EFFECT OF SURFACE PREPARATIONS ON THE ADHESION AND CORROSION BEHAVIOUR OF Zn-0.5Al COATING ON HIGH CARBON STEEL

CHAI TECK CHAW

A project report submitted in partial fulfilment of the requirements for the award of the degree of Master of Engineering (Materials Engineering)

> Faculty of Mechanical Engineering Universiti Teknologi Malaysia

> > SEPTEMBER 2014

For my dearest family, Prof Esah and friends

ACKNOWLEDGEMENT

First and foremost, I would like to express my deepest appreciation and gratitude towards my supervisor, Prof Dr Esah Hamzah for her continual supports, advices, guidance and friendship throughout this project. Without her, this project would not be as successful as it is today.

Thousand thanks also, to post doctoral fellow, Dr Saeed Farahany and technicians in the material science laboratory, Faculty of Mechanical Engineering for their technical support, friendliness and help during the laboratory work. Besides that, I also wish to thank my friends and coursemates who enlightened me with brilliant ideas during discussion.

Last but not least, I wish to thank my lovely parents for their continual moral and financial supports.

ABSTRACT

The corrosion of steel is one of the primary causes for premature failure. The ideal way to overcome this problem is to provide corrosion protection right at the time of manufacturing by hot-dip galvanizing. Conventionally, pure zinc (Zn) galvanizing provide a poor adhesion coating due to formation of brittle intermetallic of Fe–Zn; whilst single layer Zn–0.5 wt% Al alloy coating has a poor adhesion due to lack of suitable fluxes. This project is aimed to investigate the effects of various surface preparation treatments of high carbon steel plate on the adhesion and corrosion behaviour of Zn-0.5Al coating. Surface preparation prior to coating is thus crucial in order to obtain better coating adhesion and improve corrosion resistance. In this project, sand blasting and sand grinding were conducted together with chemical pretreatment on the steel substrate prior to Zn-0.5Al coating. The sand blasted, sand ground together with chemical pretreated steel samples were coated with Zn-0.5Al coating by dipping process. Adhesion property was examined on the coated samples by using pin on disc test. Electrochemical and salt spray tests were performed on the coated steels following the ASTM standard. Analysis has been carried out on the samples by using standard material characterization techniques namely optical microscopy, scanning electron microscopy and X-ray diffractometer. From this project, it was found that sand blasting followed by chemical pretreatment gave the best coating performance in terms of both adhesion and corrosion protection. This is due to the effect of high surface roughness from sand blasting which led to good mechanical interlocking effect for improved adhesion strength between Zn-0.5Al coating and steel substrate. Subsequently, corrosion protection also improved because Zn-0.5Al alloy coating did not peel off easily during corrosion tests. Sand ground galvanized steel gave the intermediate coating performance and conventional chemical pretreatment galvanized steel had a low coating performance in terms of both adhesion and corrosion protection. It can be concluded that surface roughness has significant effect on the coating adhesion and corrosion protection of galvanized coating.

ABSTRAK

Kakisan keluli adalah salah satu punca utama kegagalan pra-matang. Cara yang sesuai untuk mengatasi masalah ini adalah untuk memberi perlindungan kakisan semasa pencelupan panas penggalvanian. Sebelum ini, salutan zink (Zn) menyediakan lapisan lekatan yang lemah kerana pembentukan antara logam Fe-Zn yang rapuh; manakala lapisan Zn-0.5 %berat Al salutan aloi mempunyai lekatan yang lemah kerana kekurangan fluks yang sesuai. Projek ini bertujuan untuk mengkaji kesan pelbagai jenis rawatan permukaan pada keluli karbon tinggi ke atas lekatan dan kelakuan kakisan salutan Zn-0.5Al. Penyediaan permukaan sebelum salutan adalah penting untuk mendapatkan lapisan lekatan yang lebih baik dan meningkatkan perlindungan kakisan. Dalam projek ini, pembagasan pasir dan pengisaran pasir telah dijalankan bersama-sama dengan pra-rawatan kimia pada substrat keluli sebelum salutan celup panas Zn-0.5Al. Semua sampel kemudiannya disalut dengan salutan Zn-0.5Al melalui proses celupan panas. Kualiti lekatan telah diperiksa pada sampel bersalut dengan menggunakan mesin pin pada cakera. Ujian elektrokimia dan semburan garam telah dijalankan ke atas keluli bersalut berdasarkan standard ASTM. Analisis telah dijalankan ke atas sampel dengan menggunakan teknik pencirian bahan iaitu mikroskop optik, mikroskop pengimbasan elektron dan pembelau sinar-X. Daripada projek ini, telah didapati bahawa pembagasan pasir diikuti dengan pra-rawatan kimia memberikan prestasi salutan yang terbaik dari segi lekatan dan perlindungan kakisan. Ini adalah disebabkan oleh kesan kekasaran permukaan yang tinggi daripada pembagasan pasir yang membawa kepada kesan cengkaman mekanikal yang baik untuk kekuatan lekatan yang lebih baik antara lapisan Zn-0.5Al dan substrat keluli. Selepas itu, perlindungan kakisan juga bertambah baik kerana salutan aloi Zn-0.5Al tidak mudah menggelupas semasa ujian kakisan. Pengisaran pasir pada substrat keluli memberikan prestasi salutan yang sederhana dan konvensional pra-rawatan kimia keluli salutan Zn-0.5Al mempunyai prestasi salutan yang lemah dari segi lekatan dan kakisan perlindungan. Ia boleh disimpulkan bahawa kekasaran permukaan mempunyai kesan yang besar ke atas lekatan salutan dan perlindungan kakisan salutan tergalvani.

