HOLMIUM IONS-ACTIVATED ZINC-SULFO-BORO-PHOSPHATE NANOCOMPOSITES WITH SILVER NANOPARTICLES SENSITIZATION

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ABSTRACT

Rare-earth ions (REIs)-doped oxide glasses and glass-ceramics (GCs) became promising for various photonic applications. However, the inherent small emission cross-section of the REIs-doped systems for practical applications need substantial enhancement. Despite some studies on holmium ion (Ho3+) luminescence from different hosts, the radiative properties of Ho³⁺ in zinc-sulfo-boro-phosphate GCs for the miniaturized and inexpensive lasers development remains deficient. In addition, the lasing potency of Ho³⁺-doped phosphate-based GC nanocomposites (GCNCs) with silver nanoparticles (Ag NPs) sensitization has not widely been explored. Thus, the structural, microstructural, impedance, optical and radiative properties of some Ag NPs and Ho³⁺ co-doped zinc-sulfo-boro-phosphate GCNCs were evaluated. Three series of samples with the composition of $(40-x)P_2O_5-30B_2O_3-30ZnSO_4-xHo_2O_3$. where x = 0.0, 0.2, 0.4, 0.5, 0.6, 0.8 and 1.0 mol%; $(39.5-y)P_2O_5-30B_2O_3-30ZnSO_4-$ 0.5Ho₂O₃-yAg nanopowder, where y = 0.6, 0.7, 0.8 and 0.9 mol%; and $(39.5-z)P_2O_5 30B_2O_3-30ZnSO_4-0.5Ho_2O_3-zAgCl$, where z = 0.6, 0.7, 0.8 and 0.9 mol% were prepared using melt-quenching method. Structural characteristics of the samples were determined using X-ray diffraction (XRD), Fourier transform infrared (FTIR), Raman, energy dispersive X-ray (EDX), X-ray photoelectron spectroscopy (XPS) and density measurements. Microstructures of the samples were analyzed using high-resolution transmission electron microscope (HRTEM) and impedance spectroscopy (IS). Optical properties of the samples were measured using ultraviolet-visible-near infrared (UV-Vis-NIR) and photoluminescence (PL) spectroscopy. The XRD analyses of the as-quenched samples verified their GC nature. The observed increase and decrease in the samples density was attributed to the formation of more bridging oxygen (BO) and non-bridging oxygen (NBO), respectively. The density results of these GCs and GCNCs were supported by the FTIR, Raman and XPS spectral data analyses. The HRTEM images reconfirmed the GC nature of the samples and the existence of the Ag NPs within the network structure. The optical energy band gap, refractive index and Urbach energy were calculated from the UV absorption spectra to get the information about the local structural surroundings. The GC doped with 0.5 mol% of Ho₂O₃ exhibited the highest intensity of the red and green PL emissions. Furthermore, the GCNC doped with 0.8 mol% of Ag NPs (mean diameter of 20 nm) revealed the optimum PL intensity enhancement and strongest LSPR absorption band. The obtained larger values of the fluorescence branching ratio and emission cross-section compared to the existing state-of-the-art reports indicated the benefits of the studied samples for the construction of green and red wavelength lasers. A correlation between structural and optical properties was also established for the first time. The studied Ag NPs and Ho³⁺ co-doped phosphate-based GCNCs were asserted to be potential for the efficient photonic devices advancement. It is concluded that via the systematic composition optimization, these new types of GCNCs with customized lasing potency can be achieved.

ABSTRAK

Kaca oksida dan kaca seramik (GC) terdop ion nadir bumi (REI) menjadi sangat berpotensi untuk pelbagai aplikasi fotonik. Walau bagaimanapun, keratan rentas pancaran yang kecil dalam sistem terdop-REI untuk kegunaan praktikal memerlukan peningkatan yang besar. Walaupun terdapat beberapa kajian mengenai pendarcahaya ion holmium (Ho³⁺) daripada hos yang berbeza, kajian terhadap sifat pancaran Ho³⁺ dalam sistem GC zink-sulfo-boro-fosfat untuk pembangunan laser bersaiz mini dan murah masih kurang. Tambahan pula, keupayaan untuk laser komposit nano GC berasaskan fosfat terdop Ho³⁺ dengan pemekaan zarah nano perak (Ag NP) masih belum diterokai sepenuhnya. Oleh itu, struktur, struktur mikro, impedans, sifat-sifat optik dan pancaran beberapa GCNC zink-sulfo-boro-fosfat ko-dop Ag NP dan Ho³⁺ telah dinilai. Tiga siri sampel dengan komposisi kimia (40-x)P₂O₅ -30B₂O₃-30ZnSO₄-*x*Ho₂O₃ di mana *x* = 0.0, 0.2, 0.4, 0.5, 0.6, 0.8 dan 1.0 mol%, (39.5-*y*)P₂O₅- $30B_2O_3-30ZnSO_4-0.5Ho_2O_3-vAg$ serbuk nano di mana v = 0.6, 0.7, 0.8 dan 0.9 mol%, dan $(39.5-z)P_2O_5-30B_2O_3-30ZnSO_4-0.5Ho_2O_3-zAgCl di mana z = 0.6, 0.7, 0.8 and$ 0.9 mol% telah disediakan dengan menggunakan kaedah lindap-kejut leburan. Ciriciri struktur sampel telah ditentukan menggunakan pembelauan sinar-X (XRD), spektroskopi inframerah transformasi Fourier (FTIR), Raman, sinar-X sebaran tenaga (EDX), fotoelektron sinar-X (XPS) dan pengukuran ketumpatan. Struktur mikro sampel telah dianalisa dengan menggunakan mikroskop elektron penghantaran resolusi tinggi (HRTEM) dan spektroskopi impedans (IS). Sifat-sifat optik sampel telah diukur dengan menggunakan spektroskopi ultra ungu-cahaya nampakinframerah hampir (UV-Vis-NIR) dan kefotopendarcahayaan (PL). Analisis XRD terhadap sampel lindap-kejut yang terhasil mengesahkan sifat semula jadi GC. Peningkatan dan penurunan ketumpatan sampel yang dicerap masing-masing dikaitkan kepada pembentukan lebih banyak oksigen berangkai (BO) dan oksigen tak berangkai (NBO). Keputusan ketumpatan GC dan GCNC ini disokong oleh hasil analisis data spektra FTIR, Raman dan XPS. Imej HRTEM mengesahkan sifat semulajadi GC sampel dan kewujudan Ag NP dalam struktur rangkaian. Jurang jalur tenaga optik, indeks biasan dan tenaga Urbach dikira dari spektra penyerapan UV untuk mendapatkan maklumat mengenai persekitaran struktur setempat. Sampel GC didop dengan 0.5 mol% Ho₂O₃ menunjukkan keamatan pancaran PL merah dan hijau vang tertinggi. Seterusnya, GCNC didop dengan 0.8 mol% Ag NP dengan diameter min 20 nm, menunjukkan peningkatan keamatan PL yang optimum dan jalur penyerapan LSPR yang kuat. Nilai nisbah cabang pendarfluor dan keratan rentas pancaran yang diperolehi adalah lebih tinggi berbanding dengan nilai terkini yang telah dilaporkan, menunjukkan keberkesanan sampel yang dikaji untuk pembangunan laser panjang gelombang hijau dan merah. Korelasi antara sifat-sifat struktur dan optik juga telah diperolehi untuk pertama kalinya dalam kajian ini. GCNC berasaskan fosfat ko-dop Ag NP dan Ho³⁺ yang dikaji adalah sangat berpotensi untuk kemajuan peranti fotonik yang cekap. Adalah disimpulkan bahawa melalui keadah pengoptimuman komposisi yang sistematik, satu jenis GCNC yang baharu, sesuai dengan keupayaan laser yang dikehendaki boleh dicapai.

