan mineral bersimen, manakala ion Ca telah dilepaskan daripada sebatian bersimen. Kajian tren profil-profil migrasi menunjukkan nilai kepekatan logam-logam berat berkurangan dengan kedalaman pada semua sampel. Sebaliknya, logam-logam berat dalam cecair larutlesapan daripada tanah yang distabilkan didapati amat rendah. Kiraan keseimbangan jisim menunjukkan Zn mempunyai mobiliti yang tinggi dalam sampel KMT, manakala Ni adalah tertinggi dalam sampel-sampel KGB dan LDA. Kiraan juga menunjukkan bahawa Pb mempunyai mobiliti terendah dalam sampel KMT, manakala Cr adalah terendah dalam sampel-sampel KGB dan LDA. Kebolehtelapan kesemua sampel yang distabilkan berkurang secara mendadak selepas proses larutlesapan di sebabkan oleh penyumbatan bahan-bahan halus dan pembentukan mineral bersimen. Kajian mikroskop elektron pengimbas (SEM) menunjukkan struktur padat dengan ruang pori yang rendah dan penyumbatan bahan-bahan halus pada bahagian bawah kolum. Kuarza, kaolinit, feldspar dan mineral bersimen dalam sampel-sampel yang telah distabilkan dikesan dengan menggunakan pembelauan sinar-X (XRD). Walaubagaimanapun, mineral bersimen dilihat lebih jelas dalam SEM. Ujian-ujian pengelasan tanah, pencirian kimia, pengenalan mineralogi, dan penentuan ciri-ciri kejuruteraan juga dijalankan keatas sampel-sampel yang tidak distabilkan dan yang distabilkan. Sampel-sampel tanah dan sisa lombong yang ditambah dengan 6% kapur menyebabkan kekuatan mampatan takterkurung (UCS) yang maksima dicapai dan ia juga bertambah dengan masa. Berdasarkan ujian-ujian kolum larutlesapan dan UCS di atas, 6% kapur dicadangkan sebagai nilai optimum untuk penstabilan sampel KMT dan KGB. Penstabilan keatas KMT dan KGB dicadangkan sebagai pelapik lempung disebabkan kebolehtelapan yang rendah dan kekuatan yang tinggi. Penstabilan LDA dengan 6% kapur di samping tambahan bahan halus juga boleh dilakukan bagi mempermudah proses penstabilan, seterusnya mentakgerakan logam berat di dalam sisa lombong yang distabilkan.

## TABLE OF CONTENTS

<b>CHAPTER</b>	<b>CONTENTS</b>	<b>PAGE</b>
	<b>TITLE</b>	i
	<b>DECLARATION</b>	ii
	<b>DEDICATION</b>	iii
	<b>ACKNOWLEDGEMENTS</b>	iv
	<b>ABSTRACT</b>	v
	<b>ABSTRAK</b>	vi
	<b>TABLE OF CONTENTS</b>	vii
	<b>LIST OF TABLES</b>	xvi
	<b>LIST OF FIGURES</b>	xviii
	<b>LIST OF NOTATIONS</b>	xxvi
	<b>LIST OF APPENDICES</b>	xxviii
<b>CHAPTER I</b>	<b>INTRODUCTION</b>	
	1.1 Related Environmental Issues	1
	1.2 Study Area	4
	1.3 Statement of Problem	5

1.4	Objectives	6
1.5	Scope of the Study	7
1.6	Significance of Research	7
1.7	Thesis Organisation	8

## **CHAPTER II      LITERATURE REVIEW**

2.1	Introduction	10
2.2	Sources of Heavy Metals in Soil	10
2.2.1	Mineralisation and Weathering	11
2.2.2	Mining Activity	13
2.2.3	Landfill Site	15
2.3	Contamination and Toxicity of Heavy Metals	17
2.4	Soil Properties and Constituent for Heavy Metals Availability	19
2.4.1	Soil pH	19
2.4.2	Organic Matter	20
2.4.3	Clay Minerals	20
2.4.4	Silica Minerals	21
2.4.5	Oxides of Iron and Aluminium	21
2.5	Behaviour of Heavy Metals in Soil	22
2.5.1	Cation Exchange	22
2.5.2	Adsorption	23
2.5.3	Precipitation	26
2.5.4	Complexation	29

2.6	Remediation Techniques	29
2.6.1	Physical Process	31
2.6.2	Biological Process	31
2.6.3	Chemical Process	32
2.6.4	Thermal Process	32
2.6.5	Solidification/Stabilisation	32
2.7	Lime Stabilisation	33
2.7.1	Chemical and Mineralogical Properties	35
2.7.2	Classification Characteristics	36
2.7.3	Leaching Test and Permeability	38
2.7.4	Strength Development	39
2.8	Heavy Metals Stabilisation Cases	41
2.8.1	Landfill Area	41
2.8.2	Mining Waste Area	42
2.8.3	Contaminated Area	43
2.9	Summary	43

### **CHAPTER III TEST PROCEDURES AND MATERIALS CHARACTERISTICS**

3.1	Introduction	45
3.2	Field Work and Sampling	46
3.3	Materials	46
3.3.1	KMT Soil	49

3.3.2	KGB Soil	49
3.3.3	Mining Waste (LDA)	50
3.3.4	Fine Material (FM)	50
3.3.5	Hydrated Lime	50
3.4	Moisture Content	51
3.5	Soil Classification Tests	51
3.5.1	Atterberg Limits	52
3.5.2	Shrinkage Limit	54
3.5.3	Specific Gravity	55
3.5.4	Particle Size Analysis	55
3.6	Chemical Tests	58
3.6.1	pH Value	58
3.6.2	Organic Matter	59
3.6.3	Specific Surface Area (SSA)	60
3.6.4	Cation Exchange Capacity (CEC)	60
3.6.5	Available Lime Content	61
3.6.6	X-Ray Fluorescence (XRF)	61
3.6.7	Atomic Absorption Spectrometer (AAS)	62
3.6.8	Sulphate Content	64
3.7	Mineralogical Identification	64
3.7.1	X-Ray Diffraction (XRD)	64
3.7.2	Scanning Electron Microscope (SEM)	65
3.8	Engineering Properties	69
3.8.1	Compaction Test	69

3.8.2	Unconfined Compression Test	70
3.8.3	Permeability Test	71
3.9	Summary	72

**CHAPTER IV**                    **EFFECT OF LIME ON THE PHYSICO-CHEMICAL  
AND ENGINEERING PROPERTIES OF SOILS AND  
MINING WASTE**

4.1	Introduction	74
4.2	Sample Preparation	75
4.3	Chemical Properties	75
4.3.1	pH	75
4.3.2	Specific Surface Area (SSA)	77
4.3.3	Cation Exchange Capacity (CEC)	78
4.3.4	X-Ray Fluorescence (XRF)	80
4.4	Physical Properties	81
4.4.1	Atterberg Limits	83
4.4.2	Specific Gravity	85
4.4.3	Particle Size Analysis	87
4.5	Engineering Properties	89
4.5.1	Compaction Test	89
4.5.2	Unconfined Compressive Test	92
4.6	Summary	97

**CHAPTER V                    MINERALOGICAL AND MICROSTRUCTURAL  
PROPERTIES ON LIME STABILISED SOILS AND  
MINING WASTE**

5.1	Introduction	99
5.2	Mineralogical Identifications study	99
5.2.1	X-ray Diffraction (XRD)	100
5.2.2	Scanning Electron Microscope (SEM)	106
5.2.2.1	Microstructural Study on Stabilised KMT Samples	106
5.2.2.2	Microstructural Study on Stabilised KGB Samples	110
5.2.2.3	Microstructural Study on Stabilised LDA samples	112
5.3	Summary	116

**CHAPTER VI                    LEACHING COLUM TESTS ON LIME STABILISED  
SOIL AND MINING WASTE**

6.1	Introduction	118
6.2	Leaching Cell Apparatus	119
6.3	Sample Preparation	119
6.4	Migration Profiles and Breakthrough Curve Results	124
6.4.1	Migration Profiles and Breakthrough Curve of KMT Samples	125
6.4.1.1	pH in KMT Samples	125
6.4.1.2	Calcium in KMT Samples	127
6.4.1.3	Copper in KMT Samples	129

6.4.1.4	Chromium in KMT Samples	131
6.4.1.5	Nickel in KMT Samples	133
6.4.1.6	Lead in KMT Samples	135
6.4.1.7	Zinc in KMT Samples	138
6.4.2	Migration Profiles and Breakthrough Curve of KGB Samples	140
6.4.2.1	pH in KGB Samples	140
6.4.2.2	Calcium in KGB Samples	142
6.4.2.3	Copper in KGB Samples	142
6.4.2.4	Chromium in KGB Samples	145
6.4.2.5	Nickel in KGB Samples	147
6.4.2.6	Lead in KGB Samples	149
6.4.2.7	Zinc in KGB Samples	151
6.4.3	Migration Profiles and Breakthrough Curve of LDA Samples	153
6.4.3.1	pH in LDA Samples	153
6.4.3.2	Calcium in LDA Samples	155
6.4.3.3	Copper in LDA Samples	157
6.4.3.4	Chromium in LDA Samples	159
6.4.3.5	Nickel in LDA Samples	161
6.4.3.6	Lead in LDA Samples	163
6.4.3.7	Zinc in LDA Samples	165
6.5	Mass Balance	167
6.5.1	Mass Balance of Heavy Metals in KMT Samples	167



6.5.2	Mass Balance of Heavy Metals in KGB Samples	169
6.5.3	Mass Balance of Heavy Metals in LDA Samples	171
6.6	Permeability Studies	173
6.6.1	Permeability Studies on KMT Samples	174
6.6.2	Permeability Studies on KGB Samples	175
6.6.3	Permeability Studies on LDA Samples	177
6.7	Microstructural Studies	178
6.7.1	Microstructural Studies on KMT Samples	178
6.7.2	Microstructural Studies on KGB Samples	182
6.7.3	Microstructural Studies on LDA Samples	185
6.8	Summary	189