TABLE OF CONTENT

CHAPTER	
---------	--

1

2

TITLE

PAGE

TITI	LE PAGE	i
DEC	LARATION	ii
DED	DICATION	iii
ACK	NOWLEDGEMENT	iv
ABS	TRACT	V
ABS	TRAK	vi
TAB	LE OF CONTENTS	vii
LIST	Γ OF TABLES	xi
LIST	Γ OF FIGURES	xii
LIST	F OF SYMBOLS AND ABBREVIATIONS	xvii
LIST	Γ OF APPENDICES	xviii
INT	RODUCTION	1
1.1	Introduction	1
1.2	Problem Statement	2
1.3	Objective	3
1.4	Scopes of Study	3
LITI	ERATURE REVIEW	4
2.1	Introduction	4
2.2	Carbon Steel	4
	2.2.1 Low Carbon Steel	5
	2.2.2 Medium Carbon Steel	5
	2.2.3 High Carbon Steel	6
2.3	Steel Wire Ropes	6

	2.3.1	Types of Steel Wire Ropes 2.3.1.1 Six Strands	7 7
		2.3.1.2 Spiral Strands	8
		2.3.1.3 Multi Strands	9
	2.3.2	Adhesion strength study	9
		2.3.2.1 Pin on Disc Test	10
		2.3.2.2 Scratch Test	10
2.4	Corros	sion of Carbon Steel	11
2.5	Corros	sion of Zinc Coating	12
	2.5.1	Pitting Corrosion	13
	2.5.2	Water Corrosion	15
	2.5.3	Soil Corrosion	15
2.6	Corros	sion Control	15
	2.6.1	Protective Coatings	16
2.7	Corros	ion Testing and Monitoring	16
	2.7.1	Electrochemical Test	17
		2.7.1.1 Polarization Experiments	18
		2.7.1.2 Polarization Test Analytical Method	19
	2.7.2	Salt Spray Test	20
	2.7.3	Immersion Test	22
2.8	Hot Di	ip Galvanizing	23
	2.8.1	Pure Zinc Galvanizing	23
		2.8.1.1 Chemistry of Zinc Galvanizing	23
	2.8.2	Zn-Al Galvanizing	24
2.9	Pretrea	atment Process of Hot Dip Galvanizing	25
	2.9.1	Cleaning	26
	2.9.2	Degreasing and Pickling	26
	2.9.3	Fluxing	27
	2.9.4	Fluxing with NiCl ₂ -based Fluxes	27
2.10	Post G	alvanizing Treatments	29
2.11	Corros	sion Protection Mechanism for hot dip galvanizing	30
2.12	Differe	ent Types of Surface Preparation for Carbon Steel	31
	2.12.1	Degreasing	32
	2.12.2	Sand Blasting	33