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LIST OF SYMBOLS

А	-	Absorbance
α	-	Absorption coefficient
σ_{Ac}	-	AC conductivity
Å	-	Angstrom
A _{abs}	-	Area under the absorption peak
Mav	-	Average molecular weight
NA	-	Avogadro's number
В	-	Boron
B_2O_3	-	Boron oxide
f_{cal}	-	Calculated oscillator strength
С	-	Carbon
cm	-	Centimeter
e	-	Charge of electron
ϵ^*	-	Complex dielectric constant
Z^*	-	Complex impedance
°C	-	Degree Celsius
ρ	-	Density (Archimedes principle)
D _T	-	Density (Theoretical)
tanδ	-	Dielectric loss
θ	-	Diffracted angle of the X-Ray beam
f_{ed}	-	Electric dipole oscillator strength
S _{ed}	-	Electric dipole transition strength
J	-	Excited state's total angular momentum
ψ̈́	-	Excited state's wave function
fexp	-	Experimental oscillator strength
F	-	Farad
Fs	-	Field strength
β_R	-	Fluorescence branching ratio
f	-	Frequency

g	-	Gram		
J	-	Ground state's total angular momentum		
ψ	-	Ground state's wave function		
Hz	-	Hertz		
N	-	Ho ³⁺ concentration per cm ³		
Ho ₂ O ₃	-	Holmium oxide		
Но	-	Holmuim		
Ho ³⁺	-	Holmuim ion/s		
ε″	-	Imaginary part of the complex dielectric constant		
λ_p	-	Intensity of the PL peak		
<i>r</i> _i	-	Inter-nuclear distance		
Ω_{j}	-	J-O parameters and $j = 2,4$ and 6		
Tj	-	Judd's notation of Intensity parameters and $j = 2,4$ and 6		
k	-	Kilo (10 ³)		
f_{md}	-	Magnetic dipole oscillator strength		
S_{md}	-	Magnetic dipole transition strength		
m _e	-	Mass of electron		
Uj	-	Matrix elements and $j = 2,4$ and 6		
М	-	Mega (10 ⁶)		
m	-	Meter		
μ	-	Micro		
$\varepsilon(v)$	-	Molar extinction coefficient		
α_m	-	Molar polarizability		
R_m	-	Molar refractivity		
V_m	-	Molar volume		
x	-	Mole fraction of the component in the composition		
n	-	Nano (10 ⁻⁹)		
q	-	Number of transitions in the oscillator strength fitting		
Ω	-	Ohm		
Eg	-	Optical band gap energy		
0	-	Oxygen		
H ₃ PO ₄	-	Phosphoric acid		
Р	-	Phosphorus		

P_2O_5	-	Phosphorus pentoxide		
p	-	Pico (10^{-12})		
h	-	Planck constant		
r _p	-	Polaron radius		
φ	-	Porosity		
$ au_R$	-	Radiative lifetime		
A_R	-	Radiative transition probability		
C _{RE}	-	Rare-earth ion concentration		
ε′	-	Real part of the complex dielectric constant		
п	-	Refractive Index		
δ_{rms}	-	Root-mean-square deviation		
Sγ	-	Series where γ is 1 or 2 or 3		
Ag	-	Silver		
AgCl	-	Silver Chloride		
χ	-	Spectroscopic quality factor		
c	-	Speed of light		
$\sigma_{\mathtt{p}}^{\mathtt{E}}$	-	Stimulated emission cross-section		
S	-	Sulfur		
$\Delta\lambda_{e\!f\!f}$	-	The effective emission bandwidth		
π	-	The mathematical constant (3.14)		
m _t	-	The tangent slope value for the absorption edge.		
d	-	Thickness		
J	-	Total angular momentum		
A_T	-	Total radiative transition probability		
ΔΕ	-	Urbach energy		
λ	-	Wavelength		
Ip	-	Wavelength of the PL peak		
v	-	Wavenumber		
ct	-	Y-axis intercept value by the absorption edge tangent		
Zn	-	Zinc		
ZnSO ₄	-	Zinc sulfate		

LIST OF ABBREVIATIONS

A.U.	-	Absorption Unit
Aq.	-	Aqueous
a.u.	-	Arbitrary unit
BO	-	Bridging Oxygen
CPE	-	Constant Phase Element
CR	-	Cross Relaxation
deg.	-	Degree
eV	-	Electron Volt
EDX	-	Energy Dispersive X-Ray
ET	-	Energy Transfer
Eq.	-	Equation
FTIR	-	Fourier Transform Infrared
FWHM	-	Full Width at Half Maximum
GCNCs	-	Glass-Ceramic Nanocomposites
GCs	-	Glass-Ceramics
HRTEM	-	High Resolution Transmission Electron Microscopy
IS	-	Impedance Spectroscopy
LCR	-	Inductance Capacitance Resistance
IR	-	Infrared
IFFT	-	Inverse Fast Fourier Transform
J-O	-	Judd-Ofelt
LFP	-	Lattice Fringe Profile
LFE	-	Local Field Enhancement
LSPR	-	Localized Surface Plasmon Resonance
MR	-	Multi-Phonon Relaxations
NCs	-	Nanocomposites
NPs	-	Nanoparticles
NIR	-	Near Infrared
NBO	-	Non-Bridging Oxygen
NR	-	Non-Radiative

PL	-	Photoluminescence
REIs	-	Rare Earth Ions
X–R	-	Reactance-Resistance
Ref.	-	Reference
RQ	-	Resistance- Constant phase element
SAED	-	Selected Area Electron Diffraction
SPR	-	Surface Plasmon Resonance
TEM	-	Transmission Electron Microscopy
UV	-	Ultraviolet
Vis	-	Visible
XRD	-	X-Ray Diffraction
XPS	-	X-Ray Photoelectron Spectroscopy

LIST OF APPENDICES

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CHAPTER 1

INTRODUCTION

1.1 Research Background

The main feature that distinguishes rare-earth elements from other elements is an incompletely filled 4f subshell that is screened by the completely filled outer 5s and 5p subshells [1]. Thus, when they are incorporated into any host material, rare-earth ions (REIs) gives rise to major absorption and emission actions responsible for a wide range of applications including lasers, light emitting diodes (LEDs), and amplifiers [2– 4]. Absorption and emission cross-section of REIs are critical parameters deciding the REI's lasing potency. Therefore, the efforts have continually been made to improve the stimulated emission cross-section of REIs via the selection of appropriate host materials, suitable modifiers, and sensitizers (e.g., metal nanoparticles and nanostructures) [5–7].

Among the ternary and quaternary oxide hosts, the phosphate-based glass and glass-ceramic (GC) systems are potential because of their high thermal expansion coefficient, low phonon energy, large intake of REIs and low glass transition temperature [7–10]. Phosphate-based GCs have been acknowledged for better REIs luminescence features than glasses [11,12]. However, phosphate-based systems tend to absorb moisture, leading to the inclusion of OH^- impurities in the network structure which induce the undesirable non-radiative mechanisms in the system [13]. To overcome this limitation, the incorporating of some network modifiers was proved to be prospective to improve the chemical durability of phosphate glasses.