## **CHAPTER VII      GENERAL DISCUSSION**

7.1	Introduction	191
7.2	Physical, Chemical and Engineering Properties of Materials	192
7.3	Physical, Chemical and Engineering Properties of Stabilised Soils and Mining Waste	198
7.4	Mineralogical and Microstructural Studies on Stabilised Soils and Mining Waste	204

7.5	Leaching Column Study on Stabilised Soils and Mining Waste	207
7.6	Design and Application of Stabilised Clay Liner in Landfill System	216

## **CHAPTER VIII CONCLUSIONS AND RECOMMENDATIONS**

8.1	Introduction	221
8.2	Conclusions	221
8.2.1	Materials Characteristics	221
8.2.2	Properties of Lime Stabilised Sample	223
8.2.3	Leaching Column Tests	225
8.3	Recommendations	226

<b>REFERENCES</b>	<b>229</b>
-------------------	------------

<b>APPENDICES</b>	<b>251</b>
-------------------	------------

**LIST OF TABLES**

<b>TABLE NO</b>		<b>Page</b>
2.1	Summary of the remediation techniques (Mulligan <i>et al.</i> , 2000).	30
3.1	Moisture content of materials.	51
3.2	Classification tests of materials.	54
3.3	Chemical properties of materials.	59
3.4	Composition of major elements and heavy metals in soil samples from the study area in Sabah and in the hydrated lime.	63
3.5	The summary of mineralogical content of material from the study area.	66
3.6	Engineering properties of materials.	70
4.1	pH, specific surface area (SSA) and, cation exchange capacity (CEC) of stabilised KMT, KGB and, LDA samples.	76
4.2	Cation exchange capacity (CEC) in unstabilised and stabilised KMT, KGB, and LDA samples.	79
4.3	Composition of major elements in lime stabilised KMT, KGB, and LDA samples with 2% and 6% of lime.	80
4.4	The physical characteristics of lime stabilised KMT samples.	81
4.5	The physical characteristics of lime stabilised KGB samples.	82
4.6	The physical characteristics of lime stabilised LDA samples.	82

5.1	Summary of mineralogical content in stabilised KMT, KGB and LDA samples from study area.	100
6.1	The percentages of soil, lime, water content and dry density of KMT samples for leaching tests preparation.	121
6.2	The percentages of soil, lime, water content and dry density of KGB samples for leaching tests preparation.	122
6.3	The percentages of soil, fine material, lime, water content and dry density of LDA samples for leaching tests preparation.	122
6.4	Mass Balance Calculations of KMT samples after 7PV of leaching.	168
6.5	Mass Balance Calculations of KGB samples after 7PV of leaching.	170
6.6	Mass Balance Calculations of LDA samples after 7PV of leaching.	172
7.1	The permeability, percentages of clay, percentages of coarse sand/gravel and plasticity index.	197
7.2	Metal fixing capacities of natural materials (Benson, 1980)	215

## LIST OF FIGURES

FIGURE NO		PAGE
2.1	Schematic diagram showing specifically and non-specifically adsorbed cations in diffuse double layer (DDL) model (Modified from Yong, 2000).	25
2.2	Schematic diagram showing the inner Helmholtz plane (IHP) and outer Helmholtz plane (OHP) on the clay particle (Yong, 2000).	26
2.3	Solubility-precipitation diagram for a metal hydroxide complex (Yong, 2000).	28
3.1	Map of the study area showing the location of KMT samples near the Kayumadang Landfill area, Telipok, Sabah.	47
3.2	Map of the study area showing the sampling station of KGB samples and LDA samples at Lohan Valley, Ranau, Sabah. (Modified from Colenette, 1958)	48
3.3	Plot of plasticities of KMT, KGB, and LDA and FM samples in the plasticity chart.	53
3.4	The particle size distribution of KMT samples.	56
3.5	The particle size distribution of KGB samples.	56
3.6	The particle size distribution of LDA samples.	57
3.7	The particle size distribution of FM samples.	57
3.8	The Scanning electron photomicrographs of (A) KMT show the kaolinite layered and (B) KMT shows sub-angular shape of quartz.	67
3.9	The Scanning electron photomicrographs of (A) KGB shows the kaolinite layered and (B) KGB shows sub-angular shape of quartz.	68

3.10	The Scanning electron photomicrographs of LDA showing angular shape of quartz.	69
4.1	Plasticity chart of stabilised KMT samples.	84
4.2	Plasticity chart of stabilised KGB samples.	86
4.3	Particle size distribution of stabilised KMT samples.	88
4.4	Particle size distribution of stabilised KGB samples.	88
4.5	Particle size distribution of stabilised LDA samples.	89
4.6	Compaction curve on lime stabilised KMT samples.	90
4.7	Compaction curve on lime stabilised KGB samples.	91
4.8	Compaction curve on lime stabilised LDA samples.	91
4.9	The compression strength of stabilised KMT samples with age.	94
4.10	The compression strength of stabilised KMT samples with different percentage of lime.	94
4.11	The compression strength of stabilised KGB samples with age.	95
4.12	The compression strength of stabilised KGB samples with different percentage of lime.	95
4.13	The compression strength of stabilised LDA samples with age.	96
4.14	The compression strength of stabilised LDA samples with different percentage of lime.	97
5.1	The X-ray diffractograms of (a) untreated KMT samples, (b) treated with 2% lime, and (c) treated with 6% of lime.	102
5.2	The X-ray diffractograms of (a) untreated KGB samples, (b) treated with 2%, and (c) treated with 6% lime.	103
5.3	(A) The X-ray diffractograms of (i) untreated LDA samples, (ii) treated with 2% lime, (iii) treated with 6% lime, (iv) treated with 6% lime and 10% fine material.	104
5.3	(B) The X-ray diffractograms of (i) untreated LDA samples, with 8% lime without fine material, (ii) treated with 8% lime with 10% fine material, and (iii) treated with 8% of lime, with 20% fine material.	105

5.4	Photomicrograph of KMT2% samples.	106
5.5	Photomicrograph of KMT4% samples.	107
5.6	(A) Photomicrograph of KMT6% samples (Cured for 1 day).	107
5.6	(B) Photomicrograph of KMT6% samples (Cured for 14 days).	108
5.6	(C) Photomicrograph of KMT6% samples (Cured for 45 days).	108
5.6	(D) Photomicrograph of KMT6% samples (Cured for 100 days).	109
5.7	Photomicrograph of KMT8% samples.	109
5.8	(A) Photomicrograph of KGB2% samples (Cured for 1 day).	110
5.8	(B) Photomicrograph of KGB2% samples (Cured for 28 days).	110
5.9	Photomicrograph of KGB4% samples.	111
5.10	Photomicrograph of KGB6% samples.	111
5.11	Photomicrograph of KGB8% samples.	112
5.12	Photomicrograph of LDA2% samples without FM.	113
5.13	Photomicrograph of LDA8% samples with 10%FM.	113
5.14	Photomicrograph of LDA2% samples with 20%FM. (Cured for 28 days).	114
5.15	Photomicrograph of LDA4% samples with 20%FM. (Cured for 28 days).	114
5.16	(A) Photomicrograph of LDA6% samples with 20%FM. (Cured for 28 days).	115
5.16	(B) Photomicrograph of LDA6% samples with 20%FM. (Cured for 45 days).	115
5.17	Photomicrograph of LDA8% samples with 20%FM.	116
6.1	The schematic of laboratory leaching cell.	120
6.2A	The pH value of unstabilised and stabilised KMT samples.	126

6.2B	The pH value of leachate from unstabilised and stabilised KMT samples.	126
6.3A	The migration profiles of Ca in unstabilised and stabilised KMT samples.	128
6.3B	The concentration of Ca in leachate from unstabilised and stabilised KMT samples.	128
6.4A	The migration profiles of Cu in unstabilised and stabilised KMT samples.	130
6.4B	The concentration of Cu in leachate from unstabilised and stabilised KMT samples.	130
6.5A	The migration profiles of Cr in unstabilised and stabilised KMT samples.	132
6.5B	The concentration of Cr in leachate from unstabilised and stabilised KMT samples.	132
6.6A	The migration profiles of Ni in unstabilised and stabilised KMT samples.	134
6.6B	The concentration of Ni in leachate from unstabilised and stabilised KMT samples.	134
6.7A	The migration profiles of Pb in unstabilised and stabilised KMT samples.	137
6.7B	The concentration of Pb in leachate from unstabilised and stabilised KMT samples.	137
6.8A	The migration profiles of Zn in unstabilised and stabilised KMT samples.	139
6.8B	The concentration of Zn in leachate from unstabilised and stabilised KMT samples.	139
6.9A	The pH value of unstabilised and stabilised KGB samples.	141
6.9B	The pH value of leachate from unstabilised and stabilised KGB samples.	141
6.10A	The migration profiles of Ca in unstabilised and stabilised KGB samples.	143
6.10B	The concentration of Ca in leachate from unstabilised and stabilised KGB samples.	143



6.11A	The migration profiles of Cu in unstabilised and stabilised KGB samples.	144
6.11B	The concentration of Cu in leachate from unstabilised and stabilised KGB samples.	144
6.12A	The migration profiles of Cr in unstabilised and stabilised KGB samples.	146
6.12B	The concentration of Cr in leachate from unstabilised and stabilised KGB samples.	146
6.13A	The migration profiles of Ni in unstabilised and stabilised KGB samples.	148
6.13B	The concentration of Ni in leachate from unstabilised and stabilised KGB samples.	148
6.14A	The migration profiles of Pb in unstabilised and stabilised KGB samples.	150
6.14B	The concentration of Pb in leachate from unstabilised and stabilised KGB samples.	150
6.15A	The migration profiles of Zn in unstabilised and stabilised KGB samples.	152
6.15B	The concentration of Zn in leachate from unstabilised and stabilised KGB samples.	152
6.16A	The pH value of unstabilised and stabilised LDA samples.	154
6.16B	The pH value of leachate from unstabilised and stabilised LDA samples.	154
6.17A	The migration profiles of Ca in unstabilised and stabilised LDA samples.	156
6.17B	The concentration of Ca in leachate from unstabilised and stabilised LDA samples.	156
6.18A	The migration profiles of Cu in unstabilised and stabilised LDA samples.	158
6.18B	The concentration of Cu in leachate from unstabilised and stabilised LDA samples.	158
6.19A	The migration profiles of Cr in unstabilised and stabilised LDA samples.	160