		2.12.3 Chemical Treatment	34
3	RESI	EARCH METHODOLOGY	35
	3.1	Introduction	35
	3.2	Substrate Materials	36
	3.3	Zn-0.5Al dipping ingot preparation	37
	3.4	Thermal Analysis of the Cast Ingot	39
	3.5	Pretreatment Procedures	40
		3.5.1 Mechanical pretreatment by sand blasting	40
		3.5.2 Mechanical pretreatment by sand grinding	41
		3.5.3 Chemical pretreatment prior to hot dipping	42
	3.6	Hot Dip Galvanizing Process	43
	3.7	Corrosion test using Tafel polarization method	44
	3.8	Corrosion test using salt spray method	46
		3.8.1 Salt spray test sample preparations	47
	3.9	Materials Characterization	48
		3.9.1 Optical Microscopy	48
		3.9.2 Scanning Electron Microscopy (SEM)	49
		3.9.3 X-Ray Diffraction Spectrometer (XRD)	50
	3.10	Sample Preparation for Micrograph Analysis	51
	3.11	Vickers Hardness Measurement	52
	3.12	Surface Roughness Measurement	52
	3.13	Adhesion Strength Measurement	54
4	RESU	JLTS AND DISCUSSION	56
	4.1	Introduction	56
	4.2	Analysis of Surface after Pretreatment (before coating)	57
		4.2.1 Sample Hardness	58
	4.3	Thermal Analysis on the Zn-0.5Al ingot	58
	4.4	Hot Dip Galvanizing Coating	59
		4.4.1 Visual Inspection	59
		4.4.2 Coating Thickness	62
	4.5	Corrosion Tests	64
		4.5.1 Electrochemical Test	64

		4.5.2	Salt Spray Corrosion Test	71
		4.5.3	Salt Spray Corrosion Rate Measurement	75
		4.5.4	XRD Characterization on corrosion product	77
	4.6	Pin on	disc test	79
5	CON	CLUSI	ONS AND RECOMMENDATIONS FOR	82
	FUT	URE WO	ORK	
	5.1	Conclu	isions	82
	5.2	Recon	nmendations for Future Work	83
REF	TERENO	FS		84
KL1				04
APF	APPENDIX A – C		87	

LIST OF TABLES

TABLE	TITLE	PAGE
2.1	Variety of Unites, K value	22
2.2	Different composition of four sets of fluxes used	28
3.1	Sand blasting parameters	41
3.2	Chemical pretreatment procedures and parameters	42
3.3	Parameters for salt spray based on ASTM B117	47
3.4	Pin on Disc Parameter	54
4.1	Hardness value of as received steel substrate	58
4.2	Electrochemical data for untreated sample	66
4.3	Electrochemical data for sand ground sample	67
4.4	Electrochemical data for sand blasted sample	68
4.5	Summary of Tafel Polarization Test	70
4.6	Corrosion rate calculated from weight loss (salt spray test)	76

LIST OF FIGURES

FIGURE	TITLE	PAGE
2.0	Fe-C Phase Diagram	5
2.1	Types of wire rope constructions	7
2.2	Types of wire rope strands	8
2.3	Variation on corrosion rate with pH	13
2.4	Pitting corrosion on galvanized coating	13
2.5	Schematic of pitting corrosion	14
2.6	Electrochemical cell for polarization experiments	19
	(a) Schematic drawing of the corrosion cell set up	
	(b) Actual drawing of the corrosion cell set up	
2.7	Tafel Polarization set up	20
2.8	(a) Corrosion Process Showing Anodic and Cathodic	20
	Current Components (b) Classic Tafel Analysis	
2.9	Structure of a galvanized coating on a steel	24
2.10	Five years corrosion test on early Zn-Al alloys,	25
	effect of Al content on corrosion performance of	
	Zn-Al alloy coatings after exposure in various atmosphere	S
2.11	Average thickness of delta (δ) and zeta (ζ) interlayers of	29
	galvanized specimens pretreated by different fluxes and	
	hot dipped for various durations	

2.12	Corrosion mechanism under a water droplet with	31
	Zn coating protects the steel	
3.1	Overall experimental procedures	36
3.2	High carbon steel substrate (a) Steel wire rope and	37
	(b) Flat steel plate after cutting the wire rope.	
3.3	Pieces of Zn raw materials	37
3.4	High Temperature Furnace	38
3.5	Stirring of the dipping ingot	38
3.6	As-cast Zn-0.5% Al ingot	38
3.7	Thermal analysis set up with ceramic mould, metallic	39
	mould and thermocouple	
3.8	Sand blasting machine	40
3.9	Steel substrate surface was ground with SiC #500 paper	41
	in uni-direction to create a micro surface roughness.	
3.10	Process of chemical pretreatment (a \rightarrow b \rightarrow c \rightarrow d \rightarrow e \rightarrow f)	43
3.11	Dipping of steel sample into the Zn-Al ingot	44
3.12	Quenching of hot-dipped coated steel into water	44
3.13	Cold mounting for Tafel Polarization Test	45
3.14	Tafel polarization test setup	45
3.15	Salt Spray Chamber: (a) Schematic of the sample set up	46
	in the chamber and (b) The salt spray chamber	
3.16	Salt Spray Test Sample Preparation	47
3.17	X-Ray Diffractometer	50
3.18	Grinding Machine	51
3.19	Auto Grinding Machine	52
3.20	Vickers Hardness Tester	52
3.21	Surface roughness Profiler	53