The combination of the phosphate and borate units enhance both the glass forming ability and the chemical durability of phosphate glasses by cross-linking phosphate chains [14]. The chemical durability and structural stability of phosphate glass can further be improved by adding Zn^{2+} ions in the network structures [15,16]. It

has also been argued that the interactions between the phosphate and sulfate ions can enhance the chemical durability and create a good environment for the intake of large number of REIs, allowing the fabrication of the miniaturized lasers with the improved features [17–19]. It is worth mentioning that the zinc-sulfo-boro-phosphate composition is a new promising host for the REIs doping.

The holmium ions (Ho³⁺) among various REIs have been used in diverse technologies including lasers [20]. They exhibit unique emissions in the ultraviolet (UV), visible (Vis), and infrared (IR) regions, however, the intensities of these emissions still need to be improved for glass lasers and nanophotonic devices [6]. To evaluate the Ho³⁺ radiative properties, calculations based on Judd-Ofelt theory have been widely used [2,3,5,7,21,22] over last few decades. Through these calculations, some essential parameters can be estimated including the stimulated emission crosssection (σ_P^E). However, Ho³⁺ alike other REIs has small emission cross-section in amorphous hosts, causing high laser threshold and low gain [8]. Therefore, more studies to improve the Ho³⁺ emission cross-section are required. In this regard, some strategies have been used in order to enhanced the σ_P^E of Ho³⁺ such as the rightly selection of the host material, the co-doped with another REIs [5], and insertion of metallic nanoparticles (NPs) [7]. However, the exploration of the lasing potency of Ho³⁺ inside Ag NPs-sensitized phosphate-based GCNCs remain deficient.

Recently, a combination of the metal NPs with REIs in various host matrices has been proven to be advantageous for achieving the significant enhancement in the emission cross-section of the REIs. The Ag NPs being the common plasmonic metamaterial with abundance, strong biocompatibility, high chemical stability and resistant against oxidation has been used as the sensitizing agent in many systems to amplify the REIs lasing action [7,23–25]. The localized surface plasmon resonance (LSPR) effect of metal NPs has been demonstrated to be responsible for such significant enhancement of the optical properties [26,27]. Ag NPs size dependent improvement of LSPR field plays a vital role in enhancing the REIs photoluminescence (PL) emission intensity. The size can be controlled by means of altering the temperature and duration of thermal processing. In regard to this fact, controlling and exploring all the stages of Ag NPs formation including the starting

stage is required. Furthermore, the relative permittivity between the host material and the surface of the NP (dielectric-metal interface) play a significant role to achieve an enhancement in the PL intensity through the LSPR mediation.

The impedance spectroscopy has recently been proven to be a powerful tool to evaluate the complex permittivity and the microstructures [28]. This in turn provides a better understanding of the appropriate selection of the host matrix to improve the spectral attributes of REIs. In addition, most of the reported literatures [17,28,29] on the structural, and impedance correlation in phosphate-based systems free of REIs. However, by ascertaining such relationship in the REIs doped glasses or GCs an indepth understanding of the microscopic mechanisms can be developed.

Based on this background, this thesis took an attempt to evaluate the structural, microstructural, impedance and optical characteristics of the Ag NPs (varied size) and Ho³⁺ co-doped zinc-sulfo-boro-phosphate GCs nanocomposites (GCNCs). The main goal is to determine the lasing potency of Ho³⁺ in the newly composed system.

1.2 Problem Statement

The more the change between the local surroundings of all Ho³⁺ distributed in the phosphate-based system, the broader the spectral peaks. This in turn lowers the stimulated emission cross-section that needs to be enhanced for high optical gain laser applications. The optimum composition with efficient lasing action is required and remains an open problem. Creating some crystalline domains within the glassy matrix is believed to attain strong optical response. Interestingly, the mechanism of Ag NPs (varying size and contents) that enables LSPR assisted lasing potency in phosphatebased nanocomposites is critical to obtain the optimum composition. The composition optimization is pre-requisite to determine the modified overall properties. Therefore, the optimization of Ho³⁺ as well as Ag NPs concentration inside the phosphate-based nanocomposites need to be systematically carried out. The lasing potency of Ho³⁺ doped various host materials are primarily determined by their surrounding structures and microstructures. Thus, detail analyses are required in order to understand these basic quantities. Over the years, although diverse studies have been carried out on Ho³⁺ -activated host matrices, seldom studies have focused on its structural, microstructural, impedance and optical properties in zinc-sulfo-boro-phosphate nanocomposite for the development of Ho³⁺ based inexpensive, visible and eye-safe laser.

It is known that the Judd-Ofelt (J-O) analyses are important tools to understand the host material and the structural properties surrounding the REI as well as to determine the lasing potency via some radiative parameters. Stimulated emission cross section is the most critical parameter to decide the feasibility of achieving efficient lasing action. However, the J-O intensity and radiative parameters for Ho³⁺ -activated phosphate-based nanocomposite with Ag NPs sensitization has not yet been explored in-depth.

Interestingly, Ho³⁺ has several close-lying excited energy levels over the visible spectral region responsible for intense visible spectral transitions. However, the non-radiative processes associated with various relaxational mechanisms of these excited states lead to energy loss often limit practical applications of holmium. To surmount such shortcomings, based on the fact that is the electrical field of the host environment plays a role in the 4f energy level splitting to several sublevels, a better understanding on the appropriate selection of the host matrix is necessary. To achieve this perspective, a basic knowledge on the structure and microstructure of the chosen host material is mandatory. In this spirit, the impedance analysis is often recommended to provide useful information about the network microstructure, alongside with some impedance properties. This in turn, can be used to elucidate the relaxation mechanisms concerning the carriers transport properties in the materials under study. Therefore, detail studies on the structural, microstructural, impedance and optical properties of Ho^{3+} -activated phosphate-based nanocomposites with and without the presence of Ag NPs need to be determined so that a possible correlation amid the abovementioned traits can be established.

1.3 Research Objectives

Based on the above problem statement, the following objectives are set.

- (a) To optimize the composition of Ho³⁺ -activated zinc-sulfo-boro-phosphate GC system without and with Ag NPs sensitization.
- (b) To determine the influence of Ho³⁺ content on the structural, microstructural and optical properties of the proposed GCs and GCNCs.
- (c) To evaluate the lasing potency of the optimum GCNC (sample with highest PL intensity from each series) via Judd-Ofelt intensity and radiative parameters for supporting the experimental optical data.
- (d) To correlate the structural, microstructural, impedance and optical properties of the proposed GCs and GCNCs.

1.4 Scope of Research

To achieve the set objectives, the following scopes are included.

- (a) Selection of appropriate amount of chemicals for every 20 gram batch in three series of samples. The first GCs series is composed with varying Ho₂O₃ content and without Ag NPs embedment, while the second and third GCNCs series are formulated with changing Ag NPs concentration (Ag NPs of different mean size) at fixed Ho³⁺ content.
 - (40-x) P₂O₅ 30 B₂O₃ 30 ZnSO₄ x Ho₂O₃, (x = 0.0, 0.2, 0.4, 0.5, 0.6, 0.8 and 1.0 mol%).
 - (40-x-y) P₂O₅ 30 B₂O₃ 30 ZnSO₄ x Ho₂O₃ y Ag nanopowder, (x = the best mol% selected from (1) with highest PL intensity in visible spectral region. y = 0.6, 0.7, 0.8 and 0.9 mol%).