6.19B	The concentration of Cr in leachate from unstabilised and stabilised LDA samples.	160
6.20A	The migration profiles of Ni in unstabilised and stabilised LDA samples.	162
6.20B	The concentration of Ni in leachate from unstabilised and stabilised LDA samples.	162
6.21A	The migration profiles of Pb in unstabilised and stabilised LDA samples.	163
6.21B	The concentration of Pb in leachate from unstabilised and stabilised LDA samples.	164
6.22A	The migration profiles of Zn in unstabilised and stabilised LDA samples.	166
6.22B	The concentration of Zn in leachate from unstabilised and stabilised LDA samples.	166
6.23	Permeability of unstabilised and stabilised KMT samples with leaching 7PV.	175
6.24	Permeability of unstabilised and stabilised KGB samples with leaching 7PV.	176
6.25	Permeability of unstabilised and stabilised LDA samples with leaching 7PV.	177
6.26	(A) Photomicrograph of KMT0% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	179
6.26	(B) Photomicrograph of KMT0% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	180
6.26	(C) Photomicrograph of KMT2% samples (depth of 1.0cm) cured for 2 weeks and leached with acid at the top part.	180
6.26	(D) Photomicrograph of KMT4% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	181
6.26	(E) Photomicrograph of KMT6% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	181
6.26	(F) Photomicrograph of KMT8% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	182
6.27	(A) Photomicrograph of KGB0% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	182

6.27	(B) Photomicrograph of KGB2% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	183
6.27	(C) Photomicrograph of KGB4% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	183
6.27	(D) Photomicrograph of KGB6% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	184
6.27	(E) Photomicrograph of KGB6% samples cured (depth of 1.0cm) for 4 weeks and leached with acid.	184
6.27	(F) Photomicrograph of KGB8% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	185
6.28	(A) Photomicrograph of LDA2% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	186
6.28	(B) Photomicrograph of LDA2% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	186
6.28	(C) Photomicrograph of LDA4% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	187
6.28	(D) Photomicrograph of LDA6% samples (depth of 8.0cm) cured for 2 weeks and leached with acid.	187
6.28	(E) Photomicrograph of LDA6% samples (depth of 1.0cm) cured for 4 weeks and leached with acid.	188
6.28	(F) Photomicrograph of LDA8% samples (depth of 1.0cm) cured for 2 weeks and leached with acid.	188
6.28	(G) Photomicrograph of LDA6% samples (depth of 8.0cm) cured for 4 weeks and leached with acid.	189
7.1	The ternary composition diagram of KMT sample, KGB sample, LDA samples, and FM samples. (Modified from: Bentur, 2002)	194
7.2	The illustration of mineral assemblages in unstabilised soil showing the high percentage of pore spaces in (A) stabilised with 2% of lime shows the formation of cementitious mineral (B) stabilised with 4% of lime shows the increase of pore spaces due to flocculation reaction (C) stabilised with 6% of lime shows the flocculation-agglomeration reaction (D) stabilised with 8% of lime shows the unreacted lime created weak boundary within the particle (E).	203

- 7.3 The schematics of soil particle arrangement, pore space and of preferential channeling of (A) unstabilised and (B) stabilised KMT, KGB, and LDA samples, which controlled their permeabilities. 210
- 7.4 The plotted graphs of Ca versus pH in leachate sample from (A) KMT samples and (B) KGB samples and, (C) LDA samples. 212
- 7.5 The schematic of lime stabilised double liner system, stabilised soil cover and stabilised top layer for a hazardous waste landfill. 217
- 7.6 Detail schematic of profiles of lime stabilised double liner for hazardous waste landfill. 219

**LIST OF NOTATIONS**

AAS	-	Atomic Absorption Spectrophotometry
ASTM	-	American Society of Testing and Material.
BS	-	British Standard.
Ca	-	Calcium.
CASH	-	Calcium Aluminate Silicate Hydrate.
CEC	-	Cation Exchange Capacity.
Ce/Co		Relative Concentration
Cr	-	Chromium.
CSH	-	Calcium Silicate Hydrate.
Cu	-	Copper.
I <sub>p</sub>	-	Plasticity Index.
k	-	Permeability.
L <sub>L</sub>	-	Liquid Limit.
Meq	-	Milliequivalents.
Ni	-	Nickel.
OMC	-	Optimum Moisture Content.
Pb	-	Lead.
ρ <sub>D</sub>	-	Dry density.
P <sub>L</sub>	-	Plastic Limit.
PV	-	Pore Volume.

ppm	-	Parts Per Million.
SEM	-	Scanning Electron Microscope.
SG	-	Specific Gravity.
S <sub>L</sub>	-	Shrinkage Limit.
SSA	-	Specific Surface Area.
W <sub>o</sub>	-	Moisture content.
XRD	-	X-Ray Diffraction.
XRF	-	X-ray Fluorescence.
Zn	-	Zinc.

## LIST OF APPENDICES

APPENDIX NO		PAGE
A	List of publications.	251
B	Materials from the study area.	253
C	Composition of selected heavy metals in earth crust and different type of rock. (Turekian and Wedepohd, 1961 and Vinogradov, 1962)	256
D	Heavy metals content in rock forming minerals and the resistant of minerals with weathering process. (Modified from Deer <i>et al.</i> , 1989 and Loughnan, 1969)	257
E	Heavy metals level in soil.	258
F	The groups and common minerals in the phyllosilicate group. (Mackenzie, 1975 and Brindley <i>et al.</i> , 1968)	260
G	Leaching process of heavy meals from the improper design of landfill of abandoned of landfill. (Sharma and Lewis, 1994).	261
H	The schematic of a cover system and double liner system for hazardous waste landfill. (Landreth, 1990)	262
I	The schematic of a double liner and leachate collection system for a hazardous waste landfill (USEPA, 1989b).	263
J	pH effect on precipitation of three heavy metals Pb, Cu, and Zn (Mc Donald, 1994)	264
K	Cylindrical mould with plugs for preparation of specimens for compressive strength test (BS 1924 Part 1: 1990)	265
L	The schematic of Leaching cell.	266

M	The operating parameter use in the flame atomic absorption.	268
N	Calibration of tritestng instrument for unconfined compression test.	269
O	Calibration graph for AAS analysis.	270
P	Calibration graph for major element using XRF	281
Q	Particle Size Distribution of Stabilised Samples.	286
R	The compaction curves of the samples.	289
S	The unconfined compressive strength of samples.	290
T	X-ray diffractograms of samples.	291
U	Classification of pore system based on IUPAC classification System (Mikhail and Robens, 1983).	296
V	EDAX patterns of stabilised minerals.	297
W	Plotted graph of Ca versus heavy metals and pH versus heavy metals in leachates samples.	302
X	Bar chart of mass balance calculations.	308
Y	Flowchart for evaluation the stabilisation options (Cullinane <i>et al.</i> , 1986)	314



## CHAPTER I

### INTRODUCTION

#### 1.1 Related Environmental Issues

The global environmental issues today focused on the problems of contamination of land, water and air. Disposal of municipal solid waste and chemical contamination have become a major problem currently faced by most developing and industrial countries. This is due to the increasing population (consumer) and rapid increase in the industrial by-products. In conjunction with the continuing environmental problems, the United Nations held a Conference on Environment and Development (UNCED), in Rio de Janeiro in 1992. As a result, the conference declared the fundamental principles and the programmed action for achieving sustainable development, which is called 'Rio Declaration' (UNCED, 1992). The Principle 15 in the Rio Declaration mentioned, that *"In order to protect the environment, the precautionary approach shall be widely applied by states according to their capabilities. Where there are threats of serious or irreversible damage, lack of full scientific certainty shall not be used as a reason for postponing cost-effective measures to prevent environmental degradation"*. For that reason, part of declaration has urged the researchers to focus their research to find the implementation of the best technique to manage the toxic chemical. The scientific and technological activity in chapter 19 of Agenda 21 in Rio Declarations aims to strengthen research on safer alternatives to toxic chemicals that pose an unreasonable and otherwise unmanageable risk to the environment or human health and to those that are toxic, persistent and bio-accumulative and that cannot be adequately controlled. This indicates that scientific research is needed to prevent and to manage the toxic

chemical, which has potential to degrade our environment as well as human health. Therefore, one of the major task faced by the geoenvironmental engineering researchers today is to develop methods to evaluate the release, transport, and long term changes of contaminants from the waste disposal sites such as landfill, containment barrier and, contaminated land.

In conjunction with the Rio Declarations, the world summit on sustainable development that was held in Johannesburg in 2002 has renewed the Declarations to manage chemicals throughout their life cycle and of hazardous wastes for sustainable development and for the protections of human health and the environment (UNCED, 2002). Amongst the actions were to encourage partnership to promote activities aimed at enhancing environmentally sound management of chemicals and hazardous wastes and to promote reduction of the risks posed by heavy metals that are harmful to human health and the environment.

In Malaysia, the phenomenon of environmental degradation showed similar pattern with most other countries where contamination is a major issue. For example the trend of water quality status of river basins over the past 10 years (trend between 1990 – 1999) show the reduction in the number of clean river (DOE, 1999). However, it was also reported that in terms of heavy metal contamination out of the 5,274 river water samples analysed for heavy metals only 7% of the samples exceeded Class III (Extensive Treatment Required). It is noticed that two main sources of water pollution were manufacturing industries and sewage effluent, where the percentage is 43.7% and 45.5% respectively. In Malaysia, generally land contamination is still under control and low, nevertheless, certain areas such as mining waste containment and landfill area should be concerned due to the high risk of environmental degradation. In the long term processes the accumulation of chemicals such as heavy metals in that particular area would cause enrichment of the chemical contaminants. The heavy metals also could degrade the land and water systems by leaching processes. It could bring damage to human health through drinking water or food from plants and animals, which took the chemicals from the contaminated land and contaminated water.