3.22	Atomic Force Microscopy	53
3.23	Pin on disc machine (a) specimen set up;	54
	(b) analytical computer	
3.24	Modified sample for pin on disc test (a) CAD drawing	55
	(b) Actual pin on disc sample	
4.1	Summary of Analysis of Surface before coating	57
4.2	Cooling curve plotted from steel and ceramic moulds	59
4.3	Visual Inspection on Coated Samples	60
4.4	Scanning Electron Micrographs of coated samples	61
	after hot-dip galvanizing process.	
4.5	Scanning electron micrograph showing cross section of	62
	untreated sample	
4.6	Scanning electron micrograph showing cross section of	63
	sand ground sample	
4.7	Scanning electron micrograph showing cross section of	63
	sand blasted sample)	
4.8	Optical micrograph showing cross section of sand blasted	63
	sample with mechanical interlocking	
4.9	Scanning electron micrograph of untreated sample	65
	showing uncoated area in the coating exposed the uncoated	
	steel	
4.10	Tafel Plot for untreated sample	65
4.11	Scanning electron micrograph of sand ground	66
	sample showing uncoated area exposed the uncoated	
	steel as cathodic region.	
4.12	Tafel Plot for sand ground sample	67
4.13	Tafel Plot for sand blasted sample	68

4.14	Scanning electron micrographs of sand blasted sample	69
	(a) Before electrochemical test, (b) after electrochemical	
	Test	
4.15	Scanning electron micrograph of sand blasted sample	69
	showing magnification 250x on the pitting	
	for corroded area	
4.16	Combined Tafel plot for three samples	70
4.17	Bar chart for corrosion rate from Tafel plot	71
4.18	Samples before salt spray test	71
4.19	Samples after salt spray test	72
4.20	Scanning electron micrograph of untreated sample	72
	showing pitting corrosion.	
4.21	Scanning electron micrograph of sand ground sample	73
	showing pitting corrosion.	
4.22	Scanning electron micrograph of sand blasted sample	73
	showing pitting corrosion.	
4.23	Visual inspection showing heavy corrosion for	74
	uncoated steel	
4.24	Field emission scanning electron micrograph	74
	showing uniform corrosion on uncoated steel.	
4.25	Weight loss after seven days salt spray test	75
4.26	Salt spray corrosion rate obtained for each sample	77
	after seven days	
4.27	Powder form of corrosion product, from uncoated	77
	steel (left)	
	from Zn-Al-coated steel (right)	
4.28	XRD result for Zn-Al coated steel	78

4.29	XRD result for uncoated steel	79
4.30	Pin on disc test showing wear results for all samples	80
4.31	Scanning electron micrograph showing wear occurred	81
	on untreated sample	
4.32	Scanning electron micrograph showing wear occurred	81
	on sand ground sample.	
4.33	Scanning electron micrograph showing wear occurred	81
	on sand blasted sample.	

LIST OF SYMBOLS AND ABBREVIATIONS

Zn	- Zinc
Al	- Aluminium
Cl	- Chlorine
Fe	- Steel
%	- Percentage
°C	- Degree Celsius
m	- Meter
g	- Gram
1	- Litre
R _a	- Average roughness
SEM	- Scanning Electrode Microscope
FESEM	- Field Emission Scanning Electron Microscope
EDX	- Energy Dispersive X- Ray Analysis
XRD	- X-Ray Diffractometer
H_2O	- Water
H_2SO_4	- Sulphuric Acid
HNO ₃	- Nitric Acid
HCl	- Hydrogen Chloride (Chloride Acid)
NH ₄ Cl	- Ammonium Chloride
$ZnCl_2$	- Zinc Chloride

LIST OF APPENDICES

APPENDIX	TITLE	PAGE
А	Pitting Corrosion FESEM Images (Untreated Sample)	87
В	Pitting Corrosion FESEM Images (Sand ground Sample)	88
С	Pitting Corrosion FESEM Images (Sand blasted Sample)	89

CHAPTER 1

INTRODUCTION

1.1 Introduction

High carbon steel is used widely in the applications which require high strength property especially in the offshore corrosive environment, uncoated high carbon steel will be rendered useless due to corrosion failure. Hence, in order to solve this problem, galvanized coating was applied onto high carbon steel for barrier and cathodic protection. However, the effectiveness of the coating depends on the environment condition to which the Zinc (Zn) coating is exposed. The corrosion resistance ability of Zn in mild chloride environment is good but it has inferior resistance in aggressive chloride environment and thereby reduces the service life of the high carbon steel (Ahmad Z, 2006).