- (40-x-z) P₂O₅ 30 B₂O₃ 30 ZnSO₄ x Ho₂O₃ z AgCl, (z = 0.6, 0.7, 0.8 and 0.9 mol%).
- (b) Preparation of the mentioned three series of samples using melt quenching method.
- (c) Determination of sample density by Archimedes method.
- (d) Characterization of the structural features using XRD measurement, FTIR, Raman, XPS, EDX spectroscopies.
- (e) Determination of the microstructure characteristics by:
 - HRTEM that further identifies the existence (and morphology) of the nanoparticles.
 - IS that provides the supportive impedance data.
- (f) Characterization of optical (absorption and emission) features using UV-Vis-NIR and PL spectroscopies.
- (g) Assessment of the lasing potency of the synthesized optimum nanocomposites containing both Ho³⁺ and Ag NPs using J-O analysis (in terms of intensity and radiative parameters).

1.5 Significance of Research

- (a) New nanocomposites with optimum composition have been produced as an alternative solid-state lasing media with relatively strong lasing potency that may be useful for the devolvement of various nanophotonic devices.
 - The obtained large values of the branching ratio and respective stimulated emission cross-section (73.88 % and 46.68×10⁻²¹ cm² for the green; and 83.97 % and 41.12×10⁻²¹ cm² for the red) were greater than most of the

state-of-the-art reports indicated the effectiveness of the proposed samples for the construction of green and red lasers.

- The attained strong lasing potency of the IR transition (branching ratio: 71.40%; stimulated emission cross-section: 36.95×10⁻²¹ cm²) demanded for many photonic applications was greater than the existing one (comparative evaluations were given hereinafter in page 120).
- (b) New knowledge has been generated on the relationship among structural, impedance and optical properties in the studied nanocomposites. The optical properties can be modified by developing such correlations. For instance, the polarons that were responsible for the conduction mechanism might play a significant role in the ET mechanism.
- (c) The mechanism of the Ag NPs localized surface plasmon resonance and its influence on the enhancement of the radiative properties of Ho³⁺ in the titled system has been better understood where the impact of embedding Ag NPs with two different sizes on the Ho³⁺ transitions (while maintaining the local structure relatively fixed) was explored for the first time.

1.6 Thesis Outline

The thesis is organized as follows: Chapter 1 presents the background of the study, statement of the problems, research objectives, scope of the research and its significance. Chapter 2 provides an overview of the previous related literature. Chapter 3 displays the research methodology. Chapter 4 represents and discusses the research results. Finally, chapter 5 summarizes the findings with respect to the set research objectives, as well as some future research suggestions.

REFERENCES

- Kenyon A.J. Recent developments in rare-earth doped materials for optoelectronics. *Prog. Quantum. Electron.* 2002. 26(4–5): 225–84.
- [2] Morassuti C.Y., Andrade L.H.C., Silva J.R., Baesso M.L., Guimarães F.B., Rohling J.H., Nunes L.A.O., Boulon G., Guyot Y. and Lima S.M. Spectroscopic investigation and interest of Pr³⁺–doped calcium aluminosilicate glass. *J. Lumin.* 2019. 210 (February): 376–82.
- Li B., Liang J., Sun L., Wang S., Sun Q., Devakumar B., Annadurai G., Chen D., Huang X. and Wu Y. Cyan-emitting Ba₃Y₂B₆O₁₅ :Ce³⁺, Tb³⁺ phosphor: A potential color converter for near-UV-excited white LEDs. *J. Lumin.* 2019. 211(April): 388–93.
- [4] Henderson B. and Imbusch G.F. *Optical spectroscopy of inorganic solids*. New York: Oxford. 1989.
- [5] Prakash M.R., Neelima G., Kummara V.K., Ravi N., Viswanath C.D., Rao T.S. and Jilani S.M. Holmium doped bismuth-germanate glasses for green lighting applications: A spectroscopic study. *Opt. Mater.* 2019. 94 (April): 436–43.
- [6] Jupri S.A., Ghoshal S.K., Omar M.F. and Yusof N.N. Spectroscopic traits of holmium in magnesium zinc sulfophosphate glass host: Judd–Ofelt evaluation. *J. Alloys. Compd.* 2018. 753 (6): 446–56.
- [7] Jupri S.A., Ghoshal S.K., Yusof N.N., Omar M.F., Hamzah K. and Krishnan G. Influence of surface plasmon resonance of Ag nanoparticles on photoluminescence of Ho³⁺ ions in magnesium–zinc–sulfophosphate glass system. *Opt. Laser. Technol.* 2020. 126: 106134.
- [8] Weber M.J. Rare Earth Lasers. In: Gschneidner K.A., Jr and Eyring L. Handbook on the Physics and Chemistry of Rare Earths. Netherlands: North– Holland Publishing Company. 361–369; 1979.
- [9] Sdiri N., Elhouichet H. and Ferid M. Effects of substituting P₂O₅ for B₂O₃ on the thermal and optical properties of sodium borophosphate glasses doped with Er. J. Non. Cryst. Solids. 2014. 389: 38–45.
- [10] Zhang H., Yuan D., Mi X., Zhang X., Bai Z., Liu X. and Lin J. Crystal structure and luminescence properties of Na₂ MMg (PO₄)₂: Eu²⁺ (M = (Ca/Sr/Ba)

phosphors. J. Alloys Compd. 2019. 798: 119-28.

- [11] Liu Y., Song F., Jia G., Zhang Y. and Tang Y. Strong emission in Yb³⁺/Er³⁺ codoped phosphate glass ceramics. *Results Phys.* 2017. 7: 1987–92.
- [12] Yu X., Duan L., Ni L. and Wang Z. Fabrication and luminescence behavior of phosphate glass ceramics co-doped with Er³⁺ and Yb³⁺. *Opt. Commun.* 2012. 285 (18): 3805–8.
- [13] Shu-jiang L., An-xian L., Xiao-dong T. and Shao-bo H. Influence of addition of B₂O₃ on properties of Yb³⁺ -doped phosphate laser glass. *J. Cent. South Univ. Technol.* 2006. 13 (5): 1–5.
- [14] Brow R.K. An XPS study of oxygen bonding in zinc phosphate and zinc borophosphate glasses. J. Non. Cryst. Solids. 1996. 194: 267–73.
- [15] Kim Y., Choi W. and Ryu B. Effect of ZnO content change on the structure and properties of zinc borophosphate glasses. *Glas. Phys. Chem.* 2014. 40 (4): 408–14.
- [16] Xiu-ying L., Hua-ming Y. and Yu-xi R. Effects of Zn/P ratio on structures and properties of zinc-iron phosphate glasses. J. Cent. South. Univ. 2013. 20 (1): 44–9.
- [17] Sadok K.H., Haouari M., Gallot–lavall O. and Ouada H.B. Effect of Na₂SO₄ substitution for Na₂O on the structural and electrical properties of a sodium borophosphate glass. J. Alloys Compd. 2019. 778: 878–88.
- [18] Da N., Grassmé O., Nielsen K.H., Peters G. and Wondraczek L. Formation and structure of ionic (Na, Zn) sulfophosphate glasses. *J. Non. Cryst. Solids*. 2011.
 357 (10): 2202–6.
- [19] Lai Y.M., Liang X.F., Yang S.Y., Wang J.X. and Zhang B.T. Raman spectra study of iron phosphate glasses with sodium sulfate. *J. Mol. Struct.* 2012. 1013: 134–7.
- [20] Naumov A.V. Review of the world market of rare-earth metals. *Russ. J. Non– Ferrous Met.* 2008. 49 (1): 14–22.
- [21] Babu Y.N.C.R., Ramnaik P.S. and Kumar A.S. Photoluminescence features of Ho³⁺ ion doped PbO–Bi₂O₃ borophosphate glass systems. *J. Lumin.* 2013. 143: 510–6.
- [22] Kumar V.R., Giridhar G., Sudarsan V. and Veeraiah N. Influence of red lead on the intensity of green and orange emissions of Sm³⁺ and Ho³⁺ co-doped ZnO– SrO–P₂O₅ glass system. J. Alloys. Compd. 2017. 695: 668–81.