Normally a clay liner is used at the bottom part of a landfill to provide protection for the surrounding soil and ground water from contamination through leaching process. In general, a liner system consists of both a geomembrane barrier and a natural clay liner. Although the geomembrane is an absolute barrier nevertheless, heavy metals pollutants can permeate it by permeation due to different pressure and diffusion (Sharma & Lewis, 1994). In other words the leaching process also could occur through the geomembrane in the landfill area. In an improperly managed or old landfill area and mining waste containment, the liquid, which contained heavy metal is derived from the interaction of rainwater with solid waste or contaminated soil, contact with surface water bodies or groundwater. Hence, the heavy metals such as copper, chromium, nickel, lead, and zinc from the hazardous waste can diffuse into the groundwater or surface water systems. To prevent the contaminant from passing into the surrounding environment, the first step in landfill and containment design is to site them far from the groundwater table and far from groundwater abstraction wells. The second step is to modify or stabilise the clay liner to increase the capability of soil to adsorb and immobilise the heavy metals. Furthermore, the waste itself can be stabilised with suitable chemical such as lime and other active material. Stabilisation or solidification using lime could also improve the physical properties of the soil and capability to increase its durability and strength. Therefore, understanding the physical characteristic, chemical properties, mineralogical identification, microstructural study of both soil and waste are important for the stabilisation purposes. It is also necessary to understand the nature of the environment and the transportation processes of heavy metals in soil and water systems by means of leaching test. Leaching test is essential to evaluate the performance of lime and additive chemical as a stabilisation agent to immobilise the heavy metals in the landfill area and in the mining waste containment. Design and development of physical laboratory testing is essential to assess the potential application before suggestion can be made. This research will look into two major applications namely the stabilisation of contaminated soil and the improvement of landfill system through stabilised clay liner.

## 1.2 Study Area

The study areas are located in Kayumadang Telipok, Kg. Bongkud Ranau and Lohan Dam Ranau, Sabah. Kayumadang is one of the villages in Telipok, Sabah and located about 20 km from the northern part of Kota Kinabalu City Centre. There is a new landfill area in this village, which commenced operation in 1995. The waste disposal from domestic consumer and waste from the industrial area in the Kota Kinabalu area were disposed of in this landfill. The soil samples were collected from the road cuts near the landfill area and named as KMT samples (Kayu Madang, Telipok) i.e situated at the longitude of  $116^{\circ}10.722''\text{E}$  and latitude  $06^{\circ}06.441''\text{N}$ .

Kg Bongkud is situated in the Ranau district that is in the northern part of Lohan valley, Ranau is located at the longitude of  $116^{\circ}44.988''\text{E}$  and latitude  $06^{\circ}01.666''\text{N}$ . This village is bordered by intrusion of igneous rock in the northern part. The igneous rocks intruded the sedimentary rock of Crocker Formation. The intrusion of igneous is believed to be the factor of mineral accumulation in Ranau area and has good opportunity for the mining activity. The soil sample was collected from the hilly area. This residual soil derived from the weathering of igneous and sedimentary rocks are named as KGB samples (Kg. Bongkud, Ranau).

Mamut Copper Mining is located in Ranau district at the western coast of Sabah, Malaysia about 68 kilometers to the eastern part of Kota Kinabalu City Center, that is at the longitude of  $116^{\circ}44.166''\text{E}$  and latitude  $06^{\circ}00.691''\text{N}$ . It is the largest copper mine, and the largest producer of gold and copper in Malaysia. But, there are not much mineable reserve left when the mines end its production in the year 2000. The production of waste and haul ore average rate is 10 million tonnes and 6 million tonnes a year respectively. The mining waste is dumped in the eastern part of the mining area located at Lohan Dam, Ranau, Sabah. The mining waste sample obtained is called LDA samples (Lohan Dam).

### 1.3 Statement of Problem

Chemical and physical processes could alter the hazardous waste in a landfill system. One of the alteration or degradation products is leachate. The leachate contain heavy metal which are some of the most harmful wastes, hence will produce contamination into the environment. Abandoned and improperly managed landfill could be also the other sources of heavy metals contamination. In the mining waste containment the weathering of sulphide minerals such as pyrite and galena, which produce acid drainage is the main sources of the heavy metals. The migration of heavy metals from the hazardous waste or mining waste into the clay liner in landfill system occurred in an acidic environment. Therefore, two sources of contaminant are identified namely:

- i. Contaminant from landfill area through clay liner material represented by KMT and KGB samples.
- ii. Contaminant from mining waste containment represented by LDA sample.

The purpose of clay liner is to enhance environmental protection. They serve as a hydraulic barrier to the flow of contaminants. Therefore, they must have a low hydraulic conductivity for long periods of time. They act mainly to retain the contaminants. Understanding the leaching behaviour of heavy metals in the contaminated soil and mining waste is important in order to develop a method to restrain their mobility.

As a summary, the sources of heavy metals in the studied area were originated from the leaching of landfill leachate, and leaching from the mining waste. Therefore, understanding the leaching behaviour of heavy metals in the landfill liner and mining waste is important in order to develop a method to manage the contaminated soil and waste in this area. One of the possible solutions for the treatment of leachate generating hazardous waste or mining waste would be employing lime stabilisation. Clay liner could be stabilised with lime to increase its capacity to retain the heavy metals. Lime could react with clay and produce cementitious minerals, which crystallized with age hence increased the retention of

heavy metals in the clay liner. In the long term, the development of cementitious minerals and clogging of fine materials could decrease the permeability of the clay liner, therefore increase the capability of the soil to protect the migration of heavy metals from degrading the environment.

#### 1.4 Objectives

The objectives of this study are:

- i. To determine the chemical properties, physical properties, mineralogical identification, microstructural, and geotechnical properties of KMT, KGB, and LDA samples.
- ii. To establish the effect of lime on KMT, KGB, and LDA samples properties as follows:
  - a. Chemical properties: pH, cation exchange capacity (CEC), specific surface area (SSA), and bulk chemistry.
  - b. Physical properties: Atterberg limits specific gravity (SG), and particle size distribution.
  - c. Mineralogical identification for both of clayey soil and mining waste using X-ray diffraction (XRD), and scanning electron microscope (SEM).
  - d. Geotechnical properties such as compaction, unconfined compressive strength (UCS) and permeability.
- iii. To examine the effect of lime on leaching process of heavy metals in KMT, KGB, and LDA samples.
- iv. To produce an application design for controlling contamination from landfills systems, and mining waste.

## 1.5 Scope of the Study

- i. All of the samples were collected from Sabah i.e soil from Kg. Kayu Madang Telipok (KMT), soil from Kg. Bongkud, Ranau (KGB) and mining waste from Lohan Dam, Ranau (LDA).
- ii. Hydrated lime [ $\text{Ca}(\text{OH})_2$ ] was used as a major stabilisation agent.
- iii. Supplement pozzolana for the mining waste was collected from Kg. Melawa, Sepanggar, Kota Kinabalu (FM).
- iv. Laboratory works include of physical, chemical, mineralogical, microstructural and, geotechnical properties of the material before leaching and after leaching for both stabilised and unstabilises condition.
- v. Leaching column tests were conducted mostly until 7 pore volumes (PV) of contaminant in order to assess the suitability of lime as a stabilising agent and immobiliser for the selected heavy metals in KMT, KGB, and LDA samples.
- vi. The category of contaminant comprises of selected heavy metals (chromium, copper, nickel, lead and zinc) are manually spiked into the leaching columns to represent the contaminant from the landfill site, and mining waste.

## 1.6 Significance of Research

This study is important for the geotechnical engineers or land development agencies to plan any construction involving landfill site, contaminated soils and mining waste containment. Special measurements, treatment and design procedures are important and need to be fully understood. Huge distribution of soil from the weathering of sedimentary rock (Crocker Formation) in Kota Kinabalu area, the occurrence of contaminated soil in Kg. Bongkud area and, million of tonnes of mining waste containment in almost 1000 acres have contributed to a significant

effect on land use activity as well as the health of the people within the vicinity. Therefore, it is vital to study the suitability and effect of lime to stabilise the heavy metals in mining waste and contaminated soil.

## **1.7 Thesis Organisation**

This thesis consists of eight chapters, and the contents of each chapter is explained as follows:

**CHAPTER I:** This introductory chapter presents the current issues of environment, which relates to the statement problem in the study area. It also explains the scope and the significance of this research.

**CHAPTER II:** This chapter reviews the source of heavy metals in soil and their concentration in contaminated soil. The remediation of contaminated soil will be discussed. The lime stabilisation will also be discussed in detail, where it consists of the mineralogy of lime-soil reactions, microstructural development, physical characteristics and engineering properties.

**CHAPTER III:** Chapter III presents the sampling site of KMT, KGB, and LDA samples in Sabah. In addition to the description of the properties this chapter also will highlight the general test procedures of the physical and chemical test and also the concept of SSA, XRD, XRF, and SEM studies.

**CHAPTER IV:** Chapter IV explained the effect of lime on KMT, KGB, and LDA samples properties. The effect of lime on the physical characteristics consists of Atterberg limits, specific gravity, and particle size distributions. The analysis of chemical properties includes, cation exchange capacity, pH, specific surface area, and X-ray fluorescence will analysis for the chemical properties. Finally, the geotechnical tests such as compaction, unconfined compressive strength and, permeability will be presented.