Furthermore, ductility of pure Zn coating is found to be inferior due to the presence of brittle Fe-Zn phases such as gamma, gamma1, delta and zeta. Therefore, additional of Aluminium (Al) is proposed by researchers in order to impede the formation of the brittle interlayer phases of Fe-Zn. Adhesion is another very important property for coating process. In this study, galvanized coating was investigated in terms of their adhesion and corrosion performance on high carbon steel.

1.2 Problem Statement

Over the pass decades, offshore industry has been developing extensively due to the depletion of the onshore resources and the possibly huge profits promised by the oil and gas offshore resources. Hence, vessels, offshore plants and cranes are widely used for the extractions and production of the oil and gas resources. In order to maintain the stability and forestall the free movement of the production system on the seawater, wire ropes are used as the major part in the mooring system. With recent exploration in the subsea with depths beyond 1500 meters, wire ropes showed a few limitations towards both operation system and the environment. One of the most detrimental factors which cause the steel wire to fail in the subsea environment is corrosion problem. Several conditions such as salt (and chlorine) content, dissolved oxygen, wind and wave flow erosion and temperature difference will result in premature failure of the mooring system and possible financial lost for the production system (Chaplin, C.R., 1999).

To solve the problem, galvanizing on steel wire ropes has been used extensively in the offshore industry over the centuries due to its excellent corrosion resistance. Generally galvanizing on steel wire ropes can provide two types of corrosion protection, i.e., cathodic protection and barrier protection. However, in some circumstances, pure zinc galvanizing has proven to have corrosion problem as well despite the excellent two types of protection given. The proposed research project is therefore aimed at investigating the corrosion behavior of galvanized steel ropes for offshore mooring system and introducing a new set of parameters with Zn-Al bath for better corrosion resistance.

1.3 Objective of the Research

A study on the possible corrosion mechanisms and how to minimize the corrosion actions are required. Therefore, the main objective of the research is to investigate the effects of various surface preparation treatments of high carbon steel plate on the adhesion and corrosion behaviour of Zn-Al coating.

1.4 Scope of Study

- 1. Selection of various surface treatments suitable for high carbon steel plates.
- 2. Hot dip galvanizing process of Zn-Al on surface treated high carbon steel plates.
- Microstructural evolutions will be characterized by optical microscopy, XRD, AFM and FESEM.
- 4. Pin on disc test will be performed to determine adhesion property.
- 5. The corrosion resistance will be investigated by electrochemical test and salt spray test.

REFERENCES

- Ahmad, Z. (2006). Chapter 7 Coatings. Principles of Corrosion Engineering and Corrosion Control. Oxford, Butterworth-Heinemann: 382-437.
- Beduque A, Dou Z, Xu.R (2009) Electrochemical studies for Aluminum Electrilytic Capacitor Applications: Corrosion Analysis of Aluminium in Ethylene Glycol-Based Electrolytes, Electronic Components Assoc., Inc
- Bicao, P., W. Jianhua, et al. (2008). "Effects of zinc bath temperature on the coatings of hot-dip galvanizing." Surface and Coatings Technology 202(9): 1785-1788.
- Chaplin, C. R. (1999). "Torsional failure of a wire rope mooring line during installation in deep water" Engineering Failure Analysis 6(2): 67-82.
- Davis, G. D. and J. D. Venables (2002). Chapter 21 Surface treatments of metal adherends. Adhesion Science and Engineering. M. Chaudhury and A. V. Pocius. Amsterdam, Elsevier Science B.V.: 947-1008.
- Dutta, M., A. K. Halder, et al. (2010). "Morphology and properties of hot dip Zn–Mg and Zn–Mg–Al alloy coatings on steel sheet." Surface and Coatings Technology 205(7): 2578-2584.
- E.M. Bortoleto et.al (2012). "Experimental and Numerical Analysis of Dry Contact in the Pin on Disc Test"
- Ebnesajjad, S. (2011). 2 Introduction to Surface Preparation and Adhesion. Handbook of Adhesives and Surface Preparation. S. Ebnesajjad. Oxford, William Andrew Publishing: 15-18.
- Ebnesajjad, S. (2011). 6 Surface Preparation of Metals. Handbook of Adhesives and Surface Preparation. S. Ebnesajjad. Oxford, William Andrew Publishing: 83-106.
- Feliu Jr, S. and V. Barranco (2003). "XPS study of the surface chemistry of conventional hot-dip galvanised pure Zn, galvanneal and Zn–Al alloy coatings on steel." Acta Materialia 51(18): 5413-5424.