- [23] Vijayakumar R. and Marimuthu K. Luminescence studies on Ag nanoparticles embedded Eu³⁺ doped Boro–phosphate glasses. J. Alloys. Compd. 2016. 665: 294–303.
- [24] Soltani I., Hraiech S., Elhouichet H., Gelloz B. and Férid M. Growth of silver nanoparticles stimulate spectroscopic properties of Er³⁺ doped phosphate glasses: Heat treatment effect. *J. Alloys. Compd.* 2016. 686: 556–563
- [25] Ahmadi F., Hussin R. and Ghoshal S.K. Spectroscopic attributes of Sm³⁺ doped magnesium zinc sulfophosphate glass: Effects of silver nanoparticles inclusion. *Opt. Mater.* 2017. 73: 268–76.
- [26] Rivera V.A.G., Osorio S.P.A., Ledemi Y., Manzani D., Messaddeq Y., Nunes
 L.A.O. and Jr. E.M. Localized surface plasmon resonance interaction with Er³⁺
 -doped tellurite glass. *Opt. Express.* 2010. 18 (24): 25321–25328.
- [27] Rivera V.A.G., Ferri F.A. and Jr. E.M. localized surface plasmon resonances: Noble metal nanoparticle interaction with rare-earth ions. In: *Plasmonics– Principles and Applications*. InTech. 284–312; 2012.
- [28] Pavić L., Skoko Ž., Gajović A., Su D. and Moguš-Milanković A. Electrical transport in iron phosphate glass-ceramics. J. Non. Cryst. Solids. 2018. 502 (January): 44–53.
- [29] Barde R.V., Nemade K.R. and Waghuley S.A. AC conductivity and dielectric relaxation in V₂O₅-P₂O₅-B₂O₃ glasses. *J. Asian. Ceram. Soc.* 2015. 3 (1): 116–22.
- [30] Liu S., Song F., Cai H., Li T., Tian B. and Wu Z. Effect of thermal lens on beam quality and mode matching in LD pumped Er–Yb -codoped phosphate glass. 2008. 41(3): 035104.
- [31] Jaeger T.D. Optical Manufacturing : Material improvements bring laser glasses into mainstream applications. *Laser Focus World*. 2015. 51 (07).
- [32] Ahmadi F., Hussin R. and Ghoshal S.K. Optical transitions in Dy³⁺-doped magnesium zinc sulfophosphate glass. *J. Non. Cryst. Solids.* 2016. 452: 266–72.
- [33] Ahmadi F., Hussin R. and Ghoshal S.K. Judd–Ofelt intensity parameters of samarium–doped magnesium zinc sulfophosphate glass. J. Non. Cryst. Solids. 2016. 448: 43–51.
- [34] Ahmadi F., Hussin R. and Ghoshal S.K. Spectral characteristics of Er³⁺ doped magnesium zinc sulfophosphate glasses. J. Alloys Compd. 2017. 711: 94–102.

- [35] Danmallam I.M., Ghoshal S.K., Ariffin R., Jupri S.A. and Sharma S. Europium ions and silver nanoparticles co-doped magnesium–zinc–sulfophosphate glasses: Evaluation of ligand field and Judd-Ofelt parameters. *J. Lumin.* 2019. 216: 116713.
- [36] Denker B. and Shklovsky E. *Handbook of solid-state lasers*. Woodhead Publishing. 2013.
- [37] Shang F., Chen Y., Xu J., Yang T., Yang Y. and Chen G. Up-conversion luminescence and highly sensing characteristics of Er³⁺/Yb³⁺ co-doped borophosphate glass-ceramics. *Opt. Commun.* 2019. 441 (December): 38–44.
- [38] Ming C., Song F., An L., Ren X., Yuan Y., Cao Y. and Wang G. Research on up- and down-conversion emissions of Er³⁺/Yb³⁺ co-doped phosphate glass ceramic. *Opt. Mater.* 2012. 35 (2): 244–7.
- [39] Hermansen C., Mauro J.C. and Yue Y. A model for phosphate glass topology considering the modifying ion sub- network. J. Chem. Phys. 2014. 140 (154501): 1–8.
- [40] Yun Y.H. and Bray P.J. Nuclear magnetic resonance studies of the glasses in the system K₂O-B₂O₃-P₂O₅. J. Non. Cryst. Solids. 1978. 30: 45–60.
- [41] Zhang L., Sun H., Xu S., Li K. and Hu L. Influence of B₂O₃ to the inhomogeneous broadening and spectroscopic properties of Er³⁺/Yb³⁺ codoped fluorophosphate glasses. *J. Lumin.* 2006. 117: 46–52.
- [42] Baikova L.G., Fedorov Y.K., Pukh V.P., Tikhonova L.V., Kazannikova T.P., Sinani A.B. and Nikitina S.I. Influence of boron oxide on the physicomechanical properties of glasses in the Li₂O–B₂O₃–P₂O₅ system. *Glas. Phys. Chem.* 2003. 29 (3): 276–81.
- [43] Takebe H. Effect of B₂O₃ addition on the thermal properties and density of barium phosphate glasses. J. Non. Cryst. Solids. 2006. 352: 709–13.
- [44] Brow R.K. and Tallant D.R. Structural design of sealing glasses. J. Non. Cryst. Solids. 1997. 222: 396–406.
- [45] Kumar G.A., Martinez A. and Rosa E.D.L. Stimulated emission and radiative properties of Nd³⁺ ions in barium fluorophosphate glass containing sulphate. J. *Lumin.* 2002. 99 (2): 141–8.
- [46] Sułowska J., Wacławska I. and Szumera M. Comparative study of zinc addition effect on thermal properties of silicate and phosphate glasses. J. Therm. Anal. Calorim. 2016. 123 (2): 1091–8.