## REFERENCES

- Airoldi, C. and Critter, S.A.M. (1997). "Brazilian Red Latosol a Typic Soil as an Exchanger a Thermodynamic Study Involving Cu, Zn, Cd, Hd, Pb, Ca and Na." *Clays and Clay Minerals*. **15(2)**. 125-131.
- Alba, N, Vazquez, E, Gasso, S, and Baldasano, J.M. (2001). "Stabilization/solidification of MSW incineration residues from facilities with different air pollution control systems." Durability of matrices versus carbonation. *Waste Manage*. **21**. 313-323.
- Albino, V. Cioffi, R. DE Vito, B. and Santoro, L. (1996). "Evaluation of Solid Waste Stabilisation Processes by means of Leaching Tests." *Environmental Techno*. **17**. 309-315.
- Alloway, B. J. (1995). "*Heavy Metals In Soils*, 2<sup>nd</sup> Ed." Black Academic & Professional.
- Al-Mhaidib, A.I. and Al-Shamrani, M.A. (1996). "Swelling Characteristics of Lime-Treated Expansive Soils." *Geotech. Engr. J.* **27(2)**. 37-54.
- Ambatsian, P, Fernex, F, Bernat, M, Parron, C. and Lecolle. (1997). "High metal inputs to closed seas: the Mew Caledonian lagoon." *J. of Geochemical Exploration*. **59**. 59 – 74.
- American Standard for Testing and Materials, Part II. Test Designation D 2166-661. "Unconfined Compressive strength of cohesive soil." Philadelphia, USA.

- Anderson, W. C. 1994 (Ed). "*Innovative Site Remediation Technology: Stabilization/Solidification.*" Springer-Verlag: USA.
- Arabi, M. and Wild, S. (1986). "Microstructural development in cured soil-lime composites." *J. Mater. Sci.* **32**. 497-503.
- Assarson, K.G. (1976). "Note on Stabilisation of Soil with Lime. In: Broms, B. *Stabilisation of soil with lime columns: Design Handbook 3<sup>rd</sup>.*" Edition. Lime Column AB. Singapore.
- Barna, R. Sanchez, F, Moszkowics, P. and Mehu, J. (1997). "Leaching behavior of pollutants in stabilized/solidified wastes." *J. of Hazardous Materials.* **52**. 287 – 310.
- Basir Jasin, Sanudin Hj. Tahir and Tating, F.F. (1991). "Late Eocene planktonic foraminifera from the Crocker Formation, Pun Batu Sabah." *Warta Geologi.* **17(4)**.
- Bell, F.G. (1988). Stabilisation and treatment of clay soils with lime: Part 1 - Basic principles. *Ground Engineering*, Jan. 10-15.
- Bell, F.G. (1996). "Lime stabilization of clay minerals and soils." *Engrg. Geol.* **42**. 223-237.
- Benson R. E. Jr. (1980). "*Natural fixing materials for the containment of heavy metals in Landfills.*" In Proc. Mid-Atlantic Industrial Waste
- Benson, C. H, Daniel, D. E, Boutwell, G. P. (1999). "Field Performance of Compacted Clay Liners." *J. of Geotechnical and Geoenvironmental Engineering* **125(5)**. 390 – 403.
- Bentur, A. (2002). "Cementitious Materials – Nine millennia and a new century: Past, Present, and Future." *J. Mater. Civ. Eng.* **14(1)**, 2 – 22.

- Bertholf, R.L. (1988). Zinc. "*Handbook on Toxicity of Inorganic Compounds.*" Marcel Dekker, New York. 792-792.
- Bhatty, M.S.Y. (1973). "Fixation of metallic ions in Portland Cement Assoc. in Conner J.R. *Chemical Fixation and Solidification of Hazardous Waster.* Van Nostrand Reinhold. New York. 357 – 359.
- Borden, R.H. Baer, J.I. (1996). "Testing techniques for evaluating the shear strength of lime/fly ash slurry stabilized soil." *Geotech. Testing J.* **14(3)**. 247-256.
- Bowen, H.J.M. (1979). "*Environmental Chemistry of the Elements.*" Academic Press, London.
- Boynton, R.S. and Gutschick, K.A. (1996). "*Lime* . In Boynton, R.S. (Ed.)." *Industrial Minerals and Rocks.*
- Brady, N. C, and Weil, B.R. (1999). "*The nature and Properties of Soils.*" Prentice Hall, New Jersey. 344-346.
- Brault, J.L. (1991). "*Water Treatment Handbook.* Degremont." Paris. 229.
- Brindley, G. W. Bailey, G.T. Faust, S. A. Forman, and Rich, C. I. (1968). "Report of the nomenclature committee (1966-1967) of the Clay Minerals Society. *Clays Clay Miner.* 16:322 – 324 in Tan, K.H. *Principles of Soil Chemistry* (2<sup>nd</sup> Ed)." Marcel Dekker, Inc. 260-261.
- British Standard Institution. (1990). "*Part. 2: 4.3. Liquid Limit.*" London: BS 1377
- British Standards Institution. (1984). "Part 2. *Methods of chemical analysis. Quicklime, hydrated lime and natural calcium carbonate.*" British Standards Institutions. London: BS6463:
- British Standards Institution. (1990). "*BS 1377 Part. 2: 3.2: Moisture Content.*" London: BS 1377

- British Standards Institution. (1990). "*Part. 2: 4.0. Loss On Ignition (Organic Matter)*." London: BS 1377.
- British Standards Institution. (1990). "*Part. 2: 5.3. Plastic Limit.*" London: BS 1377
- British Standards Institution. (1990). "*Part. 2: 5.4. :Plasticity Index.*" London: BS 1377
- British Standards Institution. (1990). "*Part. 2: 6.4. Shrinkage Limit.*" London: BS 1377.
- British Standards Institution. (1990). "*Part. 2: 8.3. Specific Gravity (Particle Density)*." London: BS 1377
- British Standards Institution. (1990). "*Part. 2: 9.3. Dry Sieving.*" London: BS 1377
- British Standards Institution. (1990). "*Part. 2: 9.4. Particle Size Analysis.*" London: BS 1377.
- British Standards Institution. (1990). "*Part. 3: 9.0. pH analysis.*" London: BS 1377.
- British Standards Institution. (1990). "*BS1377: Part 4: 3.3. Compaction Test.*" London: BS 1377.
- British Standards Institution. (1990). "*BS1924: Part: 1. Stabilized Materials for Civil Engineering Purposes.*" London.
- Broms, B. (1984). "*Stabilisation of soil with lime columns: Design Handbook 3<sup>rd</sup>.*" Edition. Lime Column AB. Singapore.
- Carroll, S.A, O'day, P.A. and Piechowski, M. (1998). "Rock-Water Interactions Controlling Zinc, Cadmium, and Lead Concentrations in Surface Waters and Sediments, U.S. Tri-State Mining District. 2. Geochemical Interpretation." *Environ. Sci. Technol*, **32**: 956-965.

- Cary, E.E. (1982). In Sieler & Sigel (1988). *Handbook on Toxicity of Inorganic Compounds*. Marcel Dekker, New York. 245-248.
- Chapman, H.D. (1965). Cation exchange capacity. In Black, C.A (Ed.). "Methods of Soil Analysis, Part 2. Chemical and Microbiological Properties." *Agronomy no. 9*. American Society of Agronomy.
- Cheng, K.Y. and Bishop, P.L. (1992). Sorption, Important in Stabilized/Solidified Waste Forms. *Hazardous Waste & Hazardous Materials*. **9(3)**. 289-296.
- Christophi, A.C. and Axe. L (2000). "Competition of Cd, Cu, and Pb Adsorption on Goethite." *J. of Environmental. Eng.* **126(1)** .66 – 74.
- Chummar, A.V. (1990). "Soil stabilization by lime." *Symposium on Recent Advances in Geotechnical Engineering*, Singapore. 66-68.
- Clevenger, T. E. (1990). "Use of sequential extraction to evaluate the heavy metals in mining waste." *Water, Air, Soil Pollution Journal*. **50**. 241 –254.
- Collenette, P. (1958). "The geology and mineral resources of the Jesselton, Kinabalu Area, North Borneo." Geol. Survey Department, British Territories in Borneo Memior. **12**. 107.
- Conner J.R. (1990). "Chemical Fixation and Solidification of Hazardous Waster." Van Nostrand Reinhold. New York. 86, 291.
- Consoli, N.C. Prietto, P.D.M, Carraro, J.A.H. and Heineck, K.S. (2001). "Behavior of Compacted Soil-Fly-Ash-Carbide Lime Mixtures." *J. of Geotech. and Geoenvironmental Eng.* **127(9)**. 774-782.
- Cullinane, M. J. Jr, Jones, L. W. and Malone, P. G. (1986). "Handbook for Stabilization/ Solidification of Hazardous Waste." Hazardous Waste Engineering Research Laboratory, USEPA, Washington, D.C. 102.

- Czurda, A.K. (1993). "The triple multimineral barrier for hazardous waste encapsulation." *Engrg. Geol.* **34**. 205-209.
- Daniel, D. E, and Wu, Y. K. (1993). "Compacted Clay liners and Covers for Arid Sites." *J. Geotech. Engr.* **119(2)**. 223 –237.
- Daniel, D. E. (1993). "Case Histories of Compacted Clay Liners and Covers for Waste Disposal Facilities." *In Proceedings of the 3<sup>rd</sup> International Conference on Case Histories in Geotechnical Engineering*, St. Louis, Mo., S. Prakash, Ed. 1407 – 1425.
- Davis, A. Eary, L.E. and Helgen, S. (1999). "Assessing the Efficacy of Lime Amendment to Geochemically Stabilize Mine Tailings." *Environ. Sci. Technol.* **33**. 2626-2632.
- Deer, W.A. Howie, R.A. Zussman, J. (1989). "*An introduction to the rock forming minerals.*" Longman, England.
- Delattre-Louvel, F., Lapeyre, C., Struillou, R. and Arnould, M. (1993). "Retention mechanisms for toxic cations using artificial confinement barriers of clay-cement mixtures." *Engrg. Geol.* **34**. 151-158.
- Department of Environment Malaysia (DOE). (1999). "*Malaysia Environmental Quality Report 1999.*" Department of Environment Ministry of Science, Technology and the Environment Malaysia.
- Diamond, S. and Kinter, E.B. (1966). "Mechanisms of Soil-Lime Stabilisation: An Interpretive Review." *Public Roads*. **33(12)**. 260.
- Edil, T.B and Berthouex, P.M. (1990). "Earthen Barriers Technology for Waste Containment." *Waste Manage.* **10**. 147-153.
- Elzahabi, M. and, Yong, R. N. (2000). "PH influence on sorption characteristics of heavy metal in the vadose zone." *Engrg. Geol.* **60**. 61 – 68.