J.R. Davis, Davis et al. (2000) "Corrosion Understanding the Basics"

- Klaus Feyrer (2007) "Wire Ropes Tension, Endurance, Reliability" Wire Ropes, Elements and Definitions: 1-8.
- Li, B., A. Dong, et al. (2012). "Investigation of the corrosion behaviors of continuously hot-dip galvanizing Zn–Mg coating." Surface and Coatings Technology 206(19–20): 3989-3999.
- Li, Y. (2001). "Formation of nano-crystalline corrosion products on Zn–Al alloy coating exposed to seawater." Corrosion Science 43(9): 1793-1800.
- Liu, H., F. Li, et al. (2011). "Characterization of hot-dip galvanized coating on dual phase steels." Surface and Coatings Technology 205(11): 3535-3539.
- Manna, M. (2011). "Effect of fluxing chemical: An option for Zn–5wt.%Al alloy coating on wire surface by single hot dip process." Surface and Coatings Technology 205(12): 3716-3721.
- Marder, A. R. (2000). "The metallurgy of zinc-coated steel." Progress in Materials Science **45**(3): 191-271.
- Penney, D. J., J. H. Sullivan, et al. (2007). "Investigation into the effects of metallic coating thickness on the corrosion properties of Zn–Al alloy galvanizing coatings." Corrosion Science 49(3): 1321-1339.
- Peter Maass, Peter Peissker (2011) "Handbook of Hot-dip Galvanization"
- Renpu, W. (2011). "Chapter 11 Oil and Gas Well Corrosion and Corrosion Prevention". Advanced Well Completion Engineering (Third Edition), Gulf Professional Publishing: 617-700.
- Reumont, G., J. B. Vogt, et al. (2001). "The effects of an Fe–Zn intermetalliccontaining coating on the stress corrosion cracking behavior of a hot-dip galvanized steel." Surface and Coatings Technology 139(2–3): 265-271.
- Seré, P. R., M. Zapponi, et al. (1998). "Comparative corrosion behaviour of 55Aluminium–zinc alloy and zinc hot-dip coatings deposited on low carbon steel substrates" Corrosion Science 40(10): 1711-1723.
- Shih, H. C., J. W. Hsu, et al. (2002). "The lifetime assessment of hot-dip 5% Al–Zn coatings in chloride environments" Surface and Coatings Technology 150(1): 70-75.
- Smith F.W., Hashemi J., (2006), Foundations of Materials Science and Engineering (4th Ed), Mc Graw Hill International

- Smith, W. J. and F. E. Goodwin (2010). 4.06 Hot Dipped Coatings. Shreir's Corrosion. J. A. R. Editor-in-Chief: Tony. Oxford, Elsevier: 2556-2576.
- Tachibana, K., Y. Morinaga, et al. (2007). "Hot dip fine Zn and Zn–Al alloy double coating for corrosion resistance at coastal area" Corrosion Science 49(1): 149-157.
- V.S Sastri, Edward Ghali et al. (2007) "Corrosion Prevention and Protection Practical Solutions"
- William D. Callister, Jr. (2007) "Callister Materials Science and Engineering An Introduction 7th edition"
- Xie, Y. and H. M. Hawthorne (2002). "Effect of contact geometry on the failure modes of thin coatings in the scratch adhesion test." Surface and Coatings Technology 155(2–3): 121-129.
- Yadav, A. P., A. Nishikata, et al. (2004). "Degradation mechanism of galvanized steel in wet–dry cyclic environment containing chloride ions" Corrosion Science 46(2): 361-376.
- Yuttanant Boonyongmaneerat et. al (2010). "Effect of NiCl₂-Based Fluxes on Interfacial Layer Formation of Hot Dip galvanized Steels." Journal of Iron and Steel Research, International, 17(8):74-78
- Zhang, G. A., N. Yu, et al. (2014). "Galvanic corrosion behavior of deposit-covered and uncovered carbon steel." Corrosion Science 86(0): 202-212.