- [47] Yamusa Y.A., Hussin R. and Shamsuri W.N.W. Physical, optical and radiative properties of CaSO₄–B₂O₃–P₂O₅ glasses doped with Sm³⁺ ions. *Indian J. Phys.* 2018. 56 (3): 932–43.
- [48] Das S., Ghosh A. Structure and electrical properties of vanadium borophosphate glasses. J. Non. Cryst. Solids. 2017. 458 (October): 28–33.
- [49] Irvine J.T.S., Sinclair D.C. and West A.R. Electroceramics Characterisation by impedance sepctroscopy. *Adv. Mat.* 1990. 2 (3): 132–8.
- [50] Balaram V. Rare earth elements: A review of applications, occurrence, exploration, analysis, recycling, and environmental impact. *Geosci. Front.* 2019. 10 (4): 1285–303.
- [51] Brongersma M.L. and Kik P.G. Surface Plasmon Nanophotonics. In: Lotsch H.K.V. Optical Sciences. Netherlands: Springer. 131–268; 2007.
- [52] Srinivasa R.C., Upendra K.K., Babu P. and Jayasankar C.K. Optical properties of Ho³⁺ ions in lead phosphate glasses. *Opt. Mater.* 2012. 35 (2): 102–107.
- [53] Seshadri M., Barbosa L.C. and Radha M. Study on structural, optical and gain properties of 12 and 20 μm emission transitions in Ho³⁺ doped tellurite glasses. *J. Non. Cryst. Solids.* 2014. 406: 62–72.
- [54] Venkateswarlu M., Mahamuda S., Swapna K., Prasad M.V.V.K.S., Rao A.S., Shakya S., Babu A.M. and Prakash G.V. Holmium doped lead tungsten tellurite glasses for green luminescent applications. *J. Lumin.* 2015. 163: 64–71.
- [55] Laxmikanth C., Anjaiah J., Rao P.V., Rao B.A. and Veeraiah N. Luminescence and spectroscopic properties of ZnF₂–MO–TeO₂ glasses doped with Ho³⁺ ions. *J. Mol. Struct.* 2015. 1093: 166–71.
- [56] Rai S. and Fanai A.L. Optical properties of Ho³⁺ in sol-gel silica glass co-doped with Aluminium. J. Non. Cryst. Solids. 2016. 449: 113–8.
- [57] Kesavulu C.R., Kim H.J., Lee S.W., Kaewkhao J., Wantana N., Kothan S. and Kaewjaeng S. Optical spectroscopy and emission properties of Ho³⁺-doped gadolinium calcium silicoborate glasses for visible luminescent device applications. *J. Non. Cryst. Solids.* 2017. 474 (April): 50–7.
- [58] Damodaraiah S., Lakshmi R.P.V. and Ratnakaram Y.C. Role of bismuth content on structural and luminescence properties of Ho³⁺ doped phosphate glasses. *J. Mol. Struct.* 2020. 1200: 127157.
- [59] Jayachandra P. T., Neelima G., Ravi N., Kiran N., Nallabala N. K. R., Kummara V. K., Suresh K. and Gadige P. Optical and spectroscopic properties of Ho³⁺-

doped fluorophosphate glasses for visible lighting applications. *Mater. Res. Bull.* 2020. 124: 110753.

- [60] Burda C., Chen X., Narayanan R. and El-sayed M.A. Chemistry and properties of nanocrystals of different shapes. *Chem. Rev.* 2005. 105 (4): 1025–102.
- [61] Marchuk K. and Willets K.A. Localized surface plasmons and hot electrons. *Chem. Phys.* 2014. 445: 95–104.
- [62] Lorenz K., Alves E., Gloux F. and Ruterana P. RE implantation and annealing of III-nitrides. *Top. Appl. Phys.* 2010. 124: 25–54.
- [63] Manjavacas A., Liu J.G., Kulkarni V. and Nordlander P. Plasmon-induced hot carriers in metallic nanoparticles. ACS Nano, 2014. 8 (8): 7630–8.
- [64] Som T. and Karmakar B. Nano silver : antimony glass hybrid nanocomposites and their enhanced fluorescence application. *Solid State Sci.* 2011. 13 (5): 887–95.
- [65] Gan F. and Xu L. *Photonic glasses*. World Scientific Publishing. 2006.
- [66] Moustafa S.Y., Sahar M.R. and Ghoshal S.K. Spectroscopic attributes of Er³⁺ ions in antimony phosphate glass incorporated with Ag nanoparticles: Judd– Ofelt analysis. J. Alloys Compd. 2017. 712: 781–94.
- [67] Mahraz Z.A.S., Sahar M.R. and Ghoshal S.K. Enhanced luminescence from silver nanoparticles integrated Er³⁺-doped boro-tellurite glasses: Impact of annealing temperature. *J. Alloys Compd.* 2015. 649: 1102–9.
- [68] Amjad R.J., Dousti M.R., Sahar M.R., Shaukat S.F., Ghoshal S.K., Sazali E.S. and Nawaz F. Silver nanoparticles enhanced luminescence of Eu³⁺-doped tellurite glass. *J. Lumin.* 2014. 154: 316–21.
- [69] Amjad R.J., Dousti M.R. and Sahar M.R. Spectroscopic investigation and Judd– Ofelt analysis of silver nanoparticles embedded Er³⁺-doped tellurite glass. *Curr. Appl. Phys.* 2015. 15 (1): 1–7.
- [70] Ahmadi F., Hussin R. and Ghoshal S.K. On the optical properties of Er³⁺ ions activated magnesium zinc sulfophosphate glass: Role of silver nanoparticles sensitization. *J. Lumin.* 2018. 204 (April): 95–103.
- [71] Yusoff N.M. and Sahar M.R. The incorporation of silver nanoparticles in samarium doped magnesium tellurite glass: Effect on the characteristic of bonding and local structure. *Phys. B: Condens. Matter.* 2015. 470–471: 6–14.
- [72] Soltani I., Hraiech S., Horchani-Naifer K., Elhouichet H. and Férid M. Effect of silver nanoparticles on spectroscopic properties of Er³⁺ doped phosphate

glass. Opt. Mater. 2015. 46: 454-60.

- [73] Haouari M., Slimen F.B., Maaoui A. and Gaumer N. Structural and spectroscopic properties of Eu³⁺ doped tellurite glass containing silver nanoparticles. *J. Alloys Compd.* 2018. 743: 586–96.
- [74] Dousti M.R., Sahar M.R., Ghoshal S.K., Amjad R.J. and Ari R. Up-conversion enhancement in Er³⁺–Ag co-doped zinc tellurite glass : Effect of heat treatment. *J. Non. Cryst. Solids.* 2012. 358: 2939–42.
- [75] Saad M., Stambouli W., Mohamed S.A. and Elhouichet H. Ag nanoparticles induced luminescence enhancement of Eu³⁺ doped phosphate glasses. J. Alloys Compd. 2016. 705: 550–8.
- [76] Qi Y., Zhou Y., Wu L., Yang F., Peng S., Zheng S. and Yin D. Enhanced upconversion emissions in Ho³⁺/Yb³⁺ codoped tellurite glasses containing silver NPs. J. Non. Cryst. Solids. 2014. 402: 21–7.
- [77] Carnall W.T., Fields P.R. and Rajnak K. Electronic energy levels of the trivalent lanthanide aquo ions IV Eu⁸⁺. J. Chem. Phys. 1968. 49 (10): 4424–42.
- [78] Dodson C.M. and Zia R. Magnetic dipole and electric quadrupole transitions in the trivalent lanthanide series: Calculated emission rates and oscillator strengths. *Phys. Rev. B- Condens. Matter. Mater. Phys.* 2012. 86 (12): 1–10.
- [79] Shannon R.D. Revised effective ionic radii and systematic studies of interatomie distances in halides and chaleogenides. *Acta. Cryst.* 1976. 203: 751.
- [80] Bulus I., Dalhatu S.A., Hussin R., Shamsuri W.N.W. and Yamusa Y.A. The role of dysprosium ions on the physical and optical properties of lithium– borosulfophosphate glasses. *Int. J. Mod. Phys. B.* 2017. 31 (13): 1750101.
- [81] Moustafa Y.M. and El-Egili K. Infrared spectra of sodium phosphate glasses. J. Non. Cryst. Solids. 1998. 240 (1–3): 144–53.
- [82] Bulus I., Hussin R., Ghoshal S.K., Tamuri A.R. and Jupri S.A. Enhanced elastic and optical attributes of boro-telluro-dolomite glasses: Role of CeO₂ doping. *Ceram. Int.* 2019. 45 (15): 18648-18658.
- [83] Abdellaoui K., Ratep A., Boumaza A., Kashif I. and Donya H. The effect of the natural raw barite and the dolomite material on borate glass formation. J. Fundam. Appl. Sci. 2018. 10 (1): 281–300.
- [84] Ibrahim A.M., Hammad A.H., Abdelghany A.M. and Rabie G.O. Mixed alkali effect and samarium ions effectiveness on the structural, optical and non-linear optical properties of borate glass. J. Non. Cryst. Solids. 2018. 495 (February):

67–74.