- Emmerich, C. E, Lund, L. J, Page, A, L. Chang, A.C. (1982). "Solid phase forms of Heavy metals in sewage sludge treated soils." *J. of Environmental Quality*. **11**. 178-181.
- Faisal Hj. Ali. (1990a). "Improvement of a residual Soil." *Tenth Southeast Asian Geotechnical Conference*. Taipei. 3-8.
- Faisal Hj. Ali. (1990b). "The use of rice husk ash in soil stabilization." *Symposium on Recent Advances in Geotechnical Engineering*, Singapore: 1-6.
- Falcone, J.S. (1982). "Soluble Silicates. In Conner J.R. *Chemical Fixation and Solidification of Hazardous Waster*." Van Nostrand Reinhold. New York. 86 – 91.
- Felix Tongkul. (1987). "Sedimentology and Structure of the Crocker formation in the Kota Kinabalu Area, Sabah East, Malaysia." PhD Tesis.Univ. of London.
- Foged, N. and Baumann, J. (1999). "Clay membrane made of natural high plasticity clay: leachate migration due to advection and diffusion." *Engrg. Geol.* **54**.129-137.
- Ford, R.G, Scheinost, A.C, Scheckel, K.G, and Sparks, D.L. (1999). "The Link between Clay Mineral Weathering and the Stabilization of Ni Surface Precipitates." *Environ. Sci. Technol.* **33**. 3140-3144.
- Furuichi, T. and Tanaka, M. (1993). "Monitoring system for environmental risk management of a landfill disposal site." *Engrg. Geol.* **34**. 169-174.
- Gallego, J. L. R. Ordonez, A. and Loreda, J. (2002). "Investigation of trace element sources from an industrialized area (Alives, northern Spain) using multivariate statistical methods." *Environmental International*. **27**. 589 – 596.

- Garcia-Sanchez, A, Alastuey, A. and Querol, X. (1999). "Heavy metal adsorption by different minerals: application to the remediation of polluted soils." *The Sci. of the Total Environment*. **242**. 179 – 188.
- Gardner, L. R. (1990). "Geochemical analysis of silicate rocks and soils by XRF using pressed powders and a two-stage calibration procedure." *Chem. Geology*. **88**. 169-182.
- Gelinas, Y. Barnes, R.M. Florian, D., and Schmit, J.P. (1998). "Acid Leaching of Metals from Environmental Particles: Expressing results as a Concentration within the Leachable Fraction." *Environ. Sci. Technol.* **32**. 3622-3627.
- Ghosh, A. and Subbarao, C. (2001). "Microstructural Development in Fly ash modified with Lime and Gypsum." *J. of Material in Civil Eng.* **13(1)**. 65-70.
- Ghosh, A. and Subbarao, C. (1998). "Hydraulic Conductivity and Leachate Characteristics of Stabilized Fly Ash." *J. Environ. Engrg.* **124(9)**. 812-820.
- Goldberg, S, Su, S and Forster, H. S., (1998). "*Sorption of Molybdenum on Oxides, Clay minerals and soil. In: Adsorption of metals by Geomedia. Jeanne E. A (Ed. ).*" Academic Press, San Diego.
- Gordon, M.E. Huebner, P.M and Miazga, T.J. (1989). "Hydraulic conductivity of three Landfill Clay liners. *J. Geotech. Engrg.*" **115(8)**. 1148-1160.
- Gosh, A and Subbarao, C. (1995). "Hydraulic conductivity and Leachate characteristics of Stabilized Fly ash." *J. Environ. Engrg.* **124(9)**. 812-820.
- Grim, R. E. (1953). "*Clay Mineralogy.*" McGraw-Hill, New York.
- Hamilton, J. W and Wetterhahn, K.E. (1986). "Chromium. In Sieler & Sigel: *Handbook on Toxicity of Inorganic Compounds.*" Marcel Dekker, New York. 245-248.



- Harris, A.W. Manning, M.C, Tearle, W. M. and Tweed, C.J. (2002). "Testing of models of the dissolution of cements – leaching of synthetic CSH gels. *Cement & Concrete Res.*" **32**. 731 – 746.
- Harris, M.R and Denner J. (1997) UK Government Policy and Controls. "In Hester R.E and Harrison, R. M. (Ed)." *Contaminated land and its reclamation*. Thomas Telford. Great Britain. 39-41.
- Hausmann, M. R. (1990). "*Engineering Principles of Ground Modification.*" McGraw-Hill Publishing Company, New York. 310-311.
- Head, K. H. (1992). "*Manual of Soil Laboratory Testing.*" Pentech Press, London. 125.
- Head, K.H. (1982). "*Manual of Soil Laboratory Testing: Volume 2: Permeability, Shear Strength, and Compressibility Tests.*" Pentech Press, London.
- Heimann, R. B, Conrad, D. Florence, L. Z, Neuwirth, M, Ivey, D. G, Mikula, R. J. and Lam, W.W. (1992). Leaching of simulated heavy metals waste stabilized/solidified in different cement matrices. " *J. of Hazardous Materials*. **31**. 39 – 57.
- Herzog, A. and Mitchell, J.K. (1963). "*Reactions Accompanying Stabilization of Clay with Cement.*" Highway Research Record No. 36, Highway Research Board. 146-171.
- Hilmi, L.A and Aysen, L. M. (2000). "Microstructural development of Stabilized Fly ash as Pavement Base material." *J. of Materials in Civil Eng.* **12(2)**. 157-163.
- Hilt, G.H. and Davidson, T.T. (1960). "Lime fixation of clayey soils." *High, Res. Board. Bull.* 262. Washington, DF. 20-32.
- Holdgate, M.A. (1979). "*A Perspective of Environmental Pollution.*" University Press Cambridge.

- Hutchison, I. P. G, and Ellison, R, D. (1992). "*Mine Waste Management.*" Lewis Publishers. **270**. 142 - 144.
- Indraratna, B. Balasubramaniam, A.S. and Khan, M.J. (1995). "Effect of fly ash with lime and cement on the behaviour of a soft clay." *Quarterly J. of Engrg Geol.* **28**. 131-142.
- IUPAC Manual of Symbols and Terminology, (1972). "Appendix 1, Part I, Coll.," *Surface Chem. Pure Appl. Chem.* **31**. 578.
- Jacobson, G. (1970). "Gunong Kinabalu area, Sabah, Malaysia: Explanation of Part of Sheets 5-116-3 and 6-116-15. Report 8." Geological Survey Malaysia: Kuching, Sarawak.
- Johnson, C. A, Kersten, M, Zielger, F. and Moor, H. C. (1997). "Leaching behaviour and solubility – controlling solid phases of heavy metals in Municipal Solid Waste Incinerator Ash." *Waste Manage.* **16(1-3)**. 129 – 134.
- Kabata-pendias, A. and Pendias, H (1992). "*Trace Elements in Soils and Plants*, 2<sup>nd</sup> Ed.. CRC Press." Boca Raton, Fla.
- Kamon, M. and Nontananandh, R. (1991). "Combining Industrial Wastes with Lime for Soil Stabilization." *J. Geotech. Engrg*, **117(1)**. 1-16.
- Katz, A. (1998). "Microscopic study of alkali-activated fly ash." *Cement & Concrete Res.* **28 (2)**. 197-208.
- Krauskopf, K.B. (1956). "Dissolution and precipitation of silica at low temperatures." *Geochim. Cosmochim. Acta.* **10**. 1-26.
- Lambe, T. W. and Whitman, R.V. (1979). "*Soil Mechanics.*" New York: Wiley. 553.

- Lance, L.C. Hills, C.D. and Poole, A. B. (1997). "Effect of carbonation on properties of blended and non-blended cement solidified waste forms." *J. of Hazardous Material*. **62**. 193-212.
- Landreth, R. E. (1990). "Landfill Containment systems Regulations, in *Waste Containment Systems: Construction, Regulation, and Performance*." Geotechnical Special Publication 26. R. Bonaparte, Ed. ASCE. New York, 1 – 13.
- Li, L. Y. and Li, F. (2001). "Heavy Metal Sorption and Hydraulic Conductivity Studies Using Three types of Bentonite Admixes." *J. of Environmental Eng.* **127(5)**. 420-429.
- Lin, Z. (1997). "Mobilization and retention of heavy metals in mill-tailings from Garpenberg sulfide mines, Sweden." *The Sci. of the Total Environ.* **198**. 13-31.
- Lind, D. D, Walter, P.J. and Kingston, H.M. (1998). "Development and Validation of the EPA Microwave-Assisted Leach Method 3051A." *Environ. Sci. Technol.* **32**: 3628-3632.
- Livesey, I.D., 1988. "Lime stabilisation -current applications. In Rogers, C.D.F. Lime Stabilisation '88." *Ground Engineering*. **22(3)** .7-10.
- Locat, J. Berube, M.A. and Choquette, M., (1990). "Laboratory investigations of the Lime Stabilization of sensitive Clays: Shear Strength Development." *Can. Geotech. Journal*. **27**. 294-304.
- Locat, J. Tremblay, H. and Serge, L. (1996). "Mechanical and hydraulic behaviour of a soft inorganic clay treated with lime." *Can. Geotech. Journal*. **33**. 654-669.
- Lombardi, F, Mangilardi, T, Piga, L. and Sirini, P. (1998). "Mechanical and leaching properties of cement solidified hospital solid waste incinerator fly ash." *Waste Manage.* **18**. 99 – 106.