- [85] Möncke D., Sirotkin S., Stavrou E., Kamitsos E.I. and Wondraczek L. Partitioning and structural role of Mn and Fe ions in ionic sulfophosphate glasses partitioning and structural role of Mn and Fe ions in ionic sulfophosphate glasses. J. Chem. Phys. 2014. 141 (22): 224509.
- [86] Lakshminarayana G., Kaky K.M., Baki S.O., Lira A., Meza-Rocha A.N., Falcony C., Caldino U., Kityk I.V., Mendez-Blas A., Abas A.F., Alresheedi M.T. and Mahdi M.A. Nd³⁺-doped heavy metal oxide based multicomponent borate glasses for 106 μm solid-state NIR laser and O-band optical amplification applications. *Opt. Mater.* 2018. 78: 142–59.
- [87] Ahmadi F., Hussin R. and Ghoshal S.K. Structural and physical properties of Sm³⁺ doped magnesium zinc sulfophosphate glass. *Bull. Mater. Sci.* 2017. 40 (6): 1097–104.
- [88] Pal M., Roy B. and Pal M. Structural characterization of borate glasses containing zinc and manganese oxides. *Int. J. Mod. Phys.* 2011. 2 (9): 1062–6.
- [89] Aliyu A.M., Hussin R., Ahmad N.E. and Yamusa Y.A. Samarium doped calcium sulfate ultra-phosphate glasses: Structural, physical and luminescence investigations. *Optik*. 2018. 172 (July): 1162–71.
- [90] Samdani, Ramadevudu G., Chary M.N. and Shareefuddin M. Physical and spectroscopic studies of Cr³⁺doped mixed alkaline earth oxide borate glasses. *Mater. Chem. Phys.* 2017. 186: 382–9.
- [91] Matzkeit Y.H., Tornquist B.L., Manarin F., Botteselle G.V., Rafique J., Saba S., Braga A.L., Felix J.F. and Schneider R. Borophosphate glasses: Synthesis, characterization and application as catalyst for bis(indolyl)methanes synthesis under greener conditions. *J. Non. Cryst. Solids.* 2018. 498 (June): 153–9.
- [92] Yang Q., Sha J., Wang L., Zou Y., Niu J., Cui C. and Yang D. Crystalline boron oxide nanowires on silicon substrate. *Phys. E: Low-Dimensional Syst. Nanostructures.* 2005. 27 (3): 319–24.
- [93] Wang Y., Yu Y., Zou Y., Zhang L., Hu L. and Chen D. Broadband visible luminescence in tin fluorophosphate glasses with ultra-low glass transition temperature. *RSC Adv.* 2018. 8 (9): 4921–7.
- [94] Nurhafizah H., Rohani M.S. and Ghoshal S.K. Self cleanliness of Er³⁺/Nd³⁺ codoped lithium niobate tellurite glass containing silver nanoparticles. J. Non. Cryst. Solids. 2017. 455: 62–9.

- [95] Shaaban M.H., Ali A.A. and El-Nimr M.K. The AC conductivity of tellurite glasses doped with Ho₂O₃. *Mater. Chem. Phys.* 2006. 96: 433–8.
- [96] Abouhaswa A.S., Rammah Y.S., Ibrahim S.E. and El-Hamalawy A.A. Structural, optical, and electrical characterization of borate glasses doped with SnO₂. J. Non. Cryst. Solids. 2018. 494 (March): 59–65.
- [97] Inaba S. and Fujino S. Empirical equation for calculating the density of oxide glasses. J. Am. Ceram. Soc. 2010. 93 (1): 217–20.
- [98] Sumang R., Bongkarn T., Pimpang P. and Thongmee N. Correlation of structural, microstructure and dielectric properties of substituted and unsubstituted CaCu₃Ti_{4-x}A_xO₁₂ ceramics. *Ferroelectrics*. 2019. 552 (1): 84–94.
- [99] Jlassi I., Sdiri N., Elhouichet H. and Ferid M. Raman and impedance spectroscopy methods of P₂O₅-Li₂O-Al₂O₃ glass system doped with MgO. J. Alloys Compd. 2015. 645: 125–30.
- [100] Carnall W.T., Fields P.R. and Rajnak K. Electronic energy levels in the trivalent lanthanide aquo ions I. Pr³⁺, Nd³⁺, Pm³⁺, Sm³⁺, Dy³⁺, Ho³⁺, Er³⁺, and Tm³⁺. J. Chem. Phys. 1968. 49 (10): 4424–42.
- [101] Khor S.F., Talib Z.A. and Yunus W.M.M. Optical properties of ternary zinc magnesium phosphate glasses. *Ceram. Int.* 2012. 38 (2): 935–40.
- [102] Mott N.F. and Davis E.A. *Electronic Processes in Non-Crystalline Materials*. Oxford: Clarendon Press. 2012.
- [103] Hager I.Z. Optical properties of lithium barium haloborate glasses. J. Phys. Chem. Solids. 2009. 70: 210–7.
- [104] Davis E.A., Mott N.F. and Davis E.A. Conduction in non-crystalline systems V.conductivity, optical absorption and photoconductivity in amorphous semiconductors. *Philos. Mag.* 1970. 22 (179): 903–22.
- [105] Oliveira L.C., Lima A.M.N., Thirstrup C. and Neff H.F. Surface plasmon resonance sensors: a materials guide to design and optimization. New York: Springer. 2015.
- [106] Peng S., Mcmahon J.M., Schatz G.C., Gray S.K. and Sun Y. Reversing the sizedependence of surface plasmon resonances. *Proc. Natl. Acad. Sci. U. S. A.* 2010. 107 (33): 14530–4.
- [107] Dimitrov V. and Sakka S. Electronic oxide polarizability and optical basicity of simple oxides I. J. Appl. Phys. 1996. 79 (3): 1736–40.
- [108] Urbach F. The long-wavelength edge of photographic sensitivity and of the

electronic absorption of solids. Phys. Rev. 1953. 92 (5): 1324-1324.