- Loughnan, F. C. (1969). "*Chemical Weathering of the Silicate Minerals.*" American Elsevier Publishing Company. New York. 1-3.
- Loughnan, F.C. (1969). "*Chemical weathering of the Silicate minerals.*" American Elsevier Publishing Company, Inc. New York. 32.
- Mackenzie, R. C. (1975). "*The classification of soil silicates and oxides. In Soil Components, Vol. 2. Inorganic Components, J. E. Gieseking (ed.) Springer-Verlag, New York. 1-25. In Tan, K.H. Principles of Soil Chemistry (2<sup>nd</sup>. Ed.)*" Marcel Dekker, Inc. 260-261.
- Malone, P.G. Jones, L.W and Larson, R.J. (1982). "*Guide to the disposal of chemically stabilized and solidified waste, SW-872.*" Office of Water and Waste Management, U.S. Environmental Protection Agency, Washington, DS, September.
- Marcus J, Gires U, and Murtedza, Mohamed. (1987). "Particle size Distribution of Copper Mine Tailings from Lohan Ranau Sabah and its Relationship with heavy Metal Content." *Pertanika* **10(21)**. 37-40.
- Martinez, C.E. and McBride, M.B. (1998). "Solubility of Cd<sup>2+</sup>, Cu<sup>2+</sup>, Pb<sup>2+</sup>, and Zn<sup>2+</sup> in Aqueous Coprecipitates with Amorphous Iron Hydroxides." *Environ. Sci. Technol.* **32**. 743-748.
- Masood, G. (1986). "Leachate collection systems." *J. of Environmental Eng.* **112(3)**. 613-617.
- Mathew, P.K. and Rao, S.N. (1997). "Effect of Lime on Cation Exchange Capacity of Marine Clay." *J. of Geotech. and Geoenvironmental Eng.* **123(2)**.183 –185.
- McBean, E.A, Rovers, F. A. and Farquhar, G. J. (1995). "*Solid Waste Landfill Engineering and Design.*" Prentice Hall New Jersey. 438 – 444.
- McGrath, S.P, Chang, A.C, Page, A. Land Witter, E. (1994). "In Alloway, B.J. (Ed.) . *Heavy Metals in Soils.*" Blackie Academic & Professional. London. 355.

- Mckinley, J.D, Thomas, H.R, Williams, K. P. and Reid, J.M. (2001). "Chemical Analysis of Contaminated Soil Strengthened by the addition of Lime." *Engrg. Geol.* **60**. 181-192.
- Means, J.L, Smith, L. A, Nehring, K. W, Brauning, S, E, Gavaskar, A,R, Sass, B, M, Wiles, C. C. and Mashni, C. I. (1995). "*The Application of Solidification/ Stabilization to Waste Materials.*" Boca Raton, CRC Press, Inc. 2.
- Mehmet, A, Y. and Bilge, A. (2001). "Leaching of metals from soil contaminated by mining activities." *J. of Hazardous Materials*, B87. 289-300.
- Mitchell, J. K. (1993). "*Fundamentals of soil Behavior (2<sup>nd</sup> Ed.)*." John Wiley and Sons, Inc. New York. 114-117
- Mostafa, N.Y, El-Hemaly, S.A.S, Al-Wakeel, E.I. El-Korashy, S. A. Brown, P. W. (2001). "Characterization and evaluation of the pozzolanic activity of Egyptian industrial by-products I: Silica fume and dealuminated kaolin." *Cem. & Concrete Res.* **31**. 467-474.
- Mowat, F. S. and Bundy, K. N. (2001). "Correlation of field-measured toxicity with chemical concentration and pollutant availability." *Environ. International*, **27**. 479 – 489.
- Mulligan, C. N, Yon, R. N. and Gibbs, B. F., 2000. "Remediation technologies for metal-contaminated soils and groundwater: an evaluation." *Engrg. Geol.* **60**. 193 – 207.
- Murtedza, Mohamed. and Heng, L.Y. (1985). "Heavy Metals in Rice Field Soil of Lohan Valley, Sabah." *Sumber* (1). 179-187.
- Netherlands Ministry of Housing, Physical Planning and Environment, (1986 & 1991). "*Environmental Quality Standards for Soil and /Water.*" Netherlands Ministry of Housing, Physical Planning and Environment. Leidschendam, Netherlands.

- O'day, P.A. Carroll, S. A. and Waychunam, G.A.G. (1998). Rock-Water Interactions Controlling Zinc, Cadmium, and Lead Concentrations in Surface Waters and Sediments, U.S. Ttri-State Mining District. 1. Molecular Identification Using X-ray Absorption Spectroscopy. *Environ. Sci. Technol.*, **32**. 943-955.
- O'Driscoll, M. (1988). "Burnt Lime/dolimine: Seeking markets green." Industrial Minerals 1988.
- Odler, J. and Colan-Subauste, J. (1999). "Investigations on cement expansion associated with ettringite formation." *Cement & Concrete Res.* **29**. 731-735.
- Olleir, C. K. (1969). "Weathering." Oliver and Boyd, Edinburgh.
- Omotoso, O.E, Ivey, D.G. and Mikula, R. (1998). "Containment mechanism of trivalent chromium in tricalcium silicate." *J. of Hazardous Material.* **60**. 1-28.
- Park, C.K., (2000). "Hydration and Solidification of Hazardous Wastews Containing Heavy Metals using Modified Cementitious Materials." *Cement & Concrete Res.* **30(3)**: 429-435.
- Park, J.Y and Batchelor, B. (1999). "Prediction of chemical speciation in stabilized/solidified waste using a general chemical equilibrium model Part 1. Chemical representation of cementitious binder." *Cement & Concrete Res.* **29**. 361 – 368.
- Parrish, W.6 & Lowitzsch, K. (1959). "Geometry, Alignment and angular calibration of x-ray diffractometers." *Amer. Mineral.* **44**. 765-787.
- Perkins, R. B. and Palmer, C. B. (1999). "Solubility of ettringite ( $\text{Ca}_6[\text{Al}(\text{OH})_6]_2(\text{SO}_4)_3 \cdot 26\text{H}_2\text{O}$ ) at 5 –76 °C." *Geochim. et Cosmochim Acta.* **63 (13/14)**. 1969-1980.
- Phipps, D. A (1981). "In Effect of Heavy Metal Pollution on Plants." Applied Science Publisher, London. 1-54.

- Pierce, J.F. and Witter, K.A. (1986). "Overburden pressures exerted on clay liners." *J. of Environmental Eng.* **112(9)**. 841-854.
- Poon, C.S. and Boost, M. (1993). "The stabilization of sewage sludge by pulverized fuel ash and related materials." *Environ. Tech.* **22(6)**. 443-448.
- Poon, C.S. Chen, Z. Q. and Wai, W.H. (2001). "The effect of flow-through leaching on the diffusivity of heavy metals in stabilized/solidified wastes." *J. of Hazardous Material*. **B81**. 179-192
- Posner, A.M. (1964). "Titration curves of humic acid. Trans. Int. Congr. Soil Sci. Bucarest, Romania. In Tan, K.H." *Principles of Soil Chemistry* (2<sup>nd</sup>. Ed). Marcel Dekker, Inc. 91-93.
- Prakash, K, Sridharan, A. and Rao, S.M. (1989). "Lime Addition and Curing Effects on the Index and Compaction Characteristics of a Montmorillonitic Soil." *Geotech. Eng.* **20**: 37-47.
- Prashanth, J. P, Sivapullaiah, P. V. and Sridharan, A. (2001). "Pozzolanic fly ash as a hydraulic barrier in land fills." *Engrg. Geol.* **60**. 245-252.
- Pusch, R. and Carlsson, T. (1985). "The physical state of pore water of Na Smectite used as barrier component." *Engrg. Geol.* **21**. 257-265.
- Rabalan, R. T, Turner, D.K, Bertetti, E.P. and Prikryl, J.D. (1998). "Key Geochemical Parameter Uranium VI sorption onto selected mineral surfaces. In: *Adsorption of metals by Geomedia In: Jeanne E. A (Ed.)*." Academic Press, San Diego.
- Rajasekaran, G and Rao, S.N. (1998). "X-Ray Diffraction and Microstructural studies of Lime-Marine Clay Reaction Products." *Geotech. Engrg.* **29(1)**. 1-27.

- Rajasekaran, G. and Rao S. N., (1995). "Cation exchange studies on a lime treated marine clay." *Geotech. Engrg.* **20 (2)**. 19 – 35.
- Rajasekaran, G. and Rao S. N. (1998). "Particle size analysis of lime-treated marine clays. *Amer. Society for Testing and Materials.*" **32 (1)**. 109 – 119.
- Rajasekaran, G. and Rao, Narasimha. (1996). "Lime migration studies in marine clays. *Oc. Engineering.*" **23(4)**. 325-355.
- Rajasekaran, G. and Rao, S. N. (1998). "Particle Size Analysis of Lime-Treated Marine Clays." *Geotech. Testing Journal.* **21(2)**. 109-119.
- Rajasekaran, G. and Rao, S.N. (2022). "Permeability characteristics of lime treated marine clay." *Oc. Engr.* **29**. 113-127.
- Rajasekaran. G and Rao. S.N. (1998). "X-Ray Diffraction and Microstructural studies of Lime-Marine Clay reaction products." *Geotech. Engr. J.* **29(1)**. 1-27.
- Rakmi Ab. Rahman, Murtedza Mohamed and Ghazally Ismail. (1984). "Lombong Tembaga Mamut di Sabah –I: Kualiti Air Sungai Sugut." *Sains Malaysiana.* **13(4)**. 439-450).
- Reddi, L. N, Ming, X, Hajra, M. G. and Lee, M. (2000). "Permeability Reduction of Soil Filters due to Physical Clogging." *J. of Geotech. and Geoenvironmental Eng.* **126(3)**. 236-246.
- Reid, J.M. and Brookes, A.H. (1999). "Investigations of lime stabilised contaminated material." *Engrg. Geol.* **53**. 217-231.
- Ronov A. B. and Yaroshevsky, A.A. (1972). "Earth's Crust Geochemistry. In R.W. Fairbridge (Ed)." *The Encyclopedia of Geochemistry and Environmental Sciene.* New York. 243-254.



- Rothstein, D, Thomas, J. J, Christensen, B. J. and Jennings, H. M. (2002). "Solubility behavior of Ca-, S-, Al-, and Si-bearing solid phases in Portland Cement pore solutions as a function of hydration time." *Cement & Concrete Res.* **32**. 1663 – 1671.
- Rowe, R. K, Armstrong, M. D. and Cullimore, D. R. (2000). "Particle Size and Clogging of Granular Media Permeated with Leachate." *J. of Geotech. and Geoenvironmental Eng.* **126(9)**. 775-786.
- Sallfors, G. B. and Pierce, J. (1985). "Determination of the permeability of a clay liner." *Engrg. Geol.* **21**. 225-228.
- Sarkar, B. (1986) Copper. In Sieler & Sigel (1988). "*Handbook on Toxicity of Inorganic compounds.*" Marcel Dekker, New York. 270-275.
- Sarsby, J. R. (2000). "*Environmental geotechnics.*" Thomas Telford. London, 296 – 297.
- Sastre. J , Sahuquillo. A, Vidal, M and . Rauret. G. (2002). "Determination of Cd, Cu, Pb and Zn in environmental samples: microwave-assisted total digestion versus aqua regia and nitric acid extraction." *Analytica Chimica Acta*, **462**. 59–72.
- Schomaker, N.B. (1980). "U.S. EPA program in hazardous waste landfill research. *Hazardous and industrial solid waste testing: Fourth Symposium, ASTM STP 886*. In Petros, Jr. .K. Lacy, W.J. and Conway, R.A. (Eds.)." American Society for Testing and Materials, Philadelphia. 251-262.
- Schomaker, N.B. (1986). "U.S.EPA Program in Hazardous Waste Landfill Research." *Hazardous and Industrial Solid Waste Testing: Fourth Symposium, ASTM STP 886*. J.K. Petros, Jr., W.J. Lacy, and R.A. Conway, (Eds). American Society for Testing and Materials, Philadelphia. 251-262.

- Sengupta, M. (1993). "*Environmental Impacts of Mining: Monitoring, Restoration, and Control.*" Lewis Publisher.USA. 267-258.
- Sharma, H. D. and Lewis, S. P. (1994). "*Waste Containment Systems, Waste Stabilization, and Landfills: Design and Evaluation.*" John Wiley & Sons, New York. 17, 308-309.
- Sharma, H. H. and Lewis, S. P. (1994). "*Waste Containment Systems, Waste Stabilization and Landfills.*" John Wiley & Sons. 374 - 381.
- Shi, L. and Xu, X. (2001). "Study of the effect of fly ash on desulfurization by lime." *Fuel*. 80. 1969 – 1973.
- Shively, W. Bishop. P., Gress, D and Brown, T. (1986). "Leaching tests of Heavy Metals Stabilized with Portland Cement." *Journal WPCF*. 58. 234-241.
- Siemens. (1989). *Handbook of Diffraktometer D 5000.*" Siemens AG, Federal Republic of Germany.
- Singh, U and Uehara, G. (1999). "Electrochemistry of the double layer: Principles and Applications to Soils. In Sparks, D. L (2<sup>nd</sup>.Ed.). *Soil Physical Chemistry.*" CRC Press LLC. 1-44.
- Soliman, N.N. (1990). "*Laboratory testing of lime fixed fly ash and FGD sludge. Geotechnics of Waste Fills – Teory and practice, ASTM STP 1077.*" Landva, A and Knowles, G.D., (Eds). ASTM. Philadelphia.
- Sponza, D. and Karaoglu, N. (2002). "Environmental geochemistry and pollution studies of Aliaga metal industry district." *Environmental International*. 27. 541 – 553.
- Sposito, G. (1984). "*The surface Chemistry of Soils.*" Oxford University Press: New York. 277.

- Stephan, D. Maleki, H. Knofel, D. Eber, B. and Hardtl, R. (1999). "Influence of Cr, Ni, and Zn on the properties of pure clinker phases Part I. C<sub>3</sub>S." *Cement & Concrete Res.* **29**. 545-552.
- Sunderman, Jr. F. W. (1988). Nickel. "*Handbook on Toxicity of Inorganic Compounds.*" Marcel Dekker, New York. 456-459.
- Sustainable urban Development Project-Sabah. (SUDP). (2000). "Working Paper: Landfill improvements." Kota Kinabalu City Hall.
- Tan, B.K and Yong, R.N. (1999). "Assessing the adsorption capabilities of clay soils via the batch equilibrium test." *Eleventh Asian Regional Conference on Soil Mechanics and Geotechnical Engineering*, Hot, et al., (Ed). Balkema, Rotterdam.
- Tan, K.H. (1993). "*Principles of Soil Chemistry* (2<sup>nd</sup>. Ed)." Marcel Dekker, Inc. 260-261.
- Theodoratos, P. Moirou, A, Xenidis, A. and Paspaliaris, I. (2000). "The use of municipal sewage sludge for the stabilization of soil contaminated by mining activities." *J. of Hazardous Materials.* **B77**. 177 – 191.
- Thevenin, G. Pera, J. (1999). "Interactions between lead and different binders. *Cement & Concrete Res.*" **29**. 1605 –1610.
- Thoulow, N. Jakobse, U.H. and Clark, B. (1996). "Composition of Alkali Silica Gel and Ettringite in Concrete Railroad Ties: SEM-EDX and X-ray Diffraction Analyses." *Cement & Concrete Res.* **26(2)**. 309-318.
- Transport Research Board. (1976). "*State of the Art: Lime Stabilization: Reactions, Properties, Design, Construction.*" Transportation Research Circular, Washington.

- Tsai, T. and Vesilind, P. A. (1998). "A new Landfill Liner and Reduce Ground-water Contamination from Heavy Metals." *J. of Environmental Eng.* **124**. 1061-1095.
- Tumidajski, P.J. and Thomson, M.L. (1994). "Influence of Cadmium on the Hydration of C3A." *Cement & Concrete Res.* **24(7)**. 1359-1372.
- Turekian, K. K and Wedepohl, K. (1961). "In Rosler, H.J & Lange, H. *Geochemical Tables* ." Elsevier Publishing Company, Amsterdam.
- USDA Soil conservation Service, Soil Survey Staff. (1975). "*Soil Taxonomy- A basic System of Soil Classification for Making and Interpreting Soil Surveys.*" *Agriculture Hand book 436, USDA, SCS.* U.S. Government printing Office, Washington.
- U. S. Environmental Protection Agency (USEPA). 1989a. *Decision-Maker's Guide to Solid Waste Management, EPA/530-SW89-072, USEPA, Washington..*
- U. S. Environmental Protection Agency (USEPA). 1989b. *Requirement for Hazardous Waste Landfill Design, EPA/625/4-89/022, USEPA.* Washington. 127.
- U. S. Environmental Protection Agency. (1989). "*Stabilization / Solidification of CERCLA and RCRA wastes*". *Physical Tests, Chemical Testing Procedures, Technology Screening, and Field Activities, Report EPA / 635/ 6 – 89/ 022.*
- U. S. Environmental Protection Agency. (1989b). "*Requirements for Hazardous for Waste Landfill Design, Construction, and Closure, Seminar Publication, EPA/625/44-89/02.*" USEPA, Cincinnati, Ohio, 1989.
- United Conference on Environment and Development. (2002). "World Summit on Sustainable Development Plan of Implementation, Johannesburg Summit." [http://www.joburgsummit2002.com/docs/plan\\_final.pdf](http://www.joburgsummit2002.com/docs/plan_final.pdf)

- Vail, J.G. (1952). "Soluble silicates. in Conner J.R. Chemical Fixation and Solidification of Hazardous Waster." Van Nostrand Reinhold. New York. 86 – 91.
- Vinogradov, A. P. (1962). In Rosler, H.J & Lange, H., 1972. "Geochemical Tables." Elsevier Publishing Company, Amsterdam.
- W. K. De Jongh. (1973). "X-ray fluorescence analyses applying theoretical matrix corrections." *Stainless steel X-ray spectrometry*. **2** (151).
- Wang, Y.M, Chen, T.C, Yeh, K.J & Shue, M. F., (2001). "Stabilization of an elevated heavy metal contaminated site." *J. of Hazardous Materials*. **B88**. 63-74.
- Webb, P.A. and Orr, C. (1997). "Analytical Methods in Fine Particle Technology." Micromeritics Instrument Corporation, Norcross, USA.
- Wiles, C.C., (1987). "A review of Solidification/Stabilization Technology." *J. of Hazardous Materials*. **14**. 5-21.
- Wood, P. A. (1992). "Remediation Methods for Contaminated Site. In Hester, R. E. and Harrison, R. M. (Ed.)." *Contaminated land and its reclamation*. Thomas Telford. Great Britian. 47-58.
- Yong, R.N. Tan, B.K. Bently, S.P. Thomas, H.R. and Yaacob, W.Z.W. (1999). "On competency assessment of clay soils form South Wale as landfill liners." *Quarterly J.l of Engrg Geol.* **32(10)**. 261-270.
- Yong, R.N. Yaacob, W.Z.W, Tan, Bently, S.P, Harris, C and Tan. B. K (2001). "Partitioning of heavy metals on soil samples from column test." *Engrg. Geol.* **60**. 307-322.
- Yong, R. N. (2000). "Geoenvironmental Engineering: Contaminated Soils, Pollutant Fate, and Mitigation." CRC Press, USA. 107-109.

- Yong, R. N. Elmonayeri, D. S, and Chong, T. S. (1985). "The effect of leaching on the integrity of a natural clay." *Engrg. Geol.* **21**. 279-299.
- Younger, J. S. (1988). "Natural and lime stabilised properties of Bandung clay. Geomechanics in Tropical Soils." *Proceedings of the second international conference on Geomechanics in Tropical Soils*, Singapore. 437-444.
- Yukselen, M. A and Alpaslan, B. (2001). "Leaching of metals from soil contaminated by mining activities." *J. of Hazardous Materials.* **B87**. 289-300.
- Zachara, J.M, Smith, S. C, McKinley, J.P. and Resch, C. T. (1993). "Cadmium sorption on specimen and soil smectites in sodium and calcium electrolytes." *Soil Sci. Soc. Am. Journal.* **57**. 1491.
- Zou, D.H. and Li, L.P. (1999). "Strengthening of Solidified Dilute Tailings Slurry. *J. of Geotech. and Geoenvironmental Eng.*" **125(1)**. 11-15.