- [109] Babu S., Seshadri M., Balakrishna A., Reddy Prasad V. and Ratnakaram Y.C. Study of multicomponent fluoro-phosphate based glasses: Ho³⁺ as a luminescence center. *Phys. B: Condens. Matter.* 2015. 479: 26–34.
- [110] Pandey A. and Swart H.C. Luminescence investigation of visible light emitting Ho³⁺ doped tellurite glass. J. Lumin. 2016. 169: 93–8.
- [111] Opdenbosch D.V., Kostova M.H., Gruber S., Krolikowski S., Greil P. and Zollfrank C. Optical properties of manganese-doped nanocrystals of ZnS. *Wood Sci. Technol.* 2010. 44 (4): 416–9.
- [112] Ronda C.R. Luminescence: From theory to applications. Wiley-VCH. 2007.
- [113] Ju G., Hu Y., Chen L., Wang X. and Mu Z. Concentration quenching of persistent luminescence. *Phys. B: Condens. Matter.* 2013. 415: 1–4.
- [114] Dousti M.R., Sahar M.R., Ghoshal S.K., Amjad R.J. and Arifin R. Plasmonic enhanced luminescence in Er³⁺:Ag co-doped tellurite glass. *J. Mol. Struct.* 2013. 1033: 79–83.
- [115] Mahraz Z.A.S., Sahar M.R. and Ghoshal SK. Reduction of non-radiative decay rates in boro-tellurite glass via silver nanoparticles assisted surface plasmon impingement: Judd–Ofelt analysis. J. Lumin. 2017. 190 (April): 335–43.
- [116] Jlassi I., Elhouichet H., Ferid M., Chtourou R. and Oueslati M. Study of photoluminescence quenching in Er³⁺-doped tellurite glasses. *Opt. Mater.* 2010. 32 (7): 743–7.
- [117] Hughes M.A., Li H., Curry R.J., Suzuki T. and Ohishi Y. Energy transfer in Cr and Nd co-doped Si–B–Na–Al–Ca–Zr–O glasses. J. Non. Cryst. Solids. 2020. 530 (September): 119769.
- [118] Choi J.H., Margaryan A., Margaryan A. and Shi F.G. Judd–Ofelt analysis of spectroscopic properties of Nd³⁺-doped novel fluorophosphate glass. *J. Lumin*. 2005. 114: 167–77.
- [119] Auzel F. and Goldner P. Towards rare-earth clustering control in doped glasses. Opt. Mater. 2001. 16: 93–103.
- [120] Yusof N.N., Ghoshal S.K. and Azlan M.N. Optical properties of titania nanoparticles embedded Er³⁺-doped tellurite glass: Judd–Ofelt analysis. J. Alloys Compd. 2017. 724: 1083–92.
- [121] Awang A., Ghoshal S.K., Sahar M.R., Arifin R. and Nawaz F. Non-spherical gold nanoparticles mediated surface plasmon resonance in Er³⁺ doped zinc–

sodium tellurite glasses: Role of heat treatment. J. Lumin. 2014. 149: 138-43.

- [122] Neto A.N.C., Couto M.A. and Reisfeld R. Effects of spherical metallic nanoparticle plasmon on 4f-4f luminescence: a theoretical approach. Elsevier Ltd. 2018.
- [123] Zhang W., Lin J., Cheng M. and Zhang S. Radiative transition, local field enhancement and energy transfer microcosmic mechanism of tellurite glasses containing Er³⁺, Yb³⁺ ions and Ag nanoparticles. *J. Quant. Spectrosc. Radiat. Transf.* 2015. 159: 39–52.
- [124] Eichelbaum B.M. and Rademann K. Plasmonic enhancement or energy transfer? On the luminescence of gold-, silver-, and lanthanide-doped silicate glasses and its potential for light-emitting devices. *Adv. Funct. Mater. Adv. Funct. Mater.* 2009. 19 (13): 2045–52.
- [125] Lee T.H., Kwon Y.K. and Heo J. Local structure and its effect on the oscillator strengths and emission properties of Ho³⁺ in chalcohalide glasses. J. Non. Cryst. Solids. 2008. 354 (27): 3107–12.
- [126] Mahamuda S., Swapna K., Packiyaraj P., Rao A.S. and Prakash G.V. Visible red, NIR and Mid-IR emission studies of Ho³⁺ doped zinc alumino bismuth borate glasses. *Opt. Mater.* 2013. 36 (2): 362–71.
- [127] Seshadri M., Ratnakaram Y.C., Naidu D.T. and Rao K.V. Investigation of spectroscopic properties (absorption and emission) of Ho³⁺ doped alkali, mixed alkali and calcium phosphate glasses. *Opt. Mater.* 2010. 32 (4): 535–42.
- [128] Awang A., Ghoshal S.K., Sahar M.R. and Arifin R. Gold nanoparticles assisted structural and spectroscopic modification in Er³⁺ -doped zinc sodium tellurite glass. *Opt. Mater.* 2015. 42: 495–505.
- [129] Fang Y., Hu L., Liao M. and Wen L. Effect of bismuth oxide on the thermal stability and Judd–Ofelt parameters of Er³⁺/Yb³⁺ co-doped aluminophosphate glasses. *Spectrochim. Acta A: Mol. Biomol. Spectrosc.* 2007. 68 (3): 542–7.
- [130] Jørgensen C.K. and Reisfeld R. Judd–Ofelt parameters and chemical bonding. J. Less-Common. Met. 1983. 93 (1): 107–12.
- [131] Som T. and Karmakar B. Nephelauxetic effect of low phonon antimony oxide glass in absorption and photoluminescence of rare-earth ions. *Spectrochim. Acta A: Mol. Biomol. Spectrosc.* 2011. 79 (5): 1766–82.
- [132] Reddy C.M., Raju B.D.P., Sushma N.J., Dhoble N.S. and Dhoble S.J. A review on optical and photoluminescence studies of RE^{3+} (RE = Sm, Dy, Eu, Tb and

Nd) ions doped LCZSFB glasses. *Renew. Sustain. Energy. Rev.* 2015. 51: 566–84.

[133] Yamusa Y.A., Hussin R., Shamsuri W.N.W., Tanko Y.A. and Jupri S.A. Impact of Eu³⁺ on the luminescent, physical and optical properties of BaSO₄- B₂O₃-P₂O₅ glasses. *Optik.* 2018. 164: 324–34.

LIST OF PUBLICATIONS

- (a) ISI Journals
 - Publication: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2019. Intense red and green luminescence from holmium activated zinc-sulfo-boro-phosphate glass: Judd-Ofelt evaluation. *Journal of Alloys and Compounds*, 808, p.151706. Available at: https://doi.org/10.1016/j.jallcom.2019.151706.
 - Publication: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2020. Tailored structures and dielectric traits of holmium iondoped zinc-sulpho-boro-phosphate glass ceramics. *Ceramics International*, 46(3), pp.3282–3291. Available at: <u>https://doi.org/10.1016/j.ceramint.2019.10.034</u>.
 - Publication: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2020. Spectral features of Ho³⁺ -doped boro-phosphate glass-ceramics: Role of Ag nanoparticles sensitization. *Journal of Luminescence*, 223, p.117218. Available at: https://doi.org/10.1016/j.jlumin.2020.117218.
 - Publication: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2020. Ag nanoparticles localised surface plasmon field regulated spectral characteristics of Ho³⁺ -doped phosphate-based glass-ceramic. *Results in Physics*, 17(April), p.103102. Available at: https://doi.org/10.1016/j.rinp.2020.103102.
 - Publication: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2020. Customized physical and structural features of phosphatebased glass-ceramics: role of Ag nanoparticles and Ho³⁺ impurities. *Journal of Taibah University for Science*, 14(1), p. 954-962. Available at: <u>https://doi.org/10.1080/16583655.2020.1791536</u>.

(b) Other Contributions

- Conference: Alqarni, Areej S., Hussin, R., and Ghoshal, S.K., 2019. Modified Structure and Impedance Attributes of Holmium Ions Included Phosphate-Based Media: Synergism of Amorphous and Crystalline Phases. *Proceeding of Sixth Academic Conference on Natural Science*, Thai Nguyen University, Vietnam. ISBN 978-604-913-088-5.
- Conference: Alqarni, Areej S., Hussin, R., Alamri, S.N. and Ghoshal, S.K., 2020. A New Host with Customized Intense Lasing Action: Ag Nanoparticles and Ho³⁺ Interplay. *Proceeding of International Conference on Physics and Chemistry of Materials in Novel Engineering Applications,* Coimbatore, India. https://doi.org/10.1063/5.0